

METHOD FOR PREPARING LIGHT-ABSORBING  
POLYMERIC COMPOSITIONS

5 Related Application

This application is a continuation-in-part of our Application Serial No. 08/976,206 filed November 21, 1997, which is based upon and claims the priority of provisional application 60/031,478 filed November 27, 1996.

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Background of the Invention

This invention relates to an improved method for preparing light-absorbing polymeric compositions, which are useful as powders or pellets for incorporation into a variety of thermoplastic resins such as cellulose esters, polyesters, polyolefins, polycarbonates, polyamides, etc. by conventional melt or solution blending techniques. The colored thermoplastic resins thus produced have good clarity, good color development, excellent fastness to light and are useful for a variety of end uses where nonhazardous, nonmigrating, or nonextractable colorants are needed.

It is well-known that thermoplastic polymers may be colored by adding pigments or solvent dyes (e.g., see Thomas G. Weber, Editor, *Coloring of Plastics*, John Wiley & Sons, New York, 1979). The use of pigments, however, is accompanied by undesirable properties such as opacity, dullness of color, low tintorial strength, etc. Also, difficulties in uniformly blending the insoluble pigments with the thermoplastic resin are often encountered. Also useful for coloring thermoplastic polymers are the solvent dyes (K. Venkataraman, Editor, *The Chemistry of Synthetic Dyes*, Vol. 8, Academic Press, New York, 1978, pp. 81-131), which provide compositions having improved clarity, brightness in hue and high tintorial strength, but which

may lead to dye migration, extraction, etc. from the colored thermoplastic polymer. These problems are of particular concern when solvent dyes are used to color flexible polymers such as polyvinyl chloride, polyethylene and polypropylene which have low glass transition temperatures.

Plastics, paints, printing inks, rubber, cosmetics, and similar materials are typically colored by organic pigments when superior brilliance and tinctorial strength are important. Toxicity considerations have presented chronic problems relative to the use of organic pigments since some have been shown to be potential carcinogens and to cause contact dermatitis.

Plastics are also colored by using color concentrates consisting of physical admixtures of polymers and colorants (usually solvent dyes). However, the use of such physical admixtures to color polymeric materials such as polyester, e.g., poly(ethylene terephthalate) and blends thereof, present a number of problems, including:

Colorant migration during drying of the colored polyester pellets.

Colorant migration during extrusion and colorant accumulation on dies which can cause shutdowns for clean-up. Such colorant migration and accumulation result in time consuming and difficult clean-up, particularly when a polymer of another color is subsequently processed on the same equipment.

Colorants may not mix well, for example, when using two or more color concentrates to obtain a particular shade.

Colorants may diffuse or exude during storage and use of the colored polymeric material.

The colored polymeric compositions which are prepared by the process of this invention eliminate or minimize the aforementioned problems associated with the use of conventional dyes and pigments.

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Prior Art

To attempt to overcome some of the problems mentioned above, particularly as relates to coloring polyesters, colored polyester compositions have been prepared by  
5 copolymerizing relatively low amounts of monomeric colorants during the polymer preparation (U.S. Pat. Nos. 5,194,571; 5,106,942; 5,102,980; 5,032,670; 4,892,922; 4,740,581; 4,403,092; 4,359,570; 4,267,306 and W092/07913). However, the preparation of these colored  
10 polymers require dyes having outstanding thermal stability since the colorants are exposed to very high temperatures for prolonged periods of time necessary for polyester formation, thus severely circumscribing the selection of efficacious colorants. For example, only the nonazo type  
15 colorants have been shown to have the adequate thermal stability for copolymerization into polyesters, since azo type compounds do not have the requisite thermal stability for copolymerization.

Furthermore, it is known to prepare polymeric dyes by  
20 reacting dyes containing reactive hydroxy and amino groups with organic di-acid chlorides in solvents (U.S. Pat. Nos. 2,994,693; 3,403,200; 4,619,990; 4,778,742; 5,401,612). Although this method of polymer preparation allows the use of a wide range of chromophoric classes, including azo  
25 compounds, as colorant monomers, the polymerization reaction in each case involves the use of very reactive organic di-acid chlorides which are toxic and involve difficult to handle inorganic halogen compounds in their preparation and have accompanying problems of hydrolysis  
30 in the presence of water which causes serious handling and storage problems. The hydrolysis product (HCl) is particularly corrosive and makes storage of these compounds difficult. Furthermore, since the di-acid chlorides will react with water, the monomeric dyes must  
35 be specially dried to avoid side reactions in the polymer preparation.

In a similar attempt to prepare polymeric dyes using relatively low temperatures, polyurethanes have been prepared by reacting dyes bearing two hydroxyalkyl group with aliphatic and aromatic isocyanates (U.S. Pat. 5,194,463). However, the organic isocyanates themselves are extremely toxic and present difficult handling problems. They also are reactive with water and thus the reaction requires specially dried monomeric dyes. Also, the colored polyurethanes as a class do not have excellent thermal stability.

It is further known to prepare colored condensation polymers by reacting a polymerizable lactone or a hydroxyalkanoic acid with a dye containing reactive hydroxy group (U.S. Pat. 4,933,426). The procedure again requires relatively high reaction temperatures and prolonged times and use a large excess of the lactone reactant. The method is further hindered by the fact that some lactones are suspected carcinogens.

Light-absorbing polymeric compositions have also been produced by free radical polymerization of vinyl functionalized light-absorbing monomers (U.S. Pat. Nos. 5,310,837; 5,334,710; 5,359,008; 5,434,231 and 5,461,131).

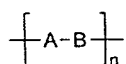
Finally, it is known that one may color plastics, in particular polyolefins, with low melting, cross-linked colored polyester compositions containing residues of terephthalic acid, isophthalic acid, or both, a low-molecular weight trimethylol alkane, i.e., 1,1,1-trimethylol propane and a copolymerizable colorant, said colorant being present at a level of 0.1-25% by weight (U.S. Pat. No. 4,116,923). Difficulties are encountered, however, in preparing these highly cross-linked colored polymers as extreme care with regard to the temperature, amount of vacuum, the level of colorant present, and the reaction time, is necessary in order to attempt to reproduce the same quality of cross-linked colored polyester composition. Further, these colored polyester

compositions are brittle or low melting and may cause deterioration in physical properties of thermoplastic polymers when added in quantities sufficient to produce a high level of coloration.

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#### Practice of the Invention

This invention relates to a method for preparing a light absorbing linear polymeric having Formula I

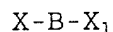


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I

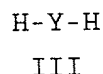
wherein A comprises the residue of a diacidic monomer comprising about 1 to 100 mole % of at least one light-absorbing monomer having a light absorption maximum between about 300 nm and about 1200 nm and wherein the remaining portion of A comprises the residue of a non-light absorbing monomer which does not absorb significantly at wavelengths above 300 nm or has a light absorption maximum below 300 nm and wherein B is a divalent organic radical selected from C<sub>2</sub>-C<sub>12</sub> alkylene, C<sub>3</sub>-C<sub>8</sub> cycloalkylene, C<sub>1</sub>-C<sub>4</sub> alkylene- C<sub>3</sub>-C<sub>8</sub>-cycloalkylene- C<sub>1</sub>-C<sub>4</sub> alkylene, C<sub>1</sub>-C<sub>4</sub> alkylene-arylene- C<sub>1</sub>-C<sub>4</sub> alkylene, C<sub>2</sub>-C<sub>4</sub> alkylene-O- C<sub>2</sub>-C<sub>4</sub> alkylene, and C<sub>2</sub>- C<sub>4</sub>-alkylene-L-arylene- C<sub>2</sub>-C<sub>4</sub> alkylene and C<sub>2</sub>-C<sub>4</sub> alkylene-(L- C<sub>2</sub>-C<sub>4</sub> alkylene)<sub>1-4</sub>, wherein L is a linking group selected from-O-, -S-, -SO<sub>2</sub>-, -NH-, -N(C<sub>1</sub>-C<sub>6</sub> alkyl)-, -N(aryl)-, -N(SO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl)-, -N(SO<sub>2</sub>aryl)-, -SO<sub>2</sub>N(C<sub>1</sub>-C<sub>6</sub> alkyl)- and combinations thereof; wherein n is at least 2.

30 The process comprises reacting said diacidic monomer with an organic compound of Formula II



II

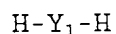
wherein B is as defined above and X and X<sub>1</sub> reactive groups and are independently selected from bromine, iodine and R-SO<sub>2</sub>O; wherein R is selected from C<sub>1</sub>-C<sub>6</sub> alkyl; C<sub>1</sub>-C<sub>6</sub> alkyl substituted with chlorine, fluorine, C<sub>1</sub>-C<sub>6</sub> alkoxy, aryl, aryloxy, arylthio or C<sub>3</sub>-C<sub>8</sub> cycloalkyl; C<sub>3</sub>-C<sub>8</sub> cycloalkyl or aryl, with said reaction being carried out in a solvent in the presence of a base; wherein the useful diacid light-absorbing monomers have Formula III



wherein H represents an acidic hydrogen atom; Y is a divalent light-absorbing moiety selected from a variety of chromophoric classes including azo, disazo, bis-azo, methine, arylidene, polymethine, azo-methine, azamethine, anthraquinone, anthrapyridone (3H-dibenz[f,ij]isoquinoline-2,7-dione, nitroarylamines anthrapyridine (7H-dibenz[f,ij]isoquinoline-7-one, phthaloylphenothiazine (14H-naphth[2,3-a]phenothiazine-8,13-dione, benzanthrone (7H (de) anthracene-7-one), anthrapyrimidine (7H-benzo[e]perimidine-7-one), anthrapyrazole, anthraisothiazole, triphenodioxazine, thiaxanthene-9-one, fluorindine (5,12-dihydroquinoxaline[2,3-b]phenazine, quinophthalone, phthalocyanine, metal phthalocyanine, naphthalocyanine, metal naphthalocyanine, nickel dithiolenes, squarylium compounds, croconium compounds, coumarin (2H-1-benzopyran-2-one), coumarin imine (2H-1-benzopyran-2-imine), perinone, benzodifuran, phthaloylacridone, phthaloylphenoxazine (14H-naphtho[2,3-a]phenoxazine-8,13-dione, phthaloylacridone (13H-naphtho[2,3-c]acridine-5,8,14-trione), anthraquinonethioxanthane (8H-naphtho[2,3-c]thioxanthene-5,8,13-trione, anthrapyridazone, pyrrolo[3,4-c]pyrrole, indigo, thioindigo, quinoline,

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xanthene, acridine, azine, cyanine, oxazine, 1,4 and 1,5-naphthoquinones, 2,5-diarylamino-terephthalic acids and esters, pyromellitic acid diimide, naphthalene-1,4,5,8-tetracarboxylic acid diimide, 3,4,9,10-perylene-tetracarboxylic acid diimide, 3-aryl-2,5-dioxypyrroline, 3-aryl-5-dicyanomethylene-2-oxopyrroline, arylisoindoline, hydroxybenzophenone, benotriazole, naphthotriazole, diminoisoindoline, naphthopyran (3H-naphtho[2,1-6]pyran-3-one and 3-imine, phthalimides, 2-arylbenzazoles, carbostyryls, 1,2-diarylethenes, 2,5-diarylthiophenes, 2,5-diaryl-1,3,4-oxadiazoles, triazines, 2,5-diarylfurans, 2,5-diaryl-1,3,4-thiadiazoles, thiophenes, 1,3-diphenyl-2-pyrazolines, 2-arylbenzofurans, 2,6-diphenylbenzofurans, quinolines, quinoxalines, 3,4-diarylfuranones, distyrylarenes, benzanthrones, polyarenes and naphthalimides; wherein the hydrogen atoms of Formula III are independently bonded to an oxygen, sulfur, or nitrogen atom which is a part of the light absorbing moiety Y; wherein the useful non light-absorbing monomers have Formula IV,



IV

wherein H represents an acidic hydrogen atom;  $\text{Y}_1$  is a divalent moiety, selected from  $-\text{O}_2\text{C}-\text{R}_1-\text{CO}_2-$  and  $-\text{O}-\text{R}_2-\text{O}-$  and  $-\text{O}_2\text{C}-\text{R}_3-\text{O}-$ , wherein  $\text{R}_1$  is selected from  $\text{C}_2-\text{C}_{12}$  alkylene, 1-4-cyclohexylene, arylene, arylene-O-arylene, arylene- $\text{SO}_2$ -arylene, arylene-S-arylene, and  $\text{C}_1-\text{C}_4$  alkylene-O- $\text{C}_1-\text{C}_4$  alkylene; wherein  $\text{R}_2$  is selected from arylene, arylene-O-arylene, arylene-S-arylene, arylene- $\text{SO}_2$ -arylene, phenylene-phenylene, and phenylene- $\text{C}(\text{R}_4)_2$ -phenylene; wherein  $\text{R}_4$  is selected from hydrogen and  $\text{C}_1-\text{C}_4$  alkyl; wherein  $\text{R}_3$  is selected from arylene.

In diacid light absorbing monomers having Formula III, the hydrogen atoms are preferably attached to an

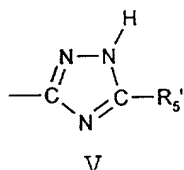
oxygen, a sulfur or a nitrogen atom which in combination provides two acidic functional group which can produce the corresponding anions under basic conditions by removal of the protons. The acidic functional groups usually have an acid dissociation constant of about  $10^{-1.5}$  to about  $10^{-12}$  (pK<sub>a</sub> of from about 1.5 to about 12). In the case of nitrogen, both protons may be attached to a single nitrogen which is attached to a sulfonyl moiety thus providing two acidic hydrogens on a single functional group.

Typical, acidic groups which provide one acidic hydrogen include-CO<sub>2</sub>H, -SH, -OH attached to an aromatic ring, -CONHCO-, -SO<sub>2</sub>-NH-CO-, -SO<sub>2</sub>-NH-SO<sub>2</sub>-, 1(H)-1,2,4-triazol-3-yl-, imidazolyl, benzimidazolyl, pyrazolyl, -SO<sub>2</sub>H attached to aromatic ring, -NHSO<sub>2</sub>R<sub>5</sub> and-SO<sub>2</sub>NHR<sub>5</sub>, wherein R<sub>5</sub> is selected from C<sub>1</sub>-C<sub>6</sub> alkyl; C<sub>1</sub>-C<sub>6</sub> alkyl substituted with at least one group selected from C<sub>1</sub>-C<sub>6</sub> alkoxy, aryl, aryloxy, arylthio or C<sub>3</sub>-C<sub>8</sub> cycloalkyl; C<sub>3</sub>-C<sub>8</sub> cycloalkyl; aryl.

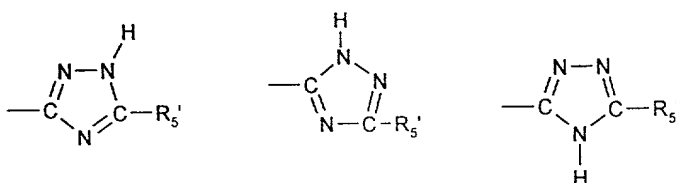
An example of an acidic functional group providing two acidic hydrogen attached to nitrogen is the sulfamoyl group (-SO<sub>2</sub>NH<sub>2</sub>).

The preferred method for producing light absorbing polymeric compositions utilizes the monomers of Formula III, wherein the protons are a part of the-CO<sub>2</sub>H, OH attached to aromatic ring, -CO-NH-CO- or 1(H)-1,2,4-triazol-3-yl functional groups. The carboxy groups are normally attached to an aromatic ring carbon or aliphatic carbon which is a part of Y. The hydroxy groups are normally attached to an unsubstituted or substituted phenyl or naphthyl radical which is a part of Y. The -CO-NHCO- groups are usually attached to an aromatic ring to provide an imide such as phthalimide or naphthalimide which are a part of Y. The 1(H)-1,2,4-triazol-3-yl group has the following Formula V, wherein R<sub>5</sub>' is



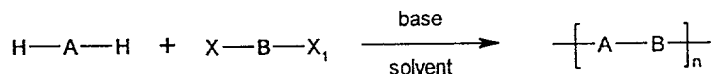


selected from hydrogen, C<sub>1</sub>-C<sub>6</sub> alkyl or aryl. It should be  
 5 observed that the triazole may exist in isomeric form as follows:



10 The 1(H)-1,2,4-triazol-3-yl group is preferably attached to a sulfur atom which is attached to the remainder of Y.

The method of the invention in the broadest sense involves the preparation of light absorbing polymeric  
 15 compositions by reacting a diacidic monomer comprising at least 1 mole % of at least one diacidic light absorbing monomer represented by H-A-H with an organic compound containing two reactive groups represented by X-B-X<sub>1</sub>, where B, X and X<sub>1</sub> are as defined above. Thus, the method  
 20 may be summarized as:



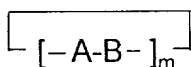
The diacidic monomer H-A-H must be acidic enough to form two nucleophiles in the presence of base under  
 25 convenient reaction conditions for the most advantageous process. This usually requires that diacidic monomers have pK<sub>a</sub> values of about 12 or below.

The dinucleophilic monomer, formed by the removal of the two hydrogen atoms by the base, attacks the

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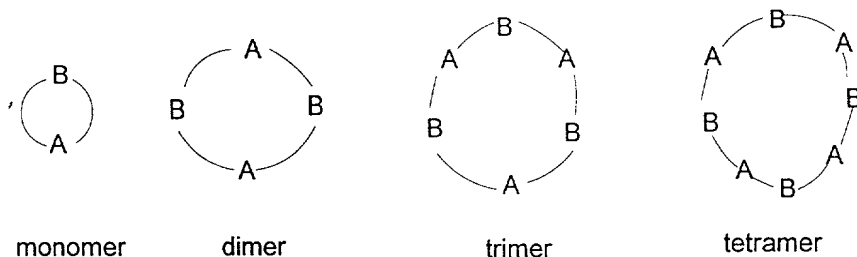
electrophilic compound II, thus displacing anions  $X^-$  and  $X_1^-$ , with head-to-tail combination with covalent bonding to produce a linear polymer  $\{-A-B\}_n$ , wherein n represents the number of repeating units. The number of repeating units must be at least 2, but usually ranges between about 2 and about 25, with the preferred number being between about 3 and about 15.

The composition produced by the method of the invention comprises, as stated above, a polymer having the general formula  $\{-A-B\}_n$ . The composition also comprises one or more cyclic compounds having the general formula



I-A

wherein m may be 1, 2, 3, or 4, e.g., the cyclic compounds having the general structures:



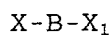
The number and concentrations of the cyclic compounds is dependent upon a variety of factors such as the structure of diacid  $H-A-H$ , the structure of the organic compound  $X-B-X_1$ , and the conditions used to facilitate the reaction to produce the composition. The cyclic compounds may constitute up to about 35 weight percent, typically about 0.5 up to 30 weight percent, of the total weight of the composition produced by the method of the invention.

Suitable bases include alkali metal carbonates; alkali metal bicarbonates; tertiary amines such as triethylamine, tri-n-butylamine, N-methylpiperidine, N,N'-dimethylpiperazine, N-methylmorpholine, N,N,N',N'-tetramethylenediamine, etc.; aromatic nitrogen bases such as pyridines, picolines, quinolines, isoquinolines, N-alkylpyrroles, N-alkylimidazoles, etc.; bicyclic nitrogen containing bases having non-hindered electron pairs, such as 1,8-diazabicyclo[4,3,0]undec-7-ene (DBU), 1,5-diazabicyclo[4,3,0]non-5-ene (DBN) and 1,4-diazadicyclo[2,2,2]octane (DABCO®).

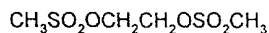
Typical solvents useful in the polymerization reaction include aprotic polar solvents such as N,N-dimethylacetamide, N,N-dimethylformamide, N-methyl-2-pyrrolidone, N-methyl-N-phenyl formamide, dimethyl sulfoxide, aliphatic nitriles, sulfolane, hexamethyl phosphoramide, etc. and mixtures thereof. Water, alcohols, ketones pyridine and ether-alcohols, such as the Cellosolves, also are sometimes useful. One requirement is that the solvent not form a stronger nucleophile in the presence of the base than that obtained from the diacidic monomer H-A-H.

The new improved process of the invention allows the preparation of near ultraviolet (UV-A, UV-B and UV-C), visible and near infrared light absorbing linear polymeric compositions at relatively low temperatures, usually at from about 75°C to about 125°C, without prolonged heating times. Furthermore, the method is adaptable to batch-process production which is advantageous for expensive products such as colorants, near infrared absorbers and near ultraviolet absorbers. The method is adaptable to a wide range of chromophoric classes since the polymer preparative reaction is carried out at relatively low temperature, which for example, allows colored polymeric compositions to be readily prepared from the very important azo class of colorants.

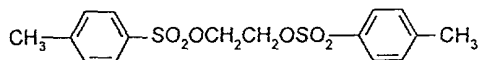
The preferred reactants of Formula II



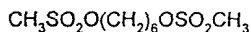
- 5 are the disulfonate compounds where X and X<sub>1</sub> are both a sulfonate ester of the formula-OSO<sub>2</sub>R, wherein R is selected from C<sub>1</sub>-C<sub>4</sub> alkyl, phenyl or p-methylphenyl and wherein B is selected from C<sub>2</sub>-C<sub>6</sub> alkylene, -CH<sub>2</sub>-1,4-cyclohexylene-CH<sub>2</sub>-, 2,2,4,4-tetramethyl-1,3-cyclobutylene, 10 1,4-cyclohexylene, -CH<sub>2</sub>CH<sub>2</sub>(OCH<sub>2</sub>CH<sub>2</sub>)<sub>2-3</sub> and -CH<sub>2</sub>CH<sub>2</sub>O-1,4-phenylene-O-CH<sub>2</sub>CH<sub>2</sub>-. Particularly, preferred reactants of Formula II are those where B is selected from-CH<sub>2</sub>CH<sub>2</sub>-, -CH<sub>2</sub>CH(CH<sub>3</sub>)CH<sub>2</sub>-, -CH<sub>2</sub>C(CH<sub>3</sub>)<sub>2</sub>CH<sub>2</sub>-, -(CH<sub>2</sub>)<sub>4</sub>-, -(CH<sub>2</sub>)<sub>6</sub>-, -CH<sub>2</sub>CH<sub>2</sub>(OCH<sub>2</sub>CH<sub>2</sub>)<sub>1-4</sub> and-CH<sub>2</sub>-1,4-cyclohexylene-CH<sub>2</sub>-.  
15 Typical reactants of Formula II are as follows:



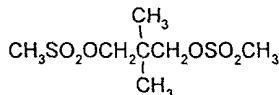
1,2-Ethandiol,  
dimethanesulfonate



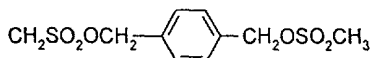
1,2-Ethandiol,  
bis(4-methylbenzenesulfonate)



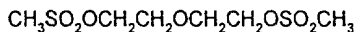
1,6-Hexanediol,  
dimethanesulfonate



1,3-Propanediol,2,2-dimethyl-,  
dimethanesulfonate

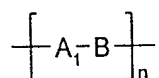


1,4-Cyclohexanedimethanol,  
dimethanesulfonate



Ethanol,2,2'-oxybis-,  
dimethanesulfonate

The invention also relates to a light absorbing linear polymeric composition having Formula Ia:



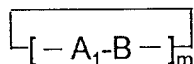
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Ia

wherein A<sub>1</sub> comprises the residue of at least one diacidic monomer having a light absorption maximum between about 300 nm and about 1200 nm, preferably between about 325 nm and 1100 nm and most preferably between about 350 nm and 1000 nm and wherein B is defined above and which has been prepared by reacting a diacid light-absorbing monomer of Formula III (H-Y-H) as defined above with an organic compound having Formula II (X-B-X<sub>1</sub>) as defined above, with the polymer producing reaction having been carried out in a solvent in the presence of base. The above-described light absorbing composition of formula Ia also contains or comprises one or more cyclic compounds having the formula

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I-B

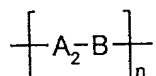
wherein A<sub>1</sub> and B are defined above and m may be 1, 2, 3, or 4. As stated hereinabove, the number and concentrations of the cyclic compounds is dependent upon a variety of factors such as the structure of diacid H-A-H, the structure of the organic compound X-B-X<sub>1</sub>, and the conditions used to facilitate the reaction to produce the composition. The cyclic compounds of formula I-B may constitute up to about 35 weight percent, typically about 1 up 30 weight percent, of the total weight of the above-described light absorbing composition.

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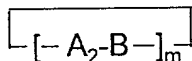
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The invention also relates to a light absorbing linear polymeric composition having Formula Ib



Ib

wherein A<sub>2</sub> comprises the residue of at least one diacidic monomer, having a light absorption maximum between about 300 nm and about 1200 nm, preferably between about 325 nm and 1100 nm and most preferably between about 350 nm and 1000 nm and which comprises at least about 50% by weight of the total of the composition of Formula Ib and wherein the remainder of A<sub>2</sub> comprises the residue of at least one non-light absorbing monomer of Formula IV above, and wherein said polymeric composition has been prepared by reacting diacidic monomers of Formula III and Formula IV with an organic compound having Formula II above, with the polymer producing reaction having been carried out in a solvent in the presence of base. The light absorbing composition of formula Ib also contains or comprises one or more cyclic compounds having the formula



I-C

wherein A<sub>2</sub> and B are defined above and m may be 1, 2, 3, or 4. Again, the number and concentrations of the cyclic compounds is dependent upon a variety of factors such as the structure of diacid H-A-H, the structure of the organic compound X-B-X<sub>1</sub>, and the conditions used to facilitate the reaction to produce the composition. The cyclic compounds of formula I-B may constitute up to about 35 weight percent, typically about 1 up 30 weight percent,

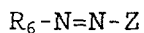
of the total weight of the above-described light absorbing composition.

The polymer compositions of Formula I, Ia, and Ib and the cyclic compositions of formulas I-A, I-B and I-C are referred to as "polydyes" herein when they absorb visible light thus rendering them strongly colored.

The invention further relates to a thermoplastic polymeric composition which comprises a thermoplastic polymer blended with at least one light absorbing linear polymeric composition of Formula I, Ia or Ib above which, as noted above, contain or comprise one or more cyclic compounds having the general formula I-A. The thermoplastic polymeric composition is usually selected from polyesters, polyolefins, polyamides, polyimides, polyvinyl chloride, polyurethanes, polycarbonates, cellulose esters, polyacrylates, polyvinylesters, polyester-amides, polystyrene, polyacrylonitrile-butadiene- styrene and polystyrene-acrylonitrile. The preferred thermoplastic polymeric composition comprises the light-absorbing polymeric compositions of Formula Ia.

The invention also relates to some of the diacidic light absorbing monomers used to prepare the light absorbing polymeric composition of Formula I, Ia, or Ib.

Preferred azo compounds useful in the practice of the invention correspond to Formula VI



VI

wherein  $R_6$  is the residue of an aromatic or heteroaromatic amine which has been diazotized and coupled with a coupling component H-Z and is preferably derived from the aromatic and heteroaromatic amine classes of aniline, 1-aminonaphthalene, 1-aminoanthraquinone, 4-aminoazobenzene, 2-aminothiazole, 2-aminobenzothiazole, 3-amino-2,1-benzisothiazole, 2-aminothieno[2,3-d]thiazole,

4-aminopyrazoloisothiazole, 2-amino-1,3,4-thiadiazolē, 5-amino-1,2,4-thiadiazole, 5-amino-1,2,3-triazole, 2-amino-1,3,4-triazole, 2(5) aminoimidazole, 3-aminopyridine, 2(3

5 aminothiophene, 2(3) aminobenzo[b]thiophene, 2-  
aminothieno[3,2-b]thiophene, 3-aminothieno[2,3-  
c]isothiazole, 3-amino-7-benz- 2,1-isothiazole, 3-  
aminobenzothienoisothiazole, 3-aminoisothiazole[3,4-  
d]pyrimidine, 5-amino- 1,2,3-triazole, 3(4)  
10 aminophthalimide and 5(6) amino-1,2-benzisothiazolon-1,1-  
dioxide with said aromatic and heteroaromatic ring systems  
being unsubstituted or substituted with one or more groups  
selected from C<sub>1</sub>-C<sub>10</sub> alkyl, C<sub>1</sub>-C<sub>6</sub> alkoxy, C<sub>3</sub>-C<sub>8</sub> cycloalkyl,  
carboxy, halogen, C<sub>1</sub>-C<sub>6</sub> alkoxycarbonyl, formyl, C<sub>1</sub>-C<sub>6</sub>  
15 alkanoyl, C<sub>1</sub>-C<sub>6</sub> alkanoyloxy, dicyanovinyl, C<sub>3</sub>-  
C<sub>8</sub>-cycloalkanoyl, thiocyano, trifluoroacetyl, cyano,  
carbamoyl, -CONH C<sub>1</sub>-C<sub>6</sub> alkyl, CONHaryl, CON(C<sub>1</sub>-C<sub>6</sub> alkyl)<sub>2</sub>,  
sulfamoyl, SO<sub>2</sub>NH C<sub>1</sub>-C<sub>6</sub> alkyl, SO<sub>2</sub>N(C<sub>1</sub>-C<sub>6</sub> alkyl)<sub>2</sub>, SO<sub>2</sub>NHaryl,  
SO<sub>2</sub>NH C<sub>3</sub>-C<sub>8</sub> cycloalkyl, CONH C<sub>3</sub>-C<sub>8</sub> cycloalkyl, aryl, aroyl,  
20 -NHSO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, -N(C<sub>1</sub>-C<sub>6</sub> alkyl)SO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, -NHSO<sub>2</sub>  
aryl, NHCO C<sub>1</sub>-C<sub>6</sub> alkyl, NHCO C<sub>3</sub>-C<sub>8</sub> cycloalkyl, NHCOaryl,  
NHCO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, NHCONH C<sub>1</sub>-C<sub>6</sub> alkyl, NHCONHaryl, N(C<sub>1</sub>-C<sub>6</sub>  
alkyl)aryl, arylazo, heteroaryl, aryloxy, arylthio, C<sub>3</sub>-C<sub>8</sub>  
cycloalkoxy, heteroarylazo, heteroarylthio, arylsulfonyl,  
25 tricyanovinyl, aryloxysulfonyl, C<sub>1</sub>-C<sub>6</sub> alkylsulfonyl,  
trifluoromethyl, fluorosulfonyl, trifluoromethylsulfonyl,  
thiocyano, hydroxy, nitro or CH=D, wherein D is the  
residue of an active methylene compound as defined below.

Z is the residue of an electron rich coupling component selected from the classes of anilines, 1-aminonaphthalenes, 1,2-dihydroquinolines, 1,2,3,4-tetrahydroquinolines, benzmorpholines (3,4-dihydro-2H-1,4-benzoxazine), pyrazolones, pyrazoles, 3-cyano-6-hydroxy-2-pyridones, 2,3-dihydroindoles, indoles, 4-hydroxycoumarins, 4-hydroxy-2-quinolones, imidazo[2,1-b]thiazoles, julolidines (2,3,6,7-tetrahydro-1H,5H-

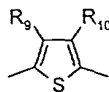
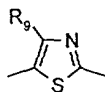
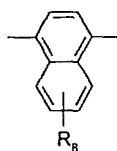
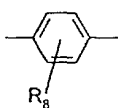


benzo[ij]quinolizines), 1-oxajulolidines, 1,2,5,6-tetrahydro-4H-pyrrolo[3,2,1-ij]quinolines, 2,6-diamino-3-cyanopyridines, 2-aminothiazoles, 2-aminothiophenes, 5,5-dimethyl-1,3-cyclohexanedione (dimedone), phenols, naphthols, 2,4-pentanediones or acetoacetarilides; with the proviso that the compounds of Formula VI contain two acidic functional groups containing one acidic hydrogen each or contain one sulfamoyl group (-SO<sub>2</sub>NH<sub>2</sub>) which contains two acidic hydrogens.

Preferred disazo compounds correspond to Formula VII

$$R_6-N=N-R_7N=N-Z \quad \text{VII}$$

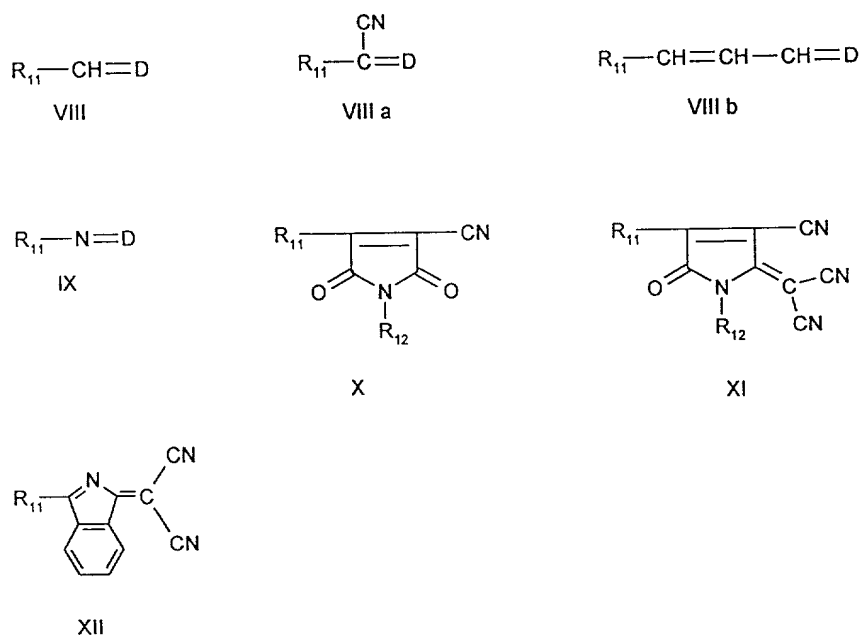
wherein R<sub>6</sub> and Z are as defined above and R<sub>7</sub> is a divalent aromatic or heteroaromatic radical selected from the classes 1,4-phenylene, naphthalene-1,4-diyl, thiazol-2,5-diyl and thiophene-2,5-diyl:



wherein R<sub>8</sub> is selected from hydrogen or 1-2 groups selected from C<sub>1</sub>-C<sub>6</sub> alkyl, C<sub>1</sub>-C<sub>6</sub> alkoxy, cyano, halogen, -NHCO C<sub>1</sub>-C<sub>6</sub> alkyl, -NHCO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, -NHCO aryl, -NHCONH aryl or NHCONH C<sub>1</sub>-C<sub>6</sub> alkyl; R<sub>9</sub> is selected from hydrogen, C<sub>1</sub>-C<sub>6</sub> alkyl, halogen, aryl, heteroaryl; R<sub>10</sub> is selected from hydrogen, C<sub>1</sub>-C<sub>6</sub> alkoxycarbonyl, cyano, carbamoyl, aryl, arylsulfonyl, aroyl, -CONH C<sub>1</sub>-C<sub>6</sub> alkyl, or C<sub>1</sub>-C<sub>6</sub> alkylsulfonyl; with the provision that two acidic functional groups containing one acidic hydrogen each or

one functional group containing two acidic hydrogens are present on compounds of Formula VII.

The preferred methine, arylidene, polymethine, azamethine, 3-aryl-2,5-dioxypyrroline, 3-aryl-5-dicyanomethylene-2-oxopyrroline and aryl isoindoline compounds correspond to Formula VIII, VIIIa, VIIIb, IX, X, XI and XII, respectively:



10

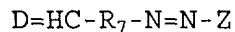
wherein R<sub>11</sub> is the residue of an aniline, 1-naphthylamine, 1,2-dihydroquinoline, 1,2,3,4-tetrahydroquinoline, 1,3,3-trimethyl-2-methyleneindole, 1,3-dihydro-2-methylene-1,1,3-trimethyl-2H-benz[e]indole, imidazo [2,1-b]thiazole, benzomorpholine (3,4-dihydro-2H-1,4,benzoxazine), indole, 2,3-dihydroindole, 2-aminothiazole, julolidine (2,3,6,7-tetrahydro-1H, 5H- benz [ij] quinolizine, 1-oxajulolidine, 4H-pyrrolo [3,2,1-ij]-quinoline, phenol, naphthol, thiophenol, pyrrole, pyrazole, furan, thiophene, carbazole, phenothiazine or phenoxazine compound; R<sub>12</sub> is selected from hydrogen, C<sub>1</sub>-C<sub>10</sub> alkyl, C<sub>3</sub>-C<sub>8</sub> alkenyl, C<sub>3</sub>-C<sub>8</sub>-alkynyl, C<sub>3</sub>-C<sub>8</sub> cycloalkyl,

15

20

aryl,  $(\text{CH}_2\text{CH}_2\text{O})_{1-3}$   $\text{R}_{13}$  and  $\text{C}_1\text{-C}_4$  alkylene-  $\text{C}_3\text{-C}_8$   
cycloalkylene, wherein the  $\text{C}_1\text{-C}_6$  alkyl groups may be  
substituted by at least one group selected from carboxy,  
 $\text{C}_1\text{-C}_6$  carbalkoxy,  $\text{C}_1\text{-C}_6$  alkanoyloxy, cyano, hydroxy,  
5 chlorine, fluorine,  $\text{C}_1\text{-C}_6$  alkoxy,  $\text{C}_3\text{-C}_8$  cycloalkyl or aryl;  
 $\text{R}_{13}$  is selected from hydrogen,  $\text{C}_1\text{-C}_6$  alkoxy or  $\text{C}_1\text{-C}_6$   
alkanoyloxy; wherein D is the residue of an active  
methylene compound selected from malononitrile, -  
cyanoacetic acid esters, malonic acid esters, -cyanacetic  
10 acid amides,  $-\text{C}_1\text{-C}_6$  alkylsulfonylacetonitriles, -  
arylsulfonylacetonitriles,  $-\text{C}_1\text{-C}_6$  alkanoylacetonitriles,  
-aroylacetonitriles, -heteroarylacetonitriles,  
bis(heteroaryl)methanes, 1,3-indanediones, 2-furanones,  
benzo-2-furanones, naphtho-2-furanones, 2-indolones, 3-  
15 cyano-1,6-dihydro-4-methyl-2,6-dioxy (2H)-pyridines, benzo  
(b) thieno-3-ylidene propane dinitrile-5,5-dioxides, 1,3-  
bis (dicyanomethylene) indanes, barbituric acid, 5-  
pyrazolones, dimedone, 3-oxo-2,3-dihydro-1-benzothiophene-  
1,1-dioxides or  $\text{aryl-C}(\text{CH}_3)\text{C}=\text{C}(\text{CN})_2$ , with the proviso that  
20 two acidic functional groups containing one acidic  
hydrogen each, or a functional group containing two acidic  
hydrogens are present in compounds of Formula VIII, VIIIA,  
VIIIB, IX, X, XI, and XII.

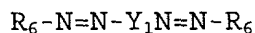
Preferred azo-methine compounds correspond to Formula  
25 XIII



XIII

30 wherein D,  $\text{R}_7$  and Z are as defined previously.

The bis-azo compound corresponds to Formula VIIa



VIIa

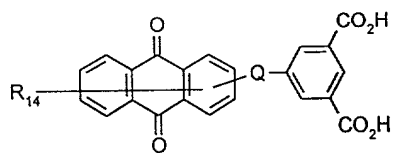
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wherein  $R_6$  is as defined above and  $Y_1$  is the residue of a bis coupling component selected from the classes of anilines, 1,2-dihydroquinolines, 1,2,3,4-tetrahydroquinolines, benzomorpholines (3,4-dihydro-  
10 2H-1,4-benzoxazines), 3-cyano-6-hydroxy-2-pyridones, 2,6-diaminopyridines, 2,3-dihydroindoles, naphthylamines, 2-aminothiazoles, or a combination of these; with the provision the compounds of Formula VIIa contain two acidic functional groups containing one acidic hydrogen each or  
15 contain one sulfamoyl group ( $-SO_2NH_2$ ) which contains two acidic hydrogens.

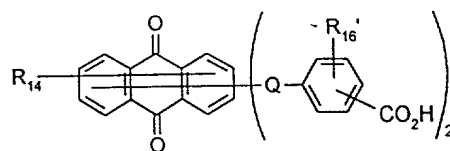
Several diacid monomers which are described in U.S. Patent Nos. 4,804,719 and 3,689,501 are useful in the practice of the invention, including various  
20 anthraquinones, anthrapyridones, anthraisothiazoles, anthraprimidines, anthrapyrimidones, phthaloylacridones, etc.

Some of the preferred anthraquinone, anthrapyridone and anthrapyrimidine compounds correspond to the light  
25 absorbing compounds of Formulae XIV- XIXf

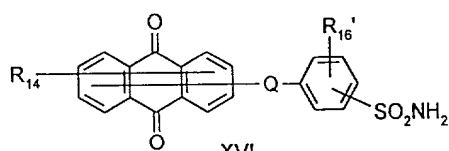
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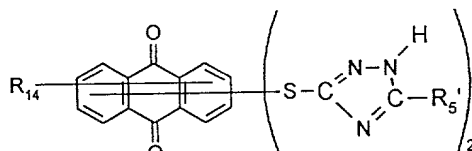
XIV



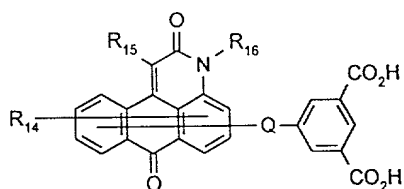
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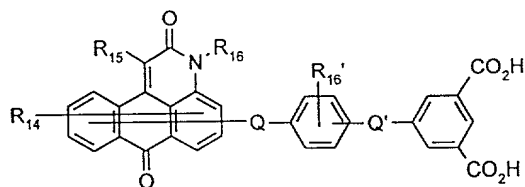
XVI



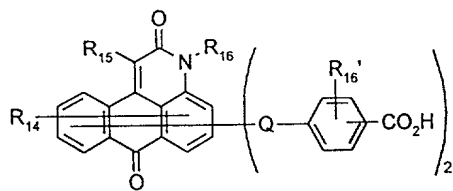
XVII



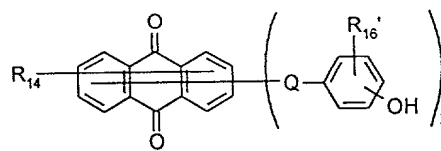
XVIII



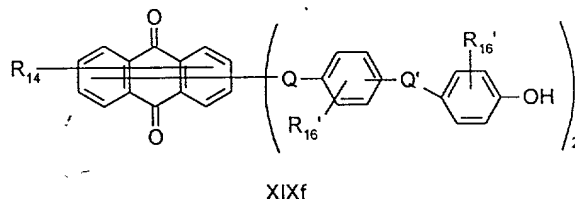
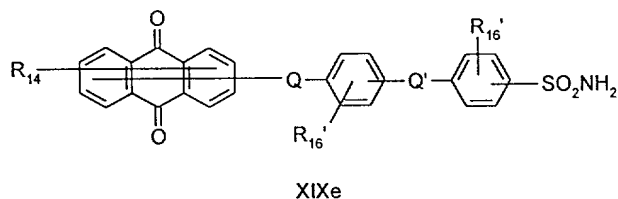
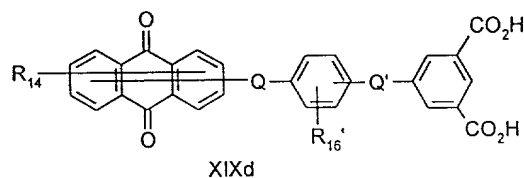
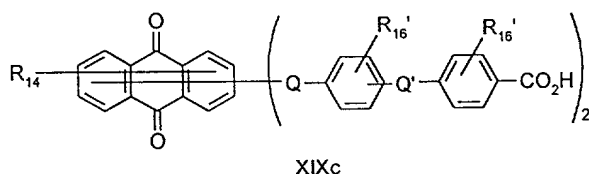
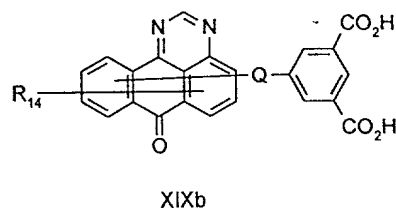
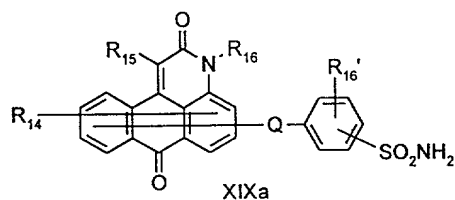
XVIII a



XVIII b

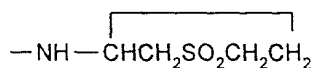


XIX



- wherein  $R_{14}$  is selected from the group consisting of hydrogen, 1-4 groups selected from amino,  $C_1$ - $C_{10}$  alkylamino,  $C_3$ - $C_8$  alkenylamino,  $C_3$ - $C_8$  alkynylamino,  $C_3$ - $C_8$  cycloalkylamino, arylamino, halogen,  $C_1$ - $C_6$  alkoxy,  $C_1$ - $C_6$  alkylthio, aryl, aroyl,  $C_1$ - $C_6$  alkanoyl,  $C_1$ - $C_6$  alkanoyloxy, NHCO  $C_1$ - $C_6$  alkyl, NHCOaryl, NHCO $_2$   $C_1$ - $C_6$  alkyl, NHSO $_2$   $C_1$ - $C_6$  alkyl, NHSO $_2$  aryl,  $C_1$ - $C_6$  alkoxycarbonyl, aryloxy, arylthio, heteroarylthio, cyano, nitro, trifluoromethyl, thiocyno, SO $_2$  $C_1$ - $C_6$  alkyl, SO $_2$  aryl, -SO $_2$ NH  $C_1$ - $C_6$  alkyl, -SO $_2$ N( $C_1$ - $C_6$  alkyl) $_2$ , -SO $_2$ N( $C_1$ - $C_6$  alkyl) aryl, CONH  $C_1$ - $C_6$  alkyl, CON( $C_1$ - $C_6$  alkyl) $_2$ , CON( $C_1$ - $C_6$  alkyl) aryl,  $C_1$ - $C_6$  alkyl, furfurylamino,

tetrahydrofurfurylamino, 4-(hydroxymethyl)  
cyclohexanemethylamino,



- 5 or hydroxy; Q and Q' are independently selected from-O-,  
-N(COR<sub>10</sub>)-, -N(SO<sub>2</sub>R<sub>10</sub>)-, -N(R<sub>10</sub>)-, -S-, -SO<sub>2</sub>-, -CO<sub>2</sub>-,  
-CON(R<sub>10</sub>)-, SO<sub>2</sub>N (R<sub>10</sub>)-, wherein R<sub>10</sub> is selected from  
hydrogen, aryl, C<sub>3</sub>-C<sub>8</sub> cycloalkyl, or C<sub>1</sub>-C<sub>10</sub> alkyl; R<sub>15</sub> is  
10 selected from hydrogen, cyano, C<sub>1</sub>-C<sub>6</sub> alkylamino, C<sub>1</sub>-C<sub>6</sub>  
alkoxy, halogen, arylthio, aryl, heteroaryl,  
heteroarylthio, C<sub>1</sub>-C<sub>6</sub> alkoxy carbonyl, aroyl or  
arylsulfonyl; R<sub>16</sub> is selected from hydrogen, C<sub>1</sub>-C<sub>10</sub> alkyl,  
C<sub>3</sub>-C<sub>8</sub> cycloalkyl and aryl; R<sub>16</sub>' is selected from the group  
15 consisting of hydrogen, one or two groups selected from  
C<sub>1</sub>-C<sub>6</sub> alkyl, halogen and C<sub>1</sub>-C<sub>6</sub> alkoxy; wherein each C<sub>1</sub>-C<sub>6</sub>  
alkyl group and C<sub>1</sub>-C<sub>6</sub> alkyl group which is a portion of  
another group may contain at least one substituent  
selected from hydroxy, cyano, chlorine, fluorine, C<sub>1</sub>-C<sub>6</sub>  
20 alkoxy, C<sub>3</sub>-C<sub>8</sub> cycloalkoxy, C<sub>1</sub>-C<sub>6</sub> alkylcyclohexyl,  
hydroxymethyl cyclohexyl, aryl and heteroaryl; with the  
provision that two acidic groups containing one acidic  
proton each or one acidic group containing two acidic  
hydrogens be present in the compounds of Formula XIV-  
25 XIXf.

Typical coupler residues which are represented by Z  
above in Formulae VI, VII, XIII for the classes of azo,  
disazo and azo-methine compounds, respectively include:

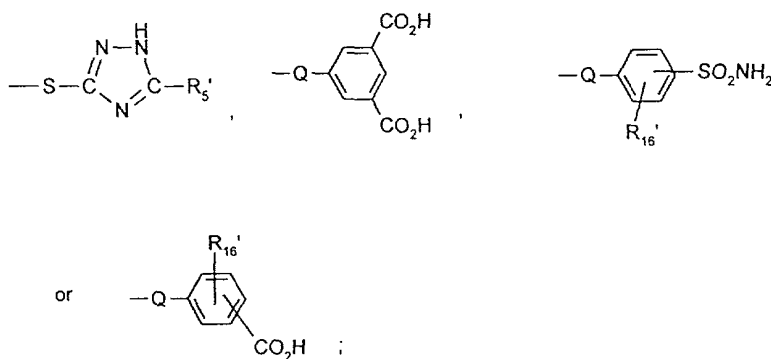
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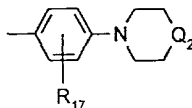
wherein R<sub>17</sub> is selected from the group consisting of hydrogen, 1-2 groups selected from C<sub>1</sub>-C<sub>6</sub> alkyl, C<sub>1</sub>-C<sub>6</sub>



- alkoxy, C<sub>1</sub>-C<sub>6</sub> alkylthio, -O-C<sub>2</sub>-C<sub>6</sub> alkylene-OH, O-C<sub>2</sub>-C<sub>6</sub> alkylene-C<sub>1</sub>-C<sub>6</sub> alkanoyloxy, C<sub>1</sub>-C<sub>6</sub> alkylene-OH, C<sub>1</sub>-C<sub>6</sub> alkylene-C<sub>1</sub>-C<sub>6</sub> alkanoyloxy, halogen, carboxy, C<sub>1</sub>-C<sub>6</sub> alkoxycarbonyl, trifluoromethyl, NHCOR<sub>24</sub>, NHCO<sub>2</sub>R<sub>24</sub>,  
 5 NHCON(R<sub>24</sub>)R<sub>25</sub>, and NHSO<sub>2</sub>R<sub>25</sub>, wherein R<sub>24</sub> is selected from hydrogen, C<sub>1</sub>-C<sub>10</sub> alkyl, C<sub>3</sub>-C<sub>8</sub> cycloalkyl or aryl, R<sub>25</sub> is selected from C<sub>1</sub>-C<sub>10</sub> alkyl, C<sub>3</sub>-C<sub>8</sub> cycloalkyl or aryl wherein each C<sub>1</sub>-C<sub>10</sub> alkyl group in R<sub>24</sub> and R<sub>25</sub> may be further substituted with one or more groups selected from C<sub>3</sub>-C<sub>8</sub> cycloalkyl, aryl, aryloxy, arylthio, CO<sub>2</sub>H, CO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, cyano, hydroxy, succinimido, C<sub>1</sub>-C<sub>6</sub> alkoxy,



- 15 wherein R<sub>5</sub>' , R<sub>16</sub>' and Q are as defined above; R<sub>18</sub> and R<sub>19</sub> are independently selected from hydrogen, unsubstituted C<sub>1</sub>-C<sub>10</sub> alkyl, substituted C<sub>1</sub>-C<sub>10</sub> alkyl, C<sub>3</sub>-C<sub>8</sub> cycloalkyl, C<sub>3</sub>-C<sub>8</sub> alkenyl, C<sub>3</sub>-C<sub>8</sub> alkynyl and aryl or R<sub>18</sub> and R<sub>19</sub> may be combined with another element to which they are attached  
 20 to form a radical Z having the formula

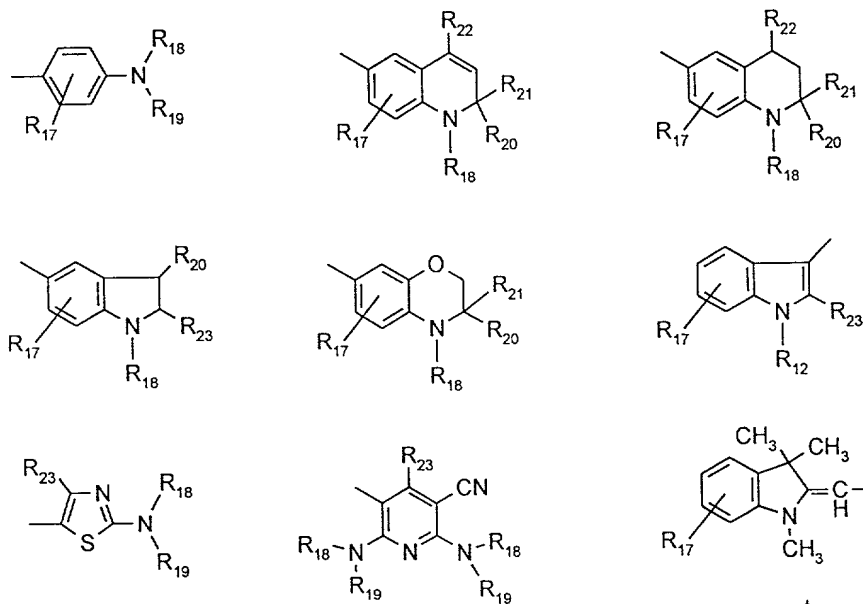


- wherein Q<sub>2</sub> is selected from a covalent bond, -O-, -S-,  
 25 -SO<sub>2</sub>-, -CO-, -CO<sub>2</sub>-, -N-(C<sub>1</sub>-C<sub>6</sub> alkyl)-, -N(CO C<sub>1</sub>-C<sub>6</sub> alkyl)-, -N(SO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl)-, -N(CO aryl)-, or -N(SO<sub>2</sub> aryl); R<sub>20</sub>, R<sub>21</sub>

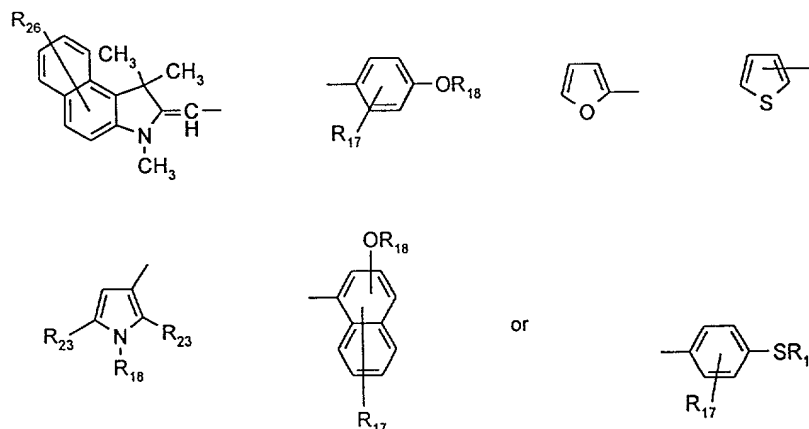
006237 994766 122900

and  $R_{22}$  are independently selected from the group consisting of or  $C_1$ - $C_6$  alkyl;  $R_{23}$  is selected from hydrogen,  $C_1$ - $C_6$  alkyl,  $C_3$ - $C_8$  cycloalkyl, heteroaryl or aryl.

- 5 Typical electron, rich aromatic residues which are represented by  $R_{11}$  in Formulae VIII- XII include:



10

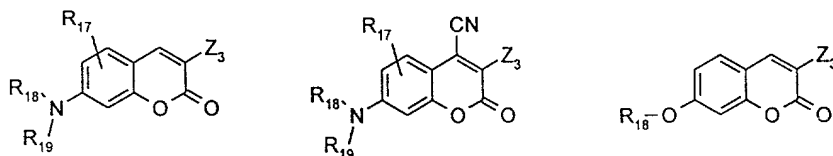


wherein  $R_{26}$  is selected from the group consisting of hydrogen, a group selected from  $C_1$ - $C_6$  alkoxy,  $CO_2H$ ,

C<sub>1</sub>-C<sub>6</sub> alkyl or C<sub>1</sub>-C<sub>6</sub> alkoxy; wherein R<sub>17</sub>-R<sub>23</sub> are as defined previously.

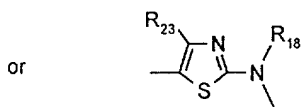
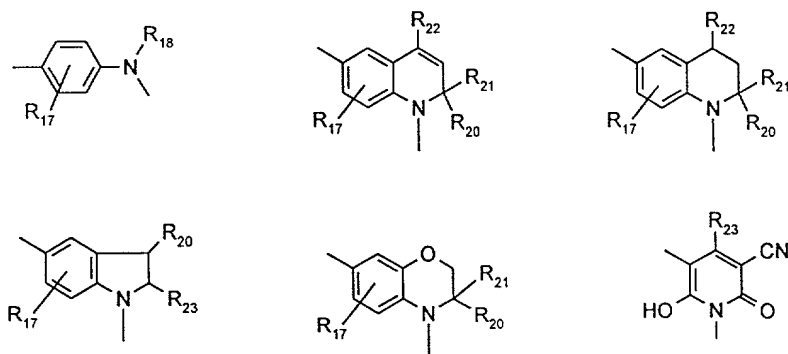
Preferred coumarin compounds useful in the practice of the invention correspond to the following formulae:

5



wherein Z<sub>3</sub> is selected from cyano, C<sub>1</sub>-C<sub>6</sub> alkoxycarbonyl, C<sub>1</sub>-C<sub>6</sub> alkylsulfonyl, arylsulfonyl, aryl, heteroaryl, formyl, aroyl, C<sub>1</sub>-C<sub>6</sub> alkanoyl or -CH=D, wherein D, R<sub>17</sub>, R<sub>18</sub> and R<sub>19</sub> are as defined previously with the provision that the coumarin compounds contain two acidic functional groups containing one acidic hydrogen each or contain one sulfamoyl (-SO<sub>2</sub>NH<sub>2</sub>) group which contains two acidic hydrogens.

Typical coupler residues which are represented by Y<sub>1</sub> in Formula VIIa above include those of the formula (Z<sub>1</sub>-L<sub>1</sub>-Z<sub>2</sub>) wherein Z<sub>1</sub> and Z<sub>2</sub> are independently selected from



wherein  $L_1$  is bonded to the nitrogen atom of  $Z_1$  and  $Z_2$ ;  
wherein  $L_1$  is selected from  $C_2$ - $C_{12}$  alkylene,  $C_3$ - $C_8$   
cycloalkylene, arylene,  $C_1$ - $C_4$  alkylene- $C_3$ - $C_8$  cycloalkylene-  
 $C_1$ - $C_4$  alkylene,  $C_1$ - $C_4$  alkylene-arylene- $C_1$ - $C_4$  alkylene,  $C_2$ - $C_4$   
5 alkylene-O-arylene-O- $C_2$ - $C_4$  alkylene, - $C_2$ - $C_4$  alkylene  $O_{1-3}$ -  
 $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$  alkylene-S- $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$   
alkylene-SO<sub>2</sub>- $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$  alkylene-N(SO<sub>2</sub>  $C_1$ - $C_6$   
alkyl)- $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$  alkylene-N(SO<sub>2</sub> aryl)- $C_2$ - $C_4$   
alkylene,  $C_2$ - $C_4$  alkylene-OCO<sub>2</sub>- $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$   
10 alkylene-O<sub>2</sub>C-arylene-CO<sub>2</sub>- $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$  alkylene-O<sub>2</sub>C-  
 $C_1$ - $C_{12}$  alkylene-CO<sub>2</sub>- $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$  alkylene-O<sub>2</sub>C- $C_3$ - $C_8$   
cycloalkylene-CO<sub>2</sub>- $C_2$ - $C_4$  alkylene,  $C_2$ - $C_4$  alkylene-NHCO- $C_2$ - $C_4$   
alkylene and  $C_2$ - $C_4$  alkylene-NHSO<sub>2</sub>- $C_2$ - $C_4$  alkylene; wherein  
 $R_{17}$ ,  $R_{18}$ ,  $R_{20}$ ,  $R_{21}$ ,  $R_{22}$ , and  $R_{23}$  are as defined previously.  
15 In the above definitions it is intended that in the  
terms  $C_1$ - $C_6$  alkyl,  $C_1$ - $C_6$  alkoxy,  $C_1$ - $C_6$  alkoxy-carbonyl,  $C_1$ - $C_6$   
alkylthio,  $C_1$ - $C_6$  alkylsulfonyl,  $C_1$ - $C_6$  alkanoyl, -CONH  $C_1$ - $C_6$   
alkyl, -SO<sub>2</sub>NH  $C_1$ - $C_6$  alkyl, -CON( $C_1$ - $C_6$  alkyl)<sub>2</sub>, -SO<sub>2</sub>N( $C_1$ - $C_6$   
alkyl)<sub>2</sub>, -NHSO<sub>2</sub>  $C_1$ - $C_6$  alkyl, -N( $C_1$ - $C_6$  alkyl) SO<sub>2</sub>  $C_1$ - $C_6$  alkyl,  
20 etc. unless otherwise stated that the  $C_1$ - $C_6$  alkyl portion  
of the group refers to a straight or branched chain alkyl  
group containing one to six carbon atoms and these  
substituted with one or more groups selected from carboxy,  
cyano, -SO<sub>2</sub>NH<sub>2</sub>, SO<sub>2</sub>NH  $C_1$ - $C_6$  alkyl, cyano, fluorine,  
25 chlorine,  $C_1$ - $C_6$  alkoxy, aryloxy, aryl, heteroaryl,  
arylthio, heteroarylthio,  $C_3$ - $C_8$ -cycloalkyl, -O<sub>2</sub>C  $C_1$ - $C_6$   
alkyl or -CO<sub>2</sub>  $C_1$ - $C_6$  alkyl.  
The terms  $C_1$ - $C_4$  alkylene,  $C_2$ - $C_4$  alkylene,  $C_1$ - $C_6$   
alkylene,  $C_2$ - $C_6$  alkylene, and  $C_2$ - $C_{12}$  alkylene are used to  
30 refer to divalent aliphatic hydrocarbon radicals  
containing one to four carbon atoms, two to four carbon  
atoms one to six carbon atoms, two to six carbon atoms, or  
two to twelve carbon atoms, respectively, and these  
optionally substituted with one or more groups selected  
35 from  $C_1$ - $C_6$  alkoxy, hydroxy, -O<sub>2</sub>C  $C_1$ - $C_6$  alkyl, carboxy, CO<sub>2</sub>  
 $C_1$ - $C_6$  alkyl, chlorine, fluorine, aryl or aryloxy.

The terms C<sub>3</sub>-C<sub>8</sub> cycloalkyl and C<sub>3</sub>-C<sub>8</sub> cycloalkylene are used to refer to fully saturated monovalent and divalent cycloaliphatic radicals, respectively, and these substituted by one or more C<sub>1</sub>-C<sub>6</sub> alkyl groups.

- 5       The terms C<sub>3</sub>-C<sub>8</sub> alkenyl and C<sub>3</sub>-C<sub>8</sub> alkynyl are used to refer to straight or branched hydrocarbon radicals containing at least one double bond or at least one triple bond, respectively.

- 10       In the terms aryl, NH aryl, aryloxy, aroyl, arylthio, arylsulfonyl, aryloxysulfonyl, -N(SO<sub>2</sub> aryl)-, -N(CO aryl)-, NHCO aryl, -NH CONH aryl, NHSO<sub>2</sub>, aryl, etc., the aryl portion of the group represents phenyl and naphthyl and these substituted with one or more groups selected from-CO<sub>2</sub>H, C<sub>1</sub>-C<sub>6</sub> alkyl, CO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, SO<sub>2</sub>NH<sub>2</sub>,  
15       SO<sub>2</sub>NH C<sub>1</sub>-C<sub>6</sub> alkyl, hydroxy, O C<sub>1</sub>-C<sub>6</sub> alkyl, S C<sub>1</sub>-C<sub>6</sub> alkyl, phenyl, O-arylene-CO<sub>2</sub>H, -S-arylene-CO<sub>2</sub>H, SO<sub>2</sub> arylene-CO<sub>2</sub>H, halogen, NHSO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, trifluoromethyl, NH CO C<sub>1</sub>-C<sub>6</sub> alkyl, cyano, or 1(H)-1,2,4-triazol-3-ylthio.

- 20       The term arylene is used to represent 1,2-, 1,3-, and 1,4- phenylene and these optionally substituted with one or more groups mentioned above as possible substituents on the aryl radical.

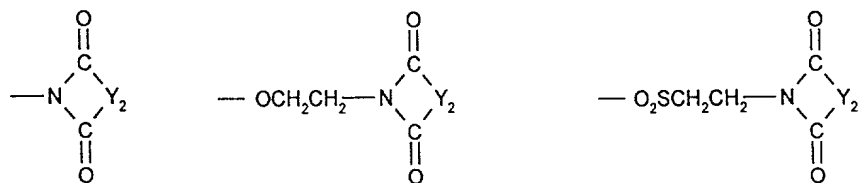
- 25       The term "heteroaryl" is used to describe a 5 or 6 membered heterocyclic aromatic ring containing one oxygen atom, and/or one sulfur atom, and/or up to three nitrogen atoms, said heterocyclic aryl ring optionally fused to one or two phenyl rings or another 5 or 6-membered heteroaryl ring. Examples of such ring systems include thienyl, furyl, pyrrolyl, imidazolyl, pyrazolyl, thiazolyl,  
30       isothiazolyl, oxazolyl, isoxazolyl, triazolyl, thiadiazolyl, oxadiazolyl, tetrazolyl, thiatriazolyl, oxatriazolyl, pyridyl, pyrimidyl, pyrazinyl, pyridazinyl, thiazinyl, oxazinyl, triazinyl, thiadiazinyl, oxadiazinyl, dithiazinyl, dioxazinyl, oxathiazinyl, tetrazynyl,  
35       thiatriazinyl, oxatriazinyl, dithiadiazinyl, imidazolynyl, dihydropyrimidyl, tetrahydropyrimidyl, tetrazolo [1,5-b]-

pyridazinyl and purinyl, benzoxazolyl, benzothiazolyl, benzimidazolyl, indolyl, and the like and those rings substituted with one or more substituents listed above in the definition of the term "aryl".

5 The term halogen is used to refer to fluorine, chlorine, bromine and iodine.

In the above definitions the unsubstituted and substituted C<sub>1</sub>-C<sub>10</sub> alkyl groups or portion of groups mentioned refer to fully saturated hydrocarbon radicals containing one to ten carbon atoms, either straight or  
10 branched chain, and such alkyl radicals substituted with one or more of the following: C<sub>3</sub>-C<sub>8</sub> cycloalkyl, aryl, hydroxy, cyano, -O-C<sub>2</sub>-C<sub>4</sub> alkylene OH, -O-C<sub>2</sub>-C<sub>4</sub> alkylene O<sub>2</sub> C-C<sub>1</sub>-C<sub>6</sub> alkyl, -S-C<sub>2</sub>-C<sub>4</sub> alkylene-OH, chlorine, fluorine, -O-C<sub>1</sub>-C<sub>6</sub> alkyl, -O-aryl, -SO<sub>2</sub> aryl, -SO<sub>2</sub>-C<sub>1</sub>-C<sub>6</sub> alkyl, 2-  
15 pyrrolidino, phthalimidino, phthalimido, succinimido, glutarimido, o-benzoic sulfimide, vinyl sulfonyl, -NHCO C<sub>1</sub>-C<sub>6</sub> alkyl, NHCOH, -NHSO<sub>2</sub>-C<sub>1</sub>-C<sub>6</sub> alkyl, NHSO<sub>2</sub> aryl, -NHCO aryl, -NH-CO<sub>2</sub>-C<sub>1</sub>-C<sub>6</sub> alkyl, -SO<sub>2</sub>NH<sub>2</sub>, -SO<sub>2</sub>-NH-C<sub>1</sub>-C<sub>6</sub> alkyl, -SO<sub>2</sub>N-(C<sub>1</sub>-C<sub>6</sub> alkyl)<sub>2</sub>, -CO<sub>2</sub>-C<sub>1</sub>-C<sub>6</sub> alkyl, CONH<sub>2</sub>, -CONH-C<sub>1</sub>-C<sub>6</sub>  
20 alkyl, -CO<sub>2</sub>-aryl, -CON(C<sub>1</sub>-C<sub>6</sub> alkyl)<sub>2</sub>, -CONH aryl, -CONH(C<sub>1</sub>-C<sub>6</sub> alkyl) aryl, -SO<sub>2</sub>N(C<sub>1</sub>-C<sub>6</sub> alkyl) aryl, -SO<sub>2</sub>-NH-C<sub>3</sub>-C<sub>8</sub> cycloalkyl, -CONH-C<sub>3</sub>-C<sub>8</sub> cycloalkyl, -OCO<sub>2</sub>-C<sub>1</sub>-C<sub>6</sub> alkyl, -O C<sub>2</sub>-C<sub>4</sub> alkylene CN; groups of the formulae:

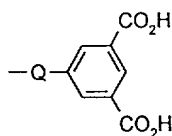
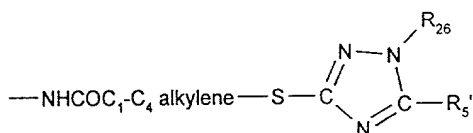
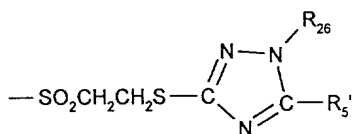
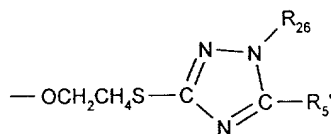
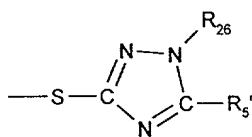
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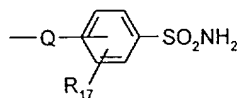
wherein Y<sub>2</sub> is selected from 1,2-phenylene; 1,2 pheylene substituted with C<sub>1</sub>-C<sub>6</sub> alkyl, C<sub>1</sub>-C<sub>6</sub> alkoxy, halogen, -CO<sub>2</sub>H,  
30 -CO<sub>2</sub> C<sub>1</sub>-C<sub>5</sub> alkyl or nitro; C<sub>2</sub>-C<sub>4</sub> alkylene, vinylene, -O CH<sub>2</sub>-, -SCH<sub>2</sub>-, -CH<sub>2</sub>OCH<sub>2</sub>-, -OCH<sub>2</sub>CH<sub>2</sub>-, -CH<sub>2</sub>SCH<sub>2</sub>-, -NHCH<sub>2</sub>-,

-NHCH<sub>2</sub>CH<sub>2</sub>, -N(C<sub>1</sub>-C<sub>6</sub> alkyl)CH<sub>2</sub>-, NHC(C<sub>1</sub>-C<sub>6</sub> alkyl)<sub>2</sub>, -N(C<sub>1</sub>-C<sub>6</sub> alkyl)CH<sub>2</sub>CH<sub>2</sub> or -NHC(aryl)<sub>2</sub>-; groups of the formulae:

-SR<sub>25</sub>, -SO<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>SR<sub>25</sub>, -OCH<sub>2</sub>CH<sub>2</sub>SR<sub>25</sub>,



or



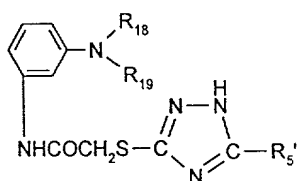
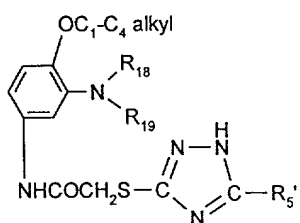
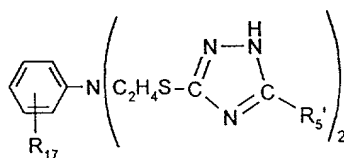
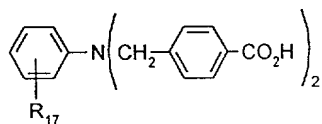
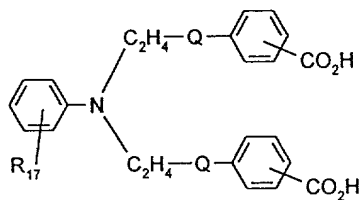
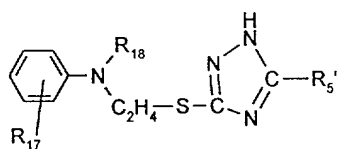
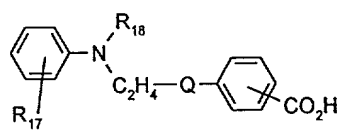
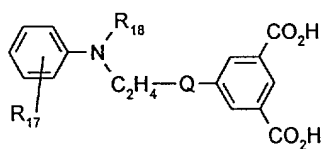
5

wherein R<sub>26</sub> is selected from hydrogen, C<sub>1</sub>-C<sub>10</sub> alkyl, C<sub>2</sub>-C<sub>4</sub> alkylene-OH, C<sub>2</sub>-C<sub>4</sub> alkylene-CO<sub>2</sub>H, C<sub>2</sub>-C<sub>4</sub> alkylene-CO<sub>2</sub> C<sub>1</sub>-C<sub>6</sub> alkyl, chloro, C<sub>1</sub>-C<sub>6</sub> alkoxy, C<sub>1</sub>-C<sub>4</sub> alkylene-arylene-CO<sub>2</sub>H, C<sub>2</sub>-C<sub>4</sub> alkylene-O-arylene-CO<sub>2</sub>H or C<sub>2</sub>-C<sub>4</sub> alkylene-S-arylene-CO<sub>2</sub>H and R<sub>5'</sub>, R<sub>17</sub>, R<sub>25</sub> and Q are as defined previously:

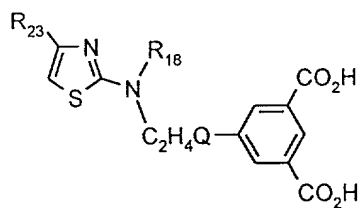
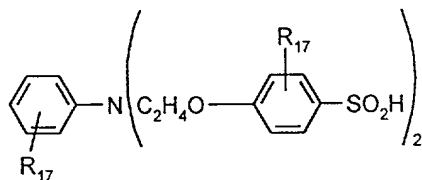
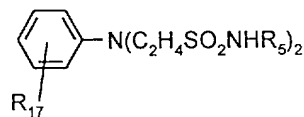
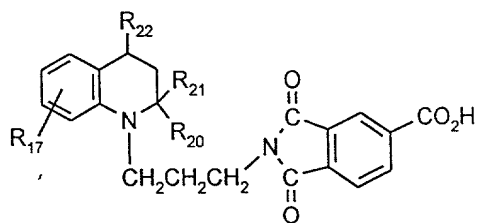
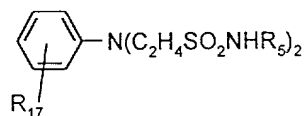
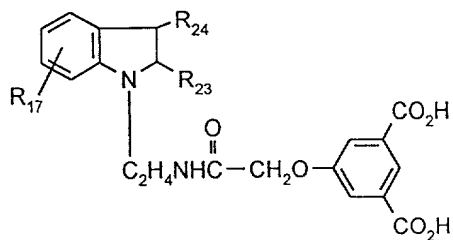
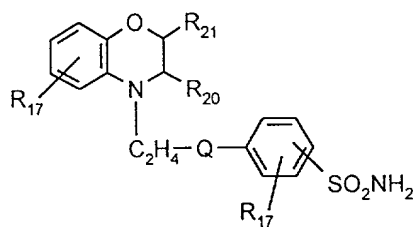
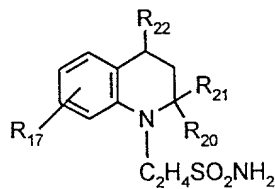
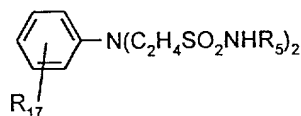
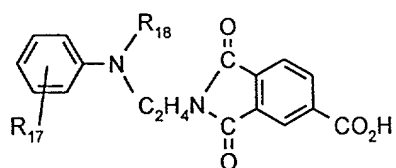
The term "light absorbing" is used to indicate the property of absorbing near ultra violet, visible or near infrared light, more particularly absorbing light between the wavelengths of 300-1200 nm, preferably between about 325 nm and 1100 nm, and most preferably between about 325 nm and 1000 nm.

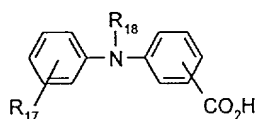
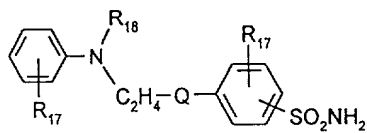
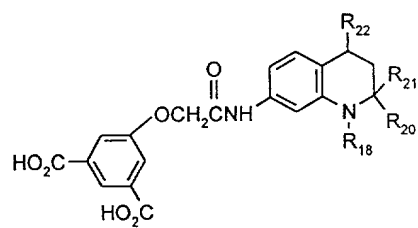
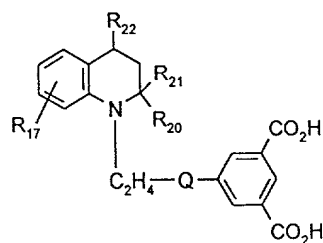
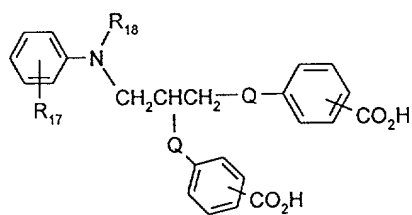
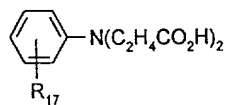
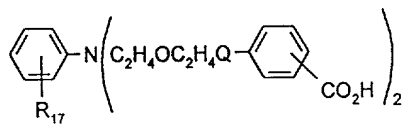
Typical aromatic amines which are useful as the coupling components to prepare compounds of Formulae VI, VII and VIII and as intermediates for preparing the compounds of Formula VIII, VIIIA, IX, X, XI and XII are as follows:

20









wherein Q, R<sub>5</sub>', R<sub>17</sub>, R<sub>18</sub>, R<sub>19</sub>, R<sub>20</sub>, R<sub>21</sub>, R<sub>22</sub> and R<sub>23</sub> are as defined previously.

5 Typical diazotizable amines (R<sub>6</sub> NH<sub>2</sub>) useful in the preparation of azo, disazo and bis-azo compounds of Formulae VI, VII, and VIIa, respectively, are adequately disclosed in the literature, e.g.:

M. Weaver and L. Shuttleworth, Dyes and Pigments, 3 (1982) 81-121;

10 L. Shuttleworth and M. Weaver, Chem. Appl. Dyes, 1990, 107-63, edited by D. Waring and G. Hallas, Plenum, New York, N.Y.;

U.S. Pat. Nos. 3,438,961; 3,573,273; 3,639,384;  
3,707,532; 3,790,557; 3,816,388; 3,816,392; 3,878,189;  
15 3,980,634; 4,012,372; 4,039,522; 4,049,643; 4,083,684;  
4,083,844; 4,097,475; 4,105,655; 4,119,621; 4,140,683;  
4,180,503; 4,189,428; 4,207,233; 4,211,696; 4,264,495;  
4,283,332; 4,400,318; 4,431,585; 4,456,551; 4,487,719;  
4,542,207; 4,564,673; 4,619,991; 4,621,136; 4,650,861;  
20 4,668,775; 4,734,490; 4,751,288; 4,760,133; 4,764,600;  
4,837,269; 4,841,036; 4,843,153; 4,888,432; 4,960,874;  
5,037,966; 5,132,411; 5,144,015; 5,283,326; 5,296,325;  
5,352,774.

Typical coupling components H-Z useful in preparing  
25 azo compounds, disazo and azo-methine compounds of Formula VI, VII and XIII, respectively, are disclosed in the literature, e.g: H. R. Schwander, Dyes and Pigments, 3(1982) 133-160; L. Shuttleworth and M. Weaver, Chem. Appl. Dyes, 1990, 107-63, edited by D. Waring and G.  
30 Hallas, Plenum, New York, NY; U.S. Patent No. 3,639,384;  
3,639,385; 3,657,215; 3,673,169; 3,816,388; 3,829,410;  
3,919,188; 3,950,130; 3,980,634; 4,041,025; 4,097,475;  
4,119,621; 4,179,435; 4,234,482; 4,283,332; 4,341,700;  
4,400,318; 4,431,585; 4,396,547; 4,619,992; 4,642,339;  
35 4,650,861; 4,668,775; 4,764,600; 4,837,269; 4,843,153;  
5,235,047; 5,283,326; 5,352,774.

Typical active methylene compounds useful in the preparation of methine, arylidene, polymethine, azamethine and azo-methine compounds corresponding to Formulae VIII, VIIIA, VIIIB, IX and XIII, respectively, are disclosed in the literature, e.g. U.S. Pat. Nos. 4,338,247; 4,617,373; 4,617,374; 4,707,537; 4,749,774; 4,826,903; 4,845,187; 4,950,732; 4,981,516 and 5,283,326.

According to the present invention the light-absorbing polymeric and cyclic compositions are incorporated into a wide variety of thermoplastic polymers using conventional techniques, e.g. solution or melt blending, such as those employed to incorporate other additives in such polymers (see R. Gächter and H. Müller, Editors: *Plastics Additives Handbook*, Hansu Publishers, New York, 1985, pp. 507-533; 729-741). For example, the light absorbing polymeric and cyclic compositions may be dry blended in the form of pellets or powders with or without adhesion promoters or dispersing agents. This premix can be subsequently processed on extruders or injection molding machines. Other conventional additives such as plasticizers, nucleating agents, flame retardants, lubricants, etc. may also be present in the final thermoplastic composition.

A wide range of thermoplastic polymers useful for blending with the light absorbing polymeric and cyclic compositions are known in the art and includes the homopolymers, copolymers and blends of polyesters, e.g., poly(ethylene terephthalate); polyolefins, e.g., polypropylene, polyethylene, linear low density polyethylene, polybutylene, and copolymers made from ethylene, propylene and/or butylene; copolymers from acrylonitrile, butadiene, and styrene; copolymers from styrene and acrylonitrile; polyamides, e.g., Nylon 6 and Nylon 66; polyvinyl chloride; polyurethanes; polyvinylidene chloride; polycarbonates; cellulose esters, e.g., cellulose acetate, propionate, butyrate, or mixed

esters; polyacrylates, e.g., poly(methyl methacrylate); polyimides; polyester-amides; polystyrene; and mixtures or blends thereof etc.

It should also be appreciated that a multiplicity of  
5 colors may be obtained by combining individual colors,  
e.g., subtractive colors such as yellow, magenta and cyan  
according to known color technology (see N. Ohta,  
*Photographic Science and Engineering*. Volume 15, No. 5,  
Sept.- Oct. 1971, pp. 395-415).

10           The particular chromophore groups present will, of course, determine the color (hue + value + chroma) of the colored polymer composition and finally the color (hue + value + chroma) of the thermoplastic polymer blends of the present invention. A large gamut of colors may be  
15   obtained, as noted above.

The actual amount of the light absorbing polymers used in combination with thermoplastic polymer will depend upon the inherent tinctorial strength of the chromophore used to prepare the light absorbing polymer, the mole % of the light absorbing monomer used to prepare the light absorbing polymer and the required level of light absorption necessary to achieve a certain property. Typically, the amount of light-absorbing polymer added to the thermoplastic polymer is such that the total amount of light-absorbing polymer in the final thermoplastic blend is from about .001% by weight to about 20% by weight, preferably from about 0.01% by weight to about 10% by weight. The final thermoplastic polymer blends thus provided are useful as a variety of molded and extruded articles, including thick and thin plastic films, plastic sheeting, molded plastic articles, containers and fibers, and the like.

When the light-absorbing polymeric compositions absorb visible light they may be used to impart light or heavy shades of a variety of colors to thermoplastics. 35 Certain compounds which possess unique visible light-

absorbing properties are useful also as toners in imparting a desirable neutral to slightly blue hue to polyesters having a yellow appearance as described in U.S. Patent 5,384,377, which discloses the copolymerization of  
5 certain thermally stable colorants for this purpose during polyester manufacture. Some of the infra-red absorbing polymeric and cyclic compositions are useful in imparting invisible markings to thermoplastics as described in U.S. Pat. No. 5,461,136, wherein the infrared absorbing  
10 compounds are fluorescent in the near infrared and are copolymerized into the thermoplastic condensation polymer during manufacture. The ultra violet absorbing polymeric and cyclic compositions may be used to impart ultra violet (UV) light screening properties to the thermoplastics; to  
15 serve as optical brighteners for the thermoplastics or to serve as UV stabilizers for the polymers themselves or for other light absorbers such as colorants.

The weight average molecular weights (Mw) and the number average molecular weights (Mn) of the polymeric  
20 compositions were determined using gel permeation chromatography (GPC) analysis.

The following examples illustrate further the practice of the invention.

25

Example 1

A mixture of 1,5-bis(2-carboxyphenylthio) anthraquinone (25.60 g, 0.05 mole), 1,2-ethanediol, dimethanesulfonate (10.90 g, 0.05 mole), potassium  
30 carbonate (13.82 g, 0.10 mole) and N-methyl-2-pyrrolidinone (NMP) (400 mL) was heated with stirring at 125°C for 1.0 hr. The reaction mixture was poured into methanol (600 mL) with stirring. The yellow polymeric product was collected by filtration and washed with  
35 methanol until filtrate was essentially clear. The methanol- wet filter cake was slurried in 1.0 L of water,

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The following examples illustrate further the practice of the invention.

25

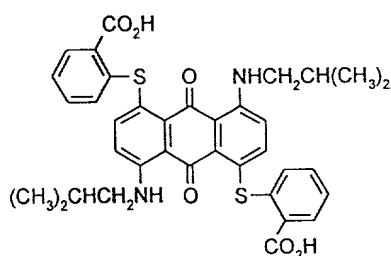
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35 methanol until filtrate was essentially clear. The methanol- wet filter cake was slurried in 1.0 L of water,

the mixture acidified by the addition of acetic acid and the yellow product was collected by filtration, washed with hot water and dried in air (yield- 21.16 g). By gel permeation chromatography (GPC) the polymeric product has  
5 a weight average molecular weight of 6,083, and number average molecular weight of 3,000 and a polydispersity value of 2.03.

### Example 2

10 A mixture of a blue anthraquinone compound (19.65 g 0.03 mole) containing two carboxy groups and having the following structure:



15 1,2-ethanediol, dimethanesulfonate (6.54g, 0.03m), potassium carbonate (8.28 g, 0.06 mole) and N,N-dimethylformamide (DMF) (100 mL) was heated with stirring at about 95°C for 1.5 hr. The reaction mixture became too  
20 thick to stir effectively and additional DMF (50 mL) was added to facilitate stirring. Stirred about 15 min. longer at about 95°C, and then added methanol (100 mL) with good stirring to the slightly cooled reaction mixture. The blue polymeric product was collected by  
25 filtration and washed with methanol. The methanol-wet filter cake was added to water (600 mL) and the mixture was acidified with acetic acid, and then the polymeric product was collected by filtration, washed with water and dried in air (yield 18.18 g). By GPC analysis the blue  
30 polymer had a molecular weight average of 3,038, a number



average molecular weight of 1,814 and a polydispersity of 1.67.

Example 2a

5 A mixture of 1,5-bis (isobutylamino)-4,8-dibromoanthraquinone (25.3 g, 0.05 mole), thiosalicylic acid (23.1 g, 0.15 mole), anhydrous  $K_2CO_3$  (20.7 g, 0.15 mole), cupric chloride dihydrate (1.2 g) and DMF (250 mL) was heated at 90-95°C with stirring for 2.0 hours. Thin  
10 layer chromatography (TLC) using 1:1 tetrahydrofuran (THF): cyclohexane showed complete conversion of the red starting material to the desired blue polar product. The reaction mixture was allowed to cool and then was drowned into water (800 mL). The blue solid was precipitated by  
15 acidification with acetic acid with stirring. The mixture was heated to about 60°C with occasional stirring and the solid was collected by filtration, washed with hot water and dried in air. Further purification was accomplished by reslurrying the product in hot methanol (300 mL),  
20 allowing to cool to room temperature, collecting by filtration, washing with methanol and air drying to yield the starting material (31.5 g) for Example 2.

Example 2b

25 1,5-Bis(isobutylamino)anthraquinone (28.0 g, 0.08 mole) was added to DMF (300 mL) and the mixture stirred at room temperature. A solution of 1,3-dibromo-5,5-dimethylhydantoin (23.0 g, 0.08 m) dissolved in DMF (75.0 mL) was added dropwise to the reaction mixture while  
30 warming to about 50°C. After complete addition of the brominating agent, the reaction mixture was heated at 50-60°C for 1.5 hours, allowed to cool and then drowned by gradual addition to water (500 mL) with stirring. The red product was collected by filtration, washed with water and  
35 dried in air. The yield of product was 39.6 g and field desorption mass spectrum analysis (FDMS) showed the

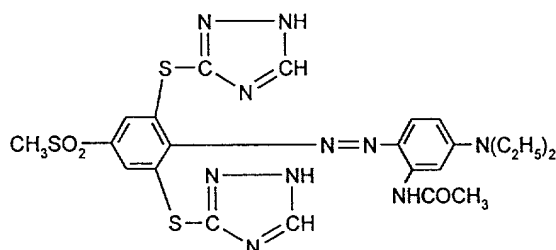
product to be 1,5-bis(isobutylamino)-4,8-dibromoanthraquinone used as the intermediate in Example 2a.

5 Example 2c

097166 13900  
A mixture of 1,5-dichloroanthraquinone (69.5 g, 0.25 mole), isobutylamine (100 g, 1.4 mole) and 2-ethoxyethanol (400 mL) was heated at reflux for 36.0 hours and allowed to cool. Methanol (400 mL) was added to make the mixture  
10 containing the crystallized product more stirrable. The dark red product was collected by filtration, washed with methanol, reslurried in hot methanol and allowed to cool, collected by filtration, washed with methanol and dried in  
15 air (yield - 67.7 g). FDMS showed the product to be the 1,5-bis(isobutylamino)anthraquinone in high purity which was used as the starting material for Example 2b.

Example 3

A mixture of an azo compound (2.93 g, 0.005 m)  
20 containing two 1(H)-1,2,4-triazol-3-thio groups and having the following structure:



25 1,2-ethanediol, dimethanesulfonate (1.08 g, 0.005 mole), potassium carbonate (1.50 g) and DMF (25.0 mL) was heated at about 95°C with stirring for 2.5 hrs. The reaction mixture was drowned into methanol (150 mL) and the red polymeric product was collected by filtration, washed with  
30 water containing a little acetic acid and then washed with hot water and dried in air (yield- 2.35 g). The polymer

by GPC analysis had a weight average molecular weight of 5,396, a number average molecular weight of 3,044 and a polydispersity value of 1.77.

5 Example 4

Eastar® PETG copolyester 6763, a poly(ethylene-1,4-cyclohexanedimethylene) terephthalate, (Eastman Chemical Co.) (400 g. of previously dried pellets) was dry blended with the yellow anthraquinone polymeric composition (0.12 g) of Example 1. The blend was extruded with a C. W. Brabender  $\frac{3}{4}$  in. extruder, equipped with a mixing screw, at 250°C into a water bath and the extrudate pelletized.

The pellets were redried at 70°C for about 17 hrs. at a pressure of about 1-5 torr. A portion of the dried pellets (1.40g) was pressed into a 18-20 mil film at 250°C using a 2-inch diameter circular mold in a Pasadena Hydraulic, Inc. press using 12,000 pounds ram force (4 inch ram). A transparent yellow film was produced with excellent color development, which contained about 300 ppm by weight of the yellow polymeric composition.

Example 5

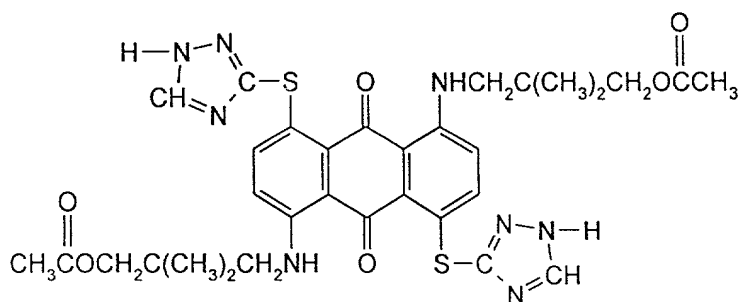
Example 4 was repeated using 0.12 g of the blue anthraquinone polymeric composition of Example 2 to give a bright blue transparent copolyester film with good color development.

Example 6

Example 4 was repeated using 0.12 g of the red azo polymeric composition of Example 3 to produce a bright red transparent film having good color development.

Example 7

A mixture of a blue anthraquinone compound (3.46 g, 0.005 mole) containing two acidic 1(H)-1,2,4-triazol-3-ylthio groups and having the following structure



1,2-ethanediol, dimethanesulfonate (1.09 g, 0.005 mole) DMF (30 mL) and potassium carbonate (1.5 g) was heated with stirring at about 95°C for 2.0 hours and then drowned into methanol (100 mL). The blue polydye was collected by filtration and washed with methanol. The methanol-wet cake was reslurried in water (400 mL) and the stirred mixture was acidified by addition of acetic acid and heated to about 60°C. The final polymeric product was collected by filtration, washed with water and dried in air (yield - 1.5 g). Absorption maxima were observed at 594,636 nm in a solution of DMF in the visible light absorption spectrum. By GPC, the polydye has a weight average molecular weight (Mw) of 3,769, a number average molecular weight (Mn) of 2,119 and a polydispersity of 1.78.

Example 7a

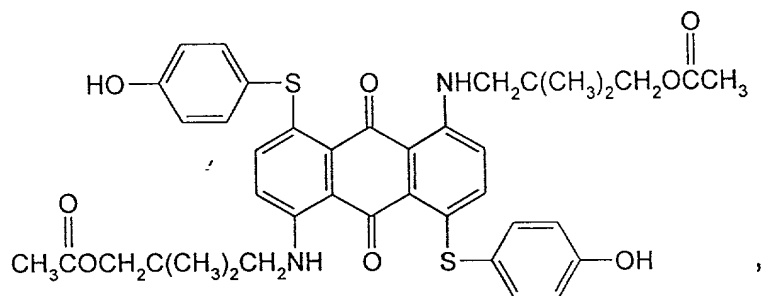
25

A mixture of 1,5-bis[(3-acetoxy-2,2-dimethylpropyl)amino-4,8-dibromoanthraquinone (6.50 g, 0.01 mole) (product of Example 2 - Invention Report Docket No. 70524), 3-

mercapto-1(H)-1,2,4-triazole (3.03 g, 0.03 mole),  
potassium carbonate (4.15 g, 0.03 mole), cupric chloride  
dihydrate (0.65 g) and DMF (100 mL) was heated 14 hours at  
about 100- 105°C. The reaction mixture was drowned into a  
5 mixture of water (400 mL) and 10% aqueous solution of  
hydrochloric acid (200 mL). The blue product was  
collected by filtration, washed with hot water and dried  
in air (yield - 6.58 g). FDMS supported the desired  
structure of the starting anthraquinone compound for  
10 Example 7.

Example 8

15 A mixture of blue anthraquinone compound (2.48 g, 0.0033  
mole) having the following structure



20 1,2-ethanediol, dimethanesulfonate (0.73 g, 0.0033 mole),  
potassium carbonate (0.5 g) and DMF (30.0 mL) was heated  
at about 95°C for 3.0 hours. The reaction mixture was  
drowned into methanol (150 mL) with stirring and the blue  
polydye product was collected by filtration and washed  
25 with methanol. The methanol-wet cake was reslurried in  
water (200 mL) and the mixture acidified with acetic acid.  
Collecting the blue solid by filtration, washing with hot  
water and air drying gave 1.21 g of polydye product, which  
has absorption maxima at 606,652 nm in DMF in the visible

absorption spectrum, a weight average molecular weight of 4,453, a number average molecular weight of 2,721 and a polydispersity of 1.6.

5 Example 8a

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A mixture of 1,5-bis[(3-acetoxy-2,2-dimethylpropyl)amino]-4,8-dibromoanthraquinone (19.56 g, 0.03 mole), p-hydroxybenzenethiol (17.64 g, 0.14 mole), potassium carbonate (19.32 g, 0.14 mole), cupric chloride dihydrate (1.0 g) and DMF (150 mL) was heated and stirred at 90-95°C for 7.0 hours and then at 120°C for about 2.0 additional hours. TLC (50:50 THF: cyclohexane) showed mostly the desired blue product, but still a small amount of violet half-reacted product was present. The reaction mixture was drowned into methanol (500 mL) and the mixture allowed to cool. After crystallization, the blue solid was collected by filtration, washed with methanol, washed with hot water and then dried in air (yield - 17.6 g). FDMS supported the desired structure of the starting anthraquinone compound for Example 8. In the visible light absorption spectrum in DMF, a maximum absorbance ( $\lambda_{\text{max}}$ ) was observed at 652 nm (extinction coefficient  $\epsilon$  of 24,638).

25

Example 9

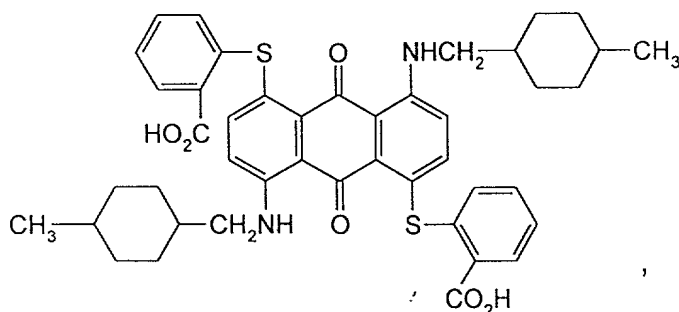
A mixture of 1,4-bis-(2,6-dimethyl-4-hydroxyanilino)anthraquinone (4.78 g, 0.01 mole) (Synthesis Example 1 of U. S. Patent 3,918,976), 1,2-ethanediol, dimethanesulfonate (2.18 g, 0.01 mole), potassium carbonate (3.0 g) and DMF (60 mL) was heated at 90-95°C with stirring for 4.0 hours. After drowning the reaction mixture into methanol (300 mL), the product was collected by filtration and washed with methanol until filtrate was essentially colorless. The methanol-wet cake

35

was reslurried in 100 mL water and acidified by adding acetic acid with stirring. After heating to about 50°C, the product was collected by filtration, washed with hot water and dried in air (yield - 1.2 g). By GPC, the blue polydye had a weight average molecular weight (Mw) of 2,764, a number average molecular weight (Mn) of 1,607 and a polydispersity of 1.72. In DMF, the visible light absorption maxima were at 586,630 nm.

10 Example 10

A mixture of an anthraquinone diacidic compound (1.52 g, .002 mole) having the following structure



Example 10a

5 A mixture of 1,5-bis-(4-methylcyclohexanemethylamino)-4,8-  
dibromoanthraquinone (20.0 g, 0.0324 mole), thiosalicyclic  
acid (11.55 g, 0.075 mole), potassium carbonate (10.35 g,  
0.075 m), cupric chloride dihydrate (1.0 g) and DMF (175  
mL) was heated at about 95°C for 4.0 hours and then  
10 drowned into acetone (400 mL). The solid which  
crystallized was collected by filtration, washed with  
acetone until the filtrate was no longer red. The  
dipotassium salt of the diacidic anthraquinone compound  
was dissolved by adding to water (500 mL) and stirring.  
The blue product which was precipitated by acidification  
15 with acetic acid was collected by filtration, washed with  
hot water and then dried in air (yield - 21.5 g). FDMS  
indicated the structure to be consistent with that given  
above in Example 10 for the starting diacidic  
anthraquinone compound.

20

Example 10b

A solution of 1,5-bis-(4-  
methylcyclohexanemethylamino)anthraquinone (65.0 g, 0.142  
25 mole) dissolved in DMF (1.0 L) by stirring at about 55°C  
was treated with a solution of N-bromosuccinimide (50.5 g,  
0.284 mole) in DMF (200 mL). After addition was  
completed, the bromination reaction was completed by  
heating at 55-60°C for 2.0 hours. Water (1.0 L) was added  
30 to precipitate the red product which was collected by  
filtration, washed with water and dried in air. After  
being reslurried in hot methanol and cooling, the product  
was collected by filtration, washed with a little methanol  
and air dried (yield - 84.0 g). FDMS indicated the  
35 structure to be that of the starting, dibrominated  
anthraquinone compound of Example 10a.



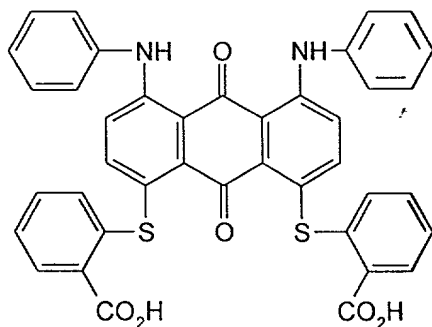
Example 10c

5 A mixture of 1,5-dichloroanthraquinone (48.0 g, 0.17 mole), 4-methyl-1-aminomethylcyclohexane (88.9 g, 0.70 mole), 2-ethoxyethanol (400 mL) was stirred and heated at reflux for 35.0 hours and the reaction mixture allowed to cool. The red product was precipitated by the addition of methanol and was collected by filtration, washed with  
10 methanol and dried in air (yield - 66.0 g). FDMS indicated the product to be the starting anthraquinone compound for Example 10b.

Example 11

15

A mixture of diacidic anthraquinone compound (0.69 g, 0.001 m) having the following structure



20

1,6-hexanediol, dimethanesulfonate (0.27 g, 0.001 mole), potassium carbonate (0.3 g) and DMF (5.0 mL) was heated with occasional stirring for 2.5 hours at about 95°C. The reaction mixture was drowned into methanol (100 mL) and  
25 the product collected by filtration, washed with methanol, water containing a little acetic acid and then finally with hot water and air dried (yield - 0.45 g). The blue polydye had an absorption maximum at 610 nm in DMF, a

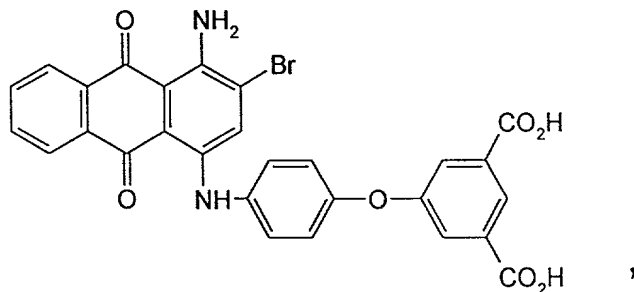
5 Example 11a

20 Example 11b

The potassium salt of thiosalicyclic acid (4.75 g, 0.03 mole) was made by addition to DMF (75 mL) and heating in the presence of potassium carbonate (8.70 g, 0.06 mole) for 2.0 hours at about 95°C. The cooled mixture was added to a solution of 1,8-dichloro-4,5-dinitroanthraquinone (5.51 g, 0.015 mole) dissolved in DMF (150 mL) at about 0-5°C with stirring. The reaction mixture was allowed to warm to about 25°C with stirring continued for 2.0 hours and then poured into water. The product was obtained in essentially quantitatively yield by slowly acidifying with 10% hydrochloric acid and was then collected by filtration, washed with water and dried in air. FDMS indicated the product to be mostly the starting anthraquinone compound used in Example 11a.

Example 12

A mixture of the diacidic anthraquinone compound (0.85 g. 0.0015 m) having the following structure



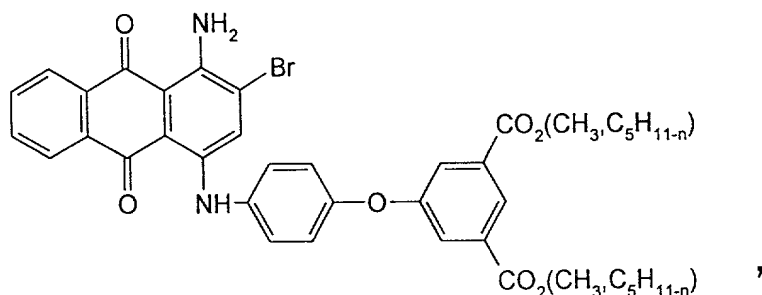
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1,6-hexanediol, dimethanesulfonate (0.41 g, 0.0015 m), potassium carbonate (0.5 g) and DMF (5.0 mL) was heated at about 95°C for 2.0 hours with occasional stirring. The reaction mixture was drowned into methanol (100 mL) and the blue polydye was collected by filtration, washed with methanol, water containing a little acetic acid and finally hot water and then dried in air (yield - 0.62 g). GPC analysis indicated a weight average molecular weight of 20,020, a number average molecular weight of 2,313 and a polydispersity of 8.66. An absorption maximum was observed at 591 nm in the visible light absorption spectrum in DMF.

20

Example 12a

The anthraquinone diester compound (4.00 g) having the following structure



50% aqueous sodium hydroxide (2.40 g) and 2-ethoxyethanol (60 mL) were combined and heated with stirring at about 95°C for 0.5 hour. Hydrolysis of ester groups appeared to be complete by TLC (50:50 THF:cyclohexane). The reaction mixture was drowned into water (600 mL) and the blue solution acidified using acetic acid. The blue solid was collected by filtration washed with water and dried in air (yield - 3.80 g). FDMS indicated the structure to be mostly that of the starting diacidic anthraquinone compound in Example 12 plus a small amount of a violet compound probably produced by displacement of the bromine atom with the 2-(ethoxy)ethoxy group.

Example 12b

A mixture of 1-amino-2,4-dibromoanthraquinone (7.62 g, 0.02 mole), dimethyl 5(4-aminophenoxy)isophthalate (9.03 g, 0.03 mole), 1-pentanol (100 mL), potassium acetate 4.0 g), and cupric acetate (0.2 g) was heated at reflux for 4.0 hours and until all of the starting material had been used up as indicated by TLC analysis (20:80 THF:cyclohexane). Several blue components presumed to be a mixture of ester products produced by transesterification were observed. The reaction mixture was drowned into methanol (100 mL) and the product was collected by filtration, washed thoroughly with methanol to remove a red by-product and then washed with water and dried in air (yield - 7.81 g). FDMS indicated ions corresponding to the dimethylester,

monopentyl ester and dipentylester of the product - the structure of the starting material for Example 12a.

Example 12c

5

A mixture of dimethyl 5-(4-nitrophenoxy)isophthalate (30.0 g, 0.09 mole), isopropanol alcohol (350 mL) and ethanol wet Raney nickel catalyst (5.0 g) was hydrogenated at 90°C for 4.0 hours at 1500 psi hydrogen pressure in an autoclave. Isopropanol (100 mL) was added to the reaction mixture from the autoclave and the solid product dissolved by heating. The Raney nickel was removed by hot filtration and the filtrate allowed to cool. The off-white solid was collected by filtration and dried in air (yield - 17.8 g). FDMS indicated the product to be dimethyl 5-(4-aminophenoxy)isophthalate used in Example 12b.

10

15

Example 12d

20

A mixture of 1-chloro-4-nitrobenzene (47.1 g, 0.30 mole), dimethyl 5-hydroxyisophthalate (63.0 g, 0.30 mole), anhydrous potassium carbonate (41.4 g), potassium iodide (0.2 g) and DMF (200 mL) was heated at 120- 125°C for 1.5 hours, under a slow nitrogen sweep allowing some distillate to be removed (about 75 mL) via a Dean-Stark trap. Additional DMF (50 mL) was added back to the reaction mixture and heating continued for an additional 1.5 hours while an additional amount of distillate (25 mL) was allowed to collect in the Dean-Stark trap. The reaction mixture was allowed to cool to about 45°C. A heavy slurry of pale yellow product resulted which was diluted further by the addition of an ice-water mixture (350 g) with good stirring. Filtration followed by washing with water and drying in air gave the pale yellow dimethyl 5-(4-nitrophenoxy)isophthalate (90.7 g)

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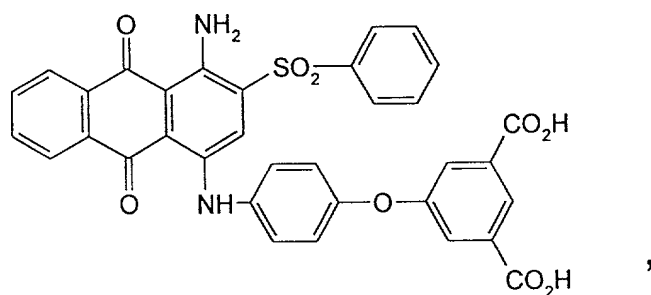
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(structure supported by FDMS) which was used in Example 12c.

### Example 13

5

A mixture of the diacidic anthraquinone compound (1.26 g, 0.002 mole) having the following structure



- 10 1,6-hexandiol, dimethanesulfonate (0.58 g, 0.002 mole), potassium carbonate (0.5 g) and DMF (6.0 mL) was heated at 90-95°C for 2.0 hours with occasional stirring. The reaction mixture was drowned into methanol (100 mL) and the dark blue-green polydye was collected by filtration,
- 15 washed with methanol, water containing a little acetic acid and finally with water and then dried in air (yield 1.13 g). GPC analysis indicated a weight average molecular weight of 14,776, a number average molecular weight of 2,514 and a polydispersity of 5.88. An
- 20 absorption maximum was observed at 620 nm in the visible light absorption spectrum in DMF.

### Example 13a

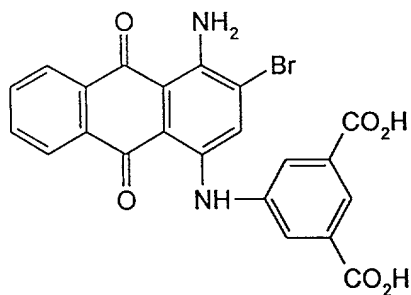
- 25 A portion (1.72 g, 0.003 mole) of the bromoanthraquinone product of Example 12a, benzenesulfinic acid, Na salt (0.98 g, 0.006 mole), potassium carbonate (1.38 g) and DMF (25 mL) were mixed and the reaction mixture heated with stirring at 90-95°C for 1.0 hour. A bathochromic shift in

color was observed as the 2-bromo substituent was replaced by the 2-phenylsulfonyl group on the anthraquinone nucleus. The greenish-blue solution was drowned into acetone (100 mL) and the solid material was collected by  
5 filtration and washed with acetone until the filtrate was pale blue. The acetone-wet solid was added with stirring to water (200 mL) and the mixture acidified with acetic acid. After being heated to about 75°C, the reaction mixture was filtered and the dark blue solid was washed  
10 with hot water and dried in air (yield - 1.50 g). FDMS indicated the structure to be that of the starting diacidic anthraquinone compound used in Example 13.

Example 14

15

A mixture of the diacidic anthraquinone compound (1.45 g, 0.003 mole) having the structure



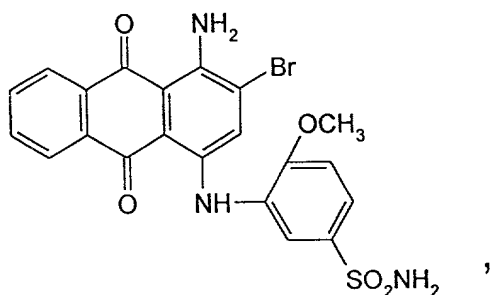
20 1,6-hexanediol, dimethanesulfonate (0.82 g, 0.003 mole), potassium carbonate (0.5 g) and DMF (8.0 mL) was heated at about 95°C for 2.0 hours. The reaction mixture was drowned into methanol (100 mL) and the blue polydye was collected by filtration and washed with methanol, water  
25 containing a little acetic acid and finally hot water and dried in air (yield - 1.10 g). GPC analysis indicated a weight average molecular weight of 3,727, a number average weight of 1,031 and a polydispersity of 3.61. Absorption

maxima were observed at 623 nm and 585 nm in the visible light absorption spectrum in DMF.

Example 15

5

A mixture of the diacidic anthraquinone compound (1.50 g, 0.003 mole) having the following structure



- 10 1,6-hexanediol, dimethanesulfonate (0.82 g, 0.003 mole), potassium carbonate (0.5 g) and DMF (8.0 mL) was heated with occasional stirring at about 95°C for 2.0 hours. The reaction mixture was then drowned into methanol (100 mL) and the blue polydye was collected by filtration, washed
- 15 with methanol, water containing a little acetic acid, and hot water and then dried in air (yield- 0.90 g). An absorption maximum at 591 nm was observed in the visible light absorption spectrum in DMF.

20 Example 15a

- To DMF (40mL) was added 1-amino-2-bromo-4-(5-chlorosulfonyl-2-methoxyanilino) anthraquinone (4.0 g) with stirring. When solution appeared to be complete, conc. ammonium
- 25 hydroxide (4.0 g) was added and stirring was continued at ambient temperature for 30 minutes. TLC using 50:50 THF:cyclohexane indicated complete reaction of the sulfonyl chloride compound to produce the desired sulfonamide. The reaction mixture was drowned into water



and the blue product was collected by filtration, washed with water and air dried (yield- 3.8 g). FDMS indicated the structure to be that of the starting compound for Example 15.

5

Example 15b

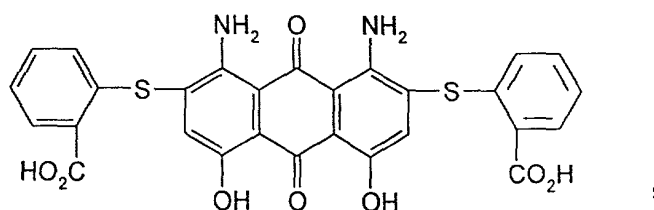
To chlorosulfonic acid (100 mL) was added 1-amino-4-o-anisidino-2-bromoanthraquinone (10.0 g, 0.0236 mole) portionwise with good stirring at 25-30°C. After addition was completed, the reaction mixture was stirred at room temperature for 1.0 hour. The reaction mixture was added in a fine stream to cold isopropanol (800 mL) with stirring. The blue product was collected by vacuum filtration on a sintered glass funnel, washed with isopropanol and dried in air (yield- 10.3 g) and used without further purification in Example 15a.

15

Example 16

20

A mixture of the diacidic anthraquinone compound (0.58 g, 0.001 m) having the following structure



1,2-ethanediol, dimethanesulfonate (0.22 g, 0.001 m), potassium carbonate (0.3 g) and DMF (5.0 mL) was heated at 95°C for 2.5 hours. The reaction mixture was drowned into methanol (100 mL) and the greenish-blue polydye was collected by filtration, washed with methanol, water containing a little acetic acid and water and then air dried (yield - 0.33 g). GPC analysis indicated a weight

25

30

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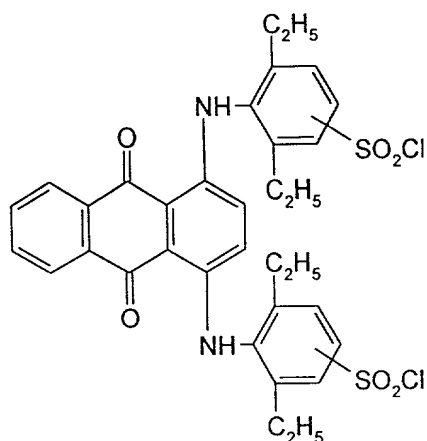
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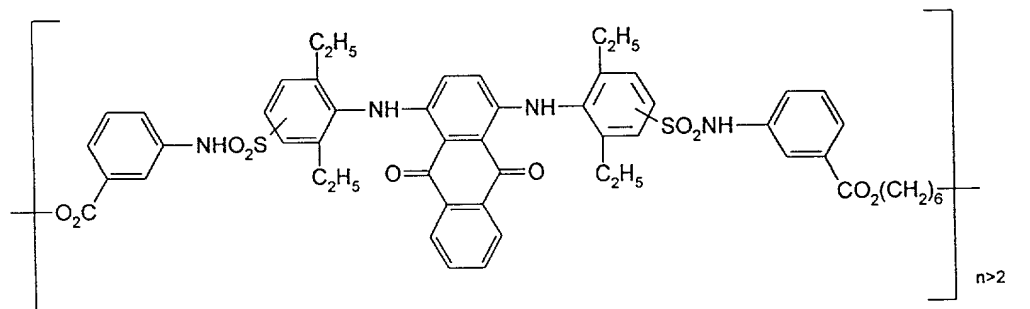
15

## 25

The anthraquinone disulfonyl chloride compound (3.50 g, 0.005 mole) having the following structure



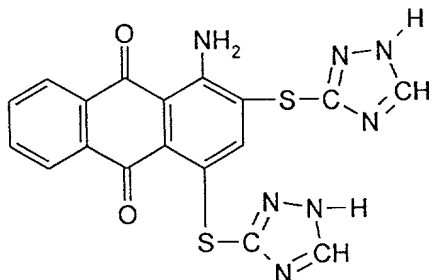
(prepared according to the procedure of U. S. Patent 5,453,482, Example 2), m-aminobenzoic acid (1.37 g, 0.10 mole), potassium carbonate (2.80 g) and DMF (30 mL) were mixed and the reaction mixture heated at 90-95°C for 30 minutes. TLC (50:50 THF:cyclohexane) indicated complete reaction of the disulfonyl chloride to produce the disulfonamide derivative. To the reaction mixture were added 1,6-hexanediol, dimethanesulfonate (1.38 g, 0.005 m), potassium carbonate (1.38 g) and heating and stirring were continued for 2.0 hours at 90-95°C. The reaction mixture was drowned into water and acidified with acetic acid. The bright blue polydye was collected by filtration, washed with water and then air dried (yield - 2.07 g) and is believed to have the following repeat unit:



GPC analysis indicated a weight average molecular weight of 5,252, a number average molecular weight of 2,179 and a polydispersity of 2.41. Absorption maxima at 583 nm and 628 nm were observed in the visible light absorption spectrum in DMF.

#### Example 18

A mixture of the diacidic anthraquinone compound (4.21 g, 0.01 mole) having the following structure



1,2-ethanediol, dimethanesulfonate (2.18 g, 0.01 mole), potassium carbonate (2.68 g, 0.02 mole) and DMF (50 mL) was heated and stirred at 90-95°C for 1.5 hours. The reaction mixture was drowned into water (400 mL) and acidified with stirring and by adding acetic acid. After being heated to about 50°C, the mixture was filtered and the red polydye washed well with water and dried in air (yield - 4.47 g). GPC analysis showed the polydye to have a weight average molecular weight of 1,603, a number average molecular weight of 922 and a polydispersity of 1.74. An absorption maximum at 524 nm was observed in the visible light absorption spectrum in DMF.

#### Example 18a

A mixture of 1-amino-2,4-dibromoanthraquinone (11.43 g, 0.03 mole), 3-mercapto-1(H)-1,2,4-triazole (9.09 g, 0.09

mole), potassium carbonate (11.52 g, 0.09 mole) and DMF (150 mL) was heated at about 95°C with stirring for 1.0 hour. The reaction mixture was drowned into water (500 mL) with stirring and acidified with acetic acid and the  
5 red product collected by filtration, washed with water and dried in air (yield - 12.64 g). FDMS indicated the product to be the diacidic anthraquinone compound used in Example 18.

10 Example 19

A mixture of 1,5-bis-(4-hydroxyphenylthio)anthraquinone (4.56 g, 0.01 mole), 1,2-ethanediol, dimethanesulfonate (2.18 g, 0.01 mole), potassium carbonate (3.0 g) and DMF  
15 (50 mL) was heated and stirred at about 95°C for 2.0 hours. The reaction mixture was drowned into methanol (100 mL) and the yellow polydye was collected by filtration and washed with methanol. The methanol-wet cake was reslurried in water (500 mL) and acidified and  
20 the polydye then collected by filtration, washed with water and dried in air (yield - 4.25 g). GPC analysis indicated the polydye to have a weight average molecular weight of 1,901, a number average molecular weight of 1,588 and a polydispersity of 1.20. An absorption maximum  
25 at 461 nm was observed in the visible light absorption spectrum in DMF.

Example 19a

30 A mixture of 1,5-dichloroanthraquinone (5.54 g, 0.02 mole), 4-hydroxybenzenethiol (6.30 g, 0.05 mole), potassium carbonate (6.90 g, 0.05 mole) and DMF (100 mL) was heated at about 95°C for 5.0 hours. The reaction mixture was drowned into water (400 mL) and the yellow  
35 product was collected by filtration, washed with water and dried in air (yield - 9.0 g). The solid was added to

acetic acid (150 mL) and the mixture heated to boiling. After being allowed to cool, the yellow solid was collected by filtration, washed with acetic acid and dried in air (yield - 6.75 g). FDMS confirmed that the product  
5 was the 1,5-bis(4-hydroxyphenylthio)anthraquinone used in Example 19.

Example 20

10 A mixture of 1,4-bis-(2-carboxyphenylthio)anthraquinone (1.53 g, 0.003 m), 1,2-ethanediol, dimethanesulfonate (0.66 g, 0.003 mole), potassium carbonate (0.75 g) and DMF (8.0 mL) was heated at about 95°C with occasional stirring for 2.0 hours. The reaction mixture was then drowned into  
15 methanol (100 mL) and the dark orange polydye was collected by filtration, washed with water containing some acetic acid then with hot water and dried in air (yield - 0.50 g). GPC analysis indicated a weight average molecular weight of 8,686, a number average molecular  
20 weight of 1,356 and a polydispersity of 6.41.

Example 20a

A mixture of 1,4-dichloroanthraquinone (2.77 g, 0.01  
25 mole), thiosalicylic acid (3.85 g, 0.025 m), potassium carbonate (3.45 g, 0.025 m), cupric chloride dihydrate (0.1 g) and DMF (50 mL) was heated at 95-100°C with stirring for 4.0 hours. The reaction mixture was drowned into acetone and the solid was collected by filtration and  
30 washed with acetone. The resulting potassium salt of the product was dissolved by stirring in water (200 mL). The red solution was neutralized to give the orange product which was collect by filtration, washed with water and dried in air (yield - 4.58 g). FDMS indicated the  
35 structure to be that of the starting material for Example

20. An absorption maximum at 501 nm was observed in the visible light absorption spectrum.

Example 21

5

A mixture of 1,8-bis-(2-carboxyphenylthio)-4,5-bis-(p-tolylthio)anthraquinone (1.51 g, 0.002 mole), 1,4-butanediol, dimethanesulfonate (0.49 g, 0.002 mole), potassium carbonate (0.60 g and DMF (8.0 mL) was heated at  
10 90-95°C with occasional stirring for 2.5 hours. The reaction mixture was drowned into methanol (100 mL) and the red polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and then dried in air (yield - 1.1 g). GPC analysis indicated a  
15 weight average molecular weight of 2,157, a number average molecular weight of 1,111 and a polydispersity of 1.94. An absorption maximum was observed at 529 nm in the visible light absorption spectrum in DMF.

20 Example 21a

A mixture of thiosalicyclic acid (4.75 g, 0.03 mole), potassium carbonate (8.70 g, 0.06 mole) and DMF (75 mL) was heated at about 100°C for 1.0 hour and the reaction  
25 mixture, which was allowed to cool, was added at 0-5°C to a solution of 1,8-dichloro-4,5-dinitroanthraquinone (5.51 g, 0.015 mole) dissolved in DMF (150 mL) with good stirring. Cooling was removed and the temperature of the reaction mixture allowed to come to ambient temperature  
30 and the mixture was stirred for about 3.0 hours. A solution of p-thiocresol (3.73 g, 0.03 mole) dissolved in DMF (80 mL) was added to the reaction mixture with stirring and the temperature raised to about 100°C and held for 2.0 hours. After allowing to cool, the reacting  
35 mixture was drowned into water (300 mL) and the mixture gradually acidified by the addition of 10% aqueous

hydrochloric acid. The red solid product was collected by filtration, washed with water and dried in air (yield - 11.28 g). FDMS analysis indicated that the product consisted mostly of the starting material for Example 21.

5

Example 22

00634T-13900  
A mixture of 1,5-bis(2-carboxyphenylthio)anthraquinone (1.54 g, 0.003 mole), 1,5-bis(2-carboxyphenylthio)-4,8-bis(isobutylamino)anthraquinone (1.31 g, 0.002 mole) (product of Example 2a), 1,2-ethandiol, dimethanesulfonate (1.09 g, 0.005 mole), potassium carbonate (1.0 g) and DMF (10 mL) was heated at 90-95°C with occasional stirring for 2.0 hours. The reaction mixture was drowned into methanol (100 mL) and the green polydye was washed with methanol, water containing acetic acid, hot water and then dried in air (yield - 1.30 g). GPC analysis indicated a weight average molecular weight of 1,839, a number average molecular weight of 1,040 and a polydispersity of 1.77. Absorption maxima were observed in the visible light absorption spectrum in DMF at 448, 603, and 645 nm.

Example 23

25 A mixture of 1,5-bis(2-carboxyphenylthio)anthraquinone (1.28 g, 0.0025 mole), 1,4-cyclohexanedimethanol, dimethanesulfonate (1.75 g, 0.0025 mole), potassium carbonate (0.82 g) and DMF (7.5 mL) was heated at about 95°C with occasional stirring for 3.0 hours. The reaction mixture was drowned into methanol (100 mL) and the yellow polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and then dried in air (yield - 0.31 g). GPC analysis indicated a weight average molecular weight of 1,158, a number average molecular weight of 1,008 and a polydispersity of 1.15.



Example 24

Example 23 was repeated except that the disulfonate used was 1,3-propanediol, 2,2-dimethyl, dimethanesulfonate  
5 (0.65 g, 0.0025 mole) to give the yellow polydye (yield - 0.76 g) which had a weight average molecular weight of 1,056, a number average molecular weight of 979 and a polydispersity of 1.08 by GPC analysis.

10 Example 25

Example 23 was repeated except that 1,6-hexanediol, dimethanesulfonate (0.68 g, 0.0025 mole) was used as the disulfonate to give the yellow polydye (yield - 1.16 g)  
15 which had a weight average molecular weight of 1,827, a number average molecular weight of 961 and a polydispersity of 1.90 by GPC analysis.

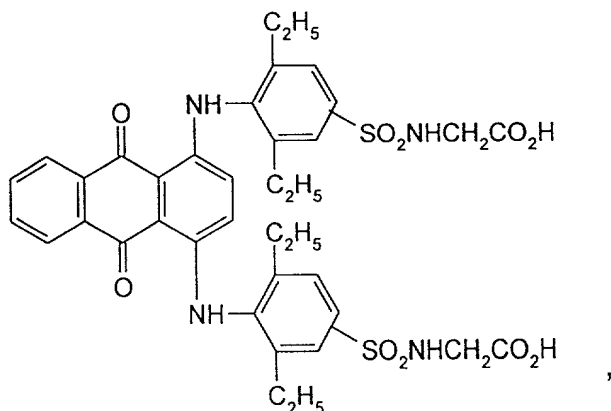
Example 26

20

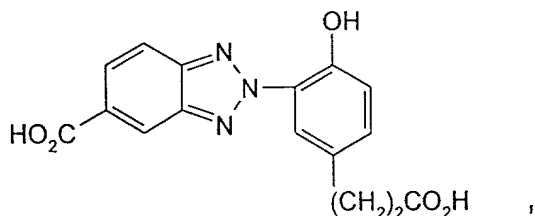
Example 23 was repeated except that 1,2-ethanediol, bis(4-methylbenzenesulfonate (0.82 g, 0.0025 mole) was used as the disulfonate to yield the yellow polydye (yield - 0.41 g) which had a weight average molecular weight of 2,442, a  
25 number average molecular weight of 1,885 and a polydispersity of 1.29 by GPC analysis.

Example 27

30 A mixture of the acidic anthraquinone compound (2.02 g, 0.0027 mole) having the structure



the acidic UV light absorbing compound (0.29 g,  $9 \times 10^{-4}$  mole) having the structure

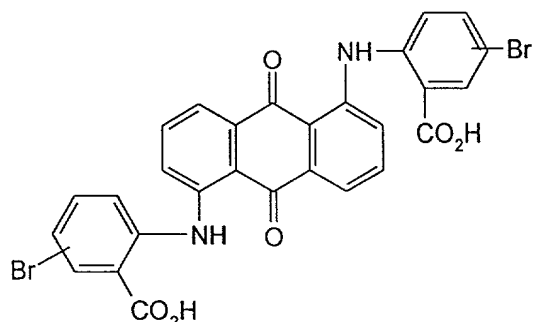


5

1,2-ethanediol, dimethanesulfonate (0.78 g, 0.0036 mole), potassium carbonate (1.0 g) and DMF (25 mL) was heated and stirred at 90-95°C for 2.0 hours. The cooled reaction mixture was drowned into water (200 mL) and made slightly acidic by the addition of acetic acid with stirring. The polymeric product was collected by filtration, washed well with water and dried in air (yield - 2.00 g). GPC analysis indicated a weight average molecular average of 5,642, a number average molecular weight of 1,720 and a polydispersity of 3.28.

#### Example 28

A mixture of the diacidic anthraquinone compound (1.27 g, 0.002 mole) having the structure



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00627F-99475460

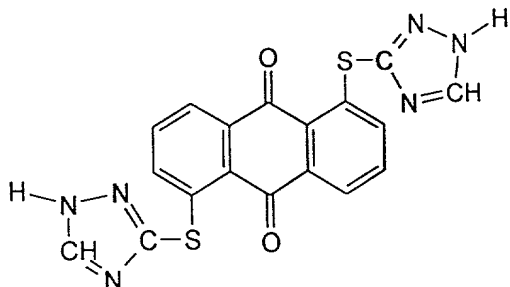
1,2-ethanediol, dimethanesulfonate (0.44 g, 0.002 mole), potassium carbonate (0.75 g) and DMF (8.0 mL) was heated at 90-95°C with occasional stirring for 2.0 hours. The reaction mixture was drowned into methanol (100 mL) and the dark red polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and then dried in air (yield - 1.23 g). GPC analysis indicated a weight average molecular weight of 1,545, a number average molecular weight of 1,213 and a polydispersity of 1.27.

#### Example 28a

To a mixture of 1,5-bis(2-carboxyanilino)anthraquinone (9.57 g, 0.02 mole) in DMF (250 mL) was added portionwise N-bromosuccinimide (7.12 g, 0.04 mole) with stirring at room temperature. The reaction mixture was then heated at about 60°C for 1.5 hours and allowed to cool. Water was added dropwise to precipitate the product, which was collected by filtration, washed with water and dried in air (yield - 11.17 g). FDMS indicated the structure of the product to be that of the starting anthraquinone compound in Example 28.

#### Example 29

A mixture of the diacidic anthraquinone compound (4.06 g, 0.01 mole) having the structure



00527 0945260  
1,2-ethanediol, dimethanesulfonate (2.18 g, 0.01 mole),  
potassium carbonate (2.76 g) and DMF (150 mL) was heated  
5 at about 100°C for 3.0 hours. The reaction mixture was  
drowned into water, acidified with acetic acid and the  
yellow polydye was collected by filtration, washed with  
water and dried in air. GPC analysis indicated a weight  
average molecular weight of 5,333, a number average  
10 molecular weight of 2,441, and a polydispersity of 2.18.

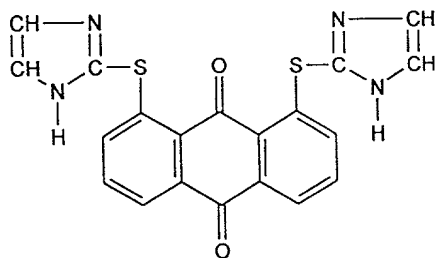
#### Example 29a

A mixture of 1,5-dichloroanthraquinone (6.93 g, 0.025  
15 mole), 3-mercapto-1(H)-1,2,4-triazole (5.56 g, 0.055  
mole), potassium carbonate (6.91 g, 0.05 mole) and DMF  
(100 mL) was heated and stirred at about 95°C for 5.0  
hours. The mixture was drowned into water and the yellow  
product was collected by filtration, washed with water and  
20 air dried. The cake was reslurried in hot isopropanol and  
the product collected by filtration, washed with  
isopropanol and dried in air (yield 8.62 g). FDMS  
indicated the product to be 1,5-bis[1(H)-1,2,4-triazol-3-  
ylthio]anthraquinone used as the diacidic anthraquinone  
25 starting material in Example 29.

Example 30

A mixture of diacidic anthraquinone compound (1.01 g, 0.0025 mole) having the structure

5



1,2-ethanediol, dimethanesulfonate (0.55 g, 0.0025 mole), potassium carbonate (0.75 g) and DMF (10 mL) was heated at about 95°C for 3.0 hours. The reaction mixture was then  
10 drowned into methanol (100 mL) and the yellow polydye was collected by filtration, water containing acetic acid, hot water and then air dried (yield - 0.35 g). GPC analysis indicated a weight average molecular weight of 2,478, a number average molecular weight of 742 and a  
15 polydispersity of 3.34. An absorption maximum was observed in the visible light absorption spectrum at 425 nm in DMF.

Example 30a

20

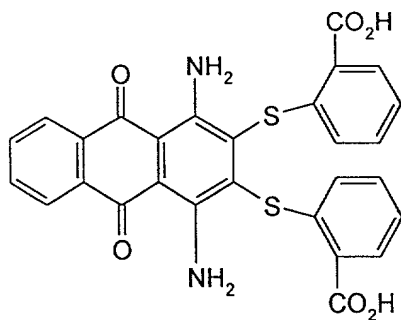
A mixture of 1,8-dichloroanthraquinone (6.93 g, 0.025 mole), 2-mercaptoimidazole (5.01 g, 0.05 mole), potassium carbonate (6.91 g) and DMF (60 mL) was heated and stirred at about 95°C for 8.0 hours. The reaction mixture was  
25 drowned into water and acidified using acetic acid. The yellow product was collected by filtration, washed with water and dried in air. FDMS indicated the product to be the 1,8-bis(imidazol-2-ylthio) anthraquinone diacidic compound used as the starting material in Example 30.

Example 31

A mixture of 1,5-bis[1(H)-1,2,4-triazol-3ylthio]  
5 anthraquinone (1.80 g, 0.00443 mole) (product of Example  
29a), 1,4-dibromobutane (0.96 g, 0.00444 mole),  
tributylamine (1.64 g, 0.00885 mole), and N-methyl-2-  
pyrrolidinone (30 mL) was heated at 8.0 hours at about  
130°C with stirring. The reaction mixture was drowned  
10 into acetone (150 mL) and the yellow polydye was collected  
by filtration, washed with acetone until filtrate was  
essentially clear and dried in air. GPC analysis  
indicated a weight average molecular weight of 5,022, a  
number average molecular weight of 3,220 and a  
15 polydispersity of 1.56.

Example 32

A mixture of the diacidic anthraquinone compound (1.63 g,  
20 0.003 mole) having the structure



1,6-hexanediol, dimethanesulfonate (0.82 g, 0.003 mole),  
potassium carbonate (0.5 g) and DMF (8.0 mL) was heated at  
25 about 95°C with occasional stirring for 2.0 hours. The  
mixture was drowned into methanol (100 mL) and the dark  
blue polydye was collected by filtration, washed with  
methanol, water containing acetic acid, hot water and

dried in air (yield - 0.92 g). Absorption maxima at 602 and 644 nm were observed in the visible light absorption spectrum in DMF. GPC analysis indicated a number average molecular weight of 1,860.

5

Example 32a

0036176-13900  
A mixture of 1,4-diamino-2,3-dichloroanthraquinone (12.24 g, 0.04 mole), thiosalicylic acid (15.4 g, 0.10 mole),  
10 potassium carbonate (13.8 g, 0.10 mole) and DMF (150 mL) was heated at about 95°C with stirring for 2.0 hours. A bathochromic shift in color from violet to blue was observed as the reaction progressed. The reaction mixture was drowned into acetone (500 mL) and the solid product  
15 was collected by filtration and washed well with acetone. The acetone-wet cake was added to water (600 mL) and the mixture acidified with acetic acid to precipitate the free acid compound, which was collected by filtration, washed with water and dried in air (yield - 21.4 g). FDMS  
20 indicated the product to be the 1,4-diamino-2,3-bis(2-carboxyphenylthio) anthraquinone used in Example 32.

Example 33

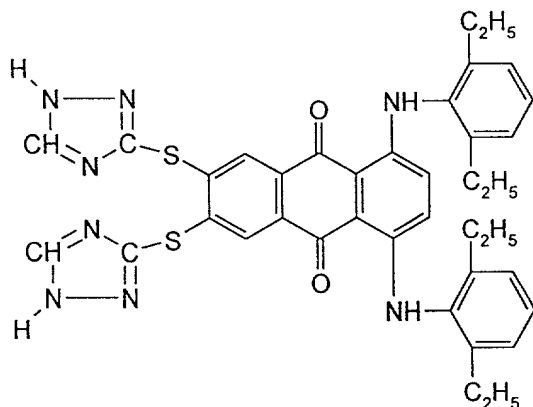
25 A mixture of 1,5-bis(2-carboxyphenylthio) anthraquinone (1.02 g, 0.002 mole), terephthalic acid (1.00 g, 0.006 mole), potassium carbonate (1.38 g) 1,2-ethanediol, dimethanesulfonate (1.74 g, 0.008 mole) and DMF (10 mL) was heated at about 95°C with occasional stirring for 2.0  
30 hours. The mixture was then drowned into methanol (100 mL) and the yellow polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 1.88 g). GPC analysis indicated a weight average molecular weight of 794, a  
35 number average molecular weight of 713 and a polydispersity of 1.11.

Example 34

Example 33 was repeated using 1,5-bis(2-carboxyphenylthio)  
5 anthraquinone (1.02 g, 0.002 mole) and terephthalic acid  
(0.33 g, 0.002 mole), 1,2-ethanediol, dimethanesulfonate  
(0.87 g, 0.004 mole) and potassium carbonate (0.87 g) to  
yield the yellow polydye (0.90 g). GPC analysis indicated  
10 a weight average molecular weight of 875, a number average  
molecular weight of 811, and a polydispersity of 1.08.

Example 35

A mixture of the diacidic anthraquinone compound (2.00 g,  
15 0.00285 mole) having the following structure (Preparation  
5 of IR Docket 70351):



1,2-ethanediol, dimethanesulfonate (0.63 g, 0.00289 mole),  
20 potassium carbonate (0.80 g) and DMF (25 mL) was heated at  
95°C for 4.0 hours with stirring. The reaction mixture  
was drowned into methanol (100 mL) and the greenish-blue  
polydye was collected by filtration, washed with methanol,  
water containing acetic acid, hot water and dried in air  
25 (yield - 1.01 g). GPC indicated a weight average  
molecular weight of 6,720, a number average molecular

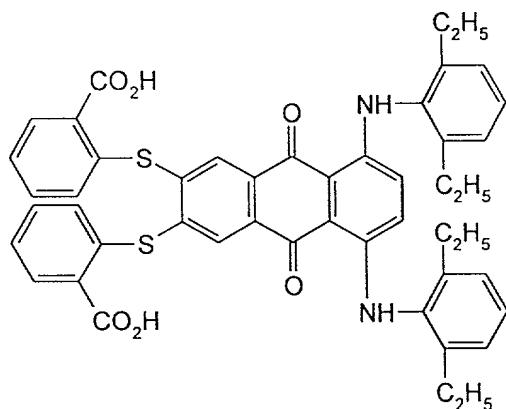


weight of 2,211 and a polydispersity of 3.04. Absorption maxima were observed at 599 and 647 nm in the visible absorption spectrum in DMF.

5 Example 36

A mixture of the diacidic anthraquinone compound (0.41 g, 0.508 mmole) having the following structure (Preparation 4 in IR Docket 70351):

10



1,2-ethanediol, dimethanesulfonate (0.11 g, 0.504 mmole), potassium carbonate (0.14 g) and DMF (5.0 mL) was heated with occasional stirring or about 95°C for 3.0 hours. The reaction mixture was drowned into methanol (50 mL) and the greenish-blue polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield 0.15 g). Absorption maxima were observed at 599 and 645 nm in the visible light absorption spectrum in DMF.

Example 37-66

Colored EASTAR<sup>®</sup> copolyester 6763 film was produced by melt blending the polydyes of Examples 7-36 and extruding

according to the following procedure to produce Examples 37-66 (Table 1).

EASTAR® PETG polyester 6763, a poly(ethylene-1,4-cyclohexanedimethylene) terephthalate (Eastman Chemical Company) (300 g of previously dried pellets) was dry blended with the anthraquinone polydye composition (0.12 g). The blend was extruded with a C. W. Brabender  $\frac{3}{4}$  in. extruder, equipped with a mixing screw, at 250°C into a water bath and the extrudate pelletized.

The pellets were redried at 70°C for 17 hrs. at a pressure of about 1-5 torr. A portion (1.40 g) of the dried pellets was pressed into a 18-20 mil film at 250°C using a 2-inch diameter circular mold in a Pasadena Hydraulic, Inc. press using 12,000 pounds ram force (4 inch ram). The transparent films contained about 300 ppm of the polydyes and each showed excellent color development to produce the colors indicated in Table 1.

#### Example 67

A mixture of 1,4-bis(2-carboxyphenylthio)anthraquinone (15.4 g, 0.03 mole) (prepared as in Example 20a), 1,5-bis(2-carboxyphenylthio)-4,8-bis(isobutylamino)anthraquinone (6.55 g, 0.01 mole) (Example 2a), 1,2-ethanediol, dimethanesulfonate (8.72 g, 0.04 mole), potassium carbonate (8.0 g) and DMF (100 mL) was stirred and heated at about 95°C for 2.0 hours with occasional stirring. The reaction mixture was drowned into methanol (500 mL) and the black polydye was collected by filtration, washed with water containing acetic acid, hot water and dried in air (yield - 9.5 g). GPC analysis indicated a weight average molecular weight of 7,512, a number average molecular weight of 1,700 and a polydispersity of 4.42.

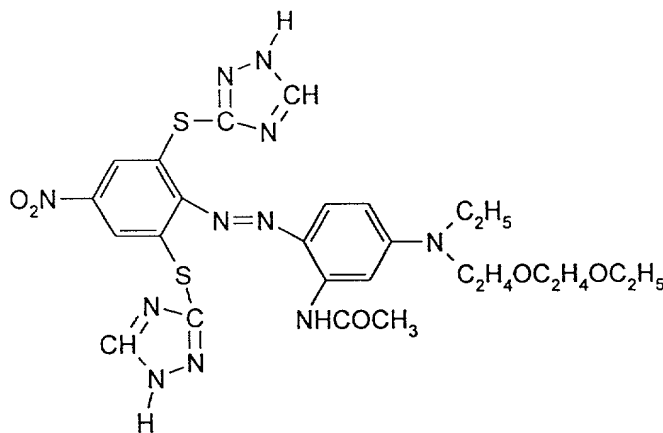
Example 68

5 EASTAR® PETG copolyester 6763 (291 g of previously dried pellets) was dry blended with the black polydye of Example 67 (9.0 g) and the blend extruded and a portion of the resulting pellets was pressed into a black film containing approximately 3.0% by weight of polydye by using the procedure described in Example 4.

10

Example 69

A mixture of the diacidic azo compound (3.20 g, 0.005 mole) having the structure



15

1,2-ethanediol, dimethanesulfonate (1.09 g, 0.005 mole), potassium carbonate (1.5 g) and DMF (25 mL) was heated and stirred at about 95°C for 3.0 hours. The reaction mixture was drowned into methanol and the violet polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 1.60 g). GPC analysis indicated a weight average molecular weight (Mw) of 6,403, a number average molecular weight (Mn) of 3,700 and a polydispersity (Mw/Mn) of 1.73.

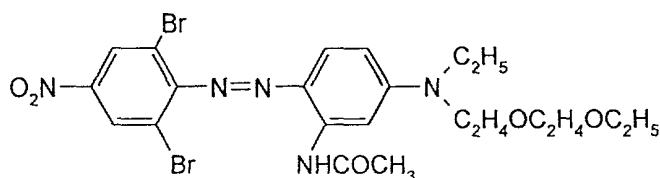
20

In the visible light absorption spectrum in DMF an absorption maximum was observed at 556 nm.

Example 69a

5

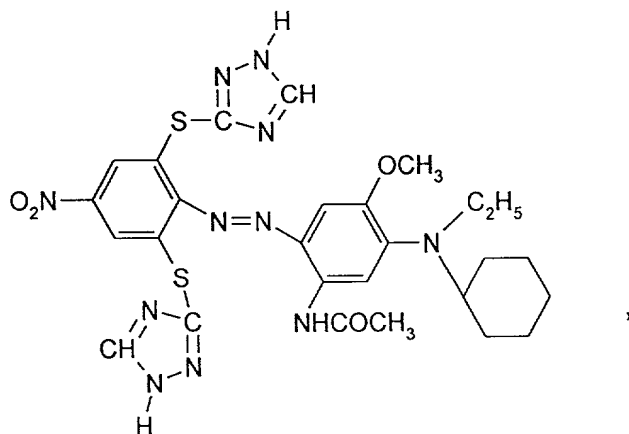
A mixture of the dibromoazobenzene dye (6.01 g, 0.010 mole) having the structure



3-mercapto-1(H)1,2,4-triazole (2.2 g, 0.022 mole),  
10 potassium carbonate (3.45 g, 0.025 mole) and DMF (100 mL)  
was stirred and heated at about 95°C for 2.0 hours. TLC  
(75 parts THF: 25 parts cyclohexane) showed incomplete  
reaction. An additional quantity (1.01 g, 0.01 m) 3-  
mercapto-1(H)-1,2,4-triazole was added and heating and  
15 stirring were continued for 2.0 additional hours. TLC  
indicated essentially complete reaction to produce the  
violet product. The reaction mixture was drowned into  
water (400 mL) and the mixture was acidified by addition  
of acetic acid, heated to about 40°C and filtered. The  
20 product was washed with warm water and dried in air (yield  
- 5.60 g). FDMS indicated the product to have the  
structure of the diacidic azobenzene compound used in  
Example 69.

25 Example 70

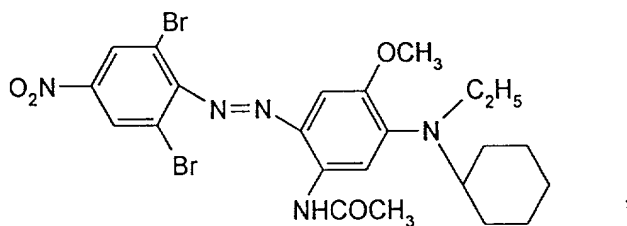
A mixture of the diacidic azo compound (1.59 g, 0.0025 mole) having the structure



1,2-ethanediol, dimethanesulfonate (0.55 g, 0.0025 mole), potassium carbonate (0.5 g) and DMF (8.0 mL) was heated at 95°C with occasional stirring for 3.0 hours. The reaction mixture was drowned into methanol (100 mL) and the blue polydye product was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 1.06 g). GPC analysis indicated a Mw of 5,497, a Mn of 2,648 and a Mw/Mn of 2.08. An absorption maximum was observed at 605 nm in DMF in the visible light absorption spectrum.

#### Example 70a

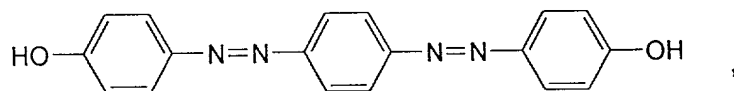
A mixture of the dibromo azobenzene dye (2.38 g, 0.004 mole) having the structure



3-mercapto-1(H)-1,2,4-triazole (1.21 g, 0.012 mole),  
potassium carbonate (1.65 g, 0.012 mole) and DMF (25 mL)  
was heated and stirred for 1.0 hour. TLC (50 parts THF:50  
parts cyclohexane) showed complete reaction to produce the  
5 product. The reaction mixture was drowned into water (100  
mL) and the mixture acidified with acetic acid. The dark  
blue product was collected by filtration, washed with  
water and dried in air (yield - 2.55g). FDMS indicated  
the product to have the structure of the diacidic  
10 azobenzene compound used in Example 70.

Example 71

A mixture of the diacidic disazo compound (1.59 g, 0.005  
15 mole) having the structure



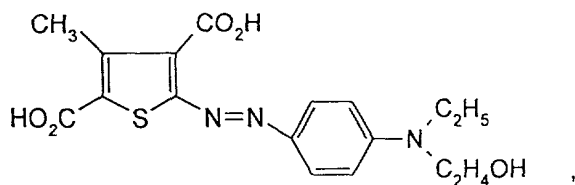
1,2-ethanediol, dimethanesulfonate (1.09 g, 0.005 mole),  
potassium carbonate (1.5 g), DMF (10 mL) was heated and  
stirred at about 95°C for 3.0 hours. The reaction mixture  
20 was drowned into methanol (100 mL) and the dark brown  
polydye was collected by filtration, washed with methanol,  
water containing acetic acid, hot water and then dried in  
air (yield - 0.66 g). GPC analysis indicated a Mw of  
4,926, a Mw of 1,574 and a Mw/Mn of 3.13.

25

Example 72

A mixture of the diacidic azo compound (1.88 g, 0.005  
mole) having the structure

30



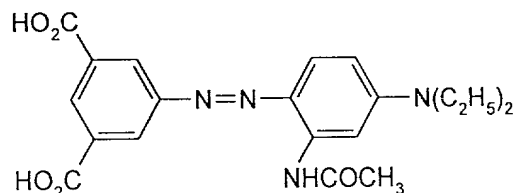
005331 99475260  
1,2-ethanediol, dimethanesulfonate (1.09 g, 0.005 mole),  
potassium carbonate (1.5 g) and DMF (20 mL) was heated at  
about 95°C with stirring for 3.0 hours. The reaction  
5 mixture was drowned in methanol (100 mL) and the red  
polydye was collected by filtration, washed with methanol,  
water containing acetic acid, hot water and dried in air  
(yield - 1.35 g). GPC analysis indicated a Mw of 6,888, a  
Mn of 2,127 and a Mw/Mn of 3.24. An absorption maximum  
10 was observed at 527 nm in the visible light absorption  
spectrum in DMF.

#### Example 72a

15 To a stirred mixture of the azo compound (4.05 g, 0.01  
mole) [4-(3',5'-dicarbomethoxy-4'-methylthiophene-2-  
ylazo)-N-ethyl-N(2-hydroxyethyl)aniline] and 2-  
ethoxyethanol (50 mL) at room temperature was added  
aqueous 50% NaOH solution(3.75 g). After being heated at  
20 about 95°C for 1.0 hour, the reaction product was drowned  
into acetone (300 mL). The disodium salt of the diacidic  
azo dye was collected by filtration washed with acetone  
and then quickly dissolved in water (200 mL).  
Acidification with acetic acid precipitated the free  
25 diacid dye, which was collected by filtration, washed with  
water and dried in air (yield - 2.35 g). FDMS indicated  
the product to have the structure of the diacidic azo  
compound used in Example 72.

Example 73

A mixture of the diacidic azobenzene compound (1.19 g, 0.003 mole) having the structure



5

1,2-ethanediol, dimethanesulfonate (0.66 g, 0.003 mole), potassium carbonate (0.75 g), and DMF (8.0 mL) was stirred occasionally and heated at about 95°C for 2.0 hours. The reaction mixture was drowned into methanol (100 mL) and the orange polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 0.65 g). GPC analysis showed a Mw of 3,015, a Mn of 2,128 and a Mw/Mn of 1.42. An absorption maximum was observed in the visible light absorption at 479 nm in DMF.

15

Example 73a

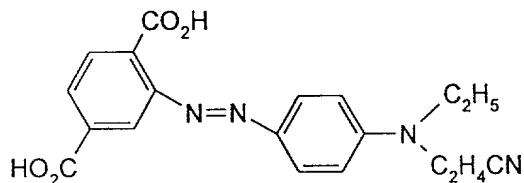
To a mixture of 3-acetamido-4-(3',5'-dicarbomethoxyphenylazo)-N,N-diethylaniline (1.7 g, 0.004 mole) in 2-ethoxyethanol (20 mL) was added aqueous 50% NaOH (1.6 g). The reaction mixture was heated with stirring of 95°C for 10 minutes and then drowned into water (100 mL). The solution was acidified with acetic acid to precipitate the diacid dye which was collected by filtration, washed with water and dried in air (yield - 1.6 g). FDMS indicated the structure to be that of the starting diacid azobenzene compound in Example 73.

30



Example 74

A mixture of the diacidic azobenzene compound (1.10 g, 0.003 mole) having the structure



5  
1,6-hexanediol, dimethanesulfonate (0.82 g, 0.003 mole), potassium carbonate (0.45 g) and DMF (8.0 mL) was heated at 95°C with occasional stirring for 2.0 hours. The  
10 reaction mixture was drowned into methanol (100 mL). A slightly sticky yellow product resulted. The methanol was removed by decantation and the product dissolved in DMF (10 mL) by heating and stirring. Water (100 mL) was added and the mixture acidified by addition of acetic acid. The  
15 solid yellow polydye was collected by filtration, washed with water and dried in air (yield - 0.47 g). GPC analysis indicated a Mw of 9,314, a Mn of 3,208 and a Mw/Mn of 2.90. An absorption maximum at 428 nm was observed in the visible light absorption spectrum in DMF.

20

Example 74a

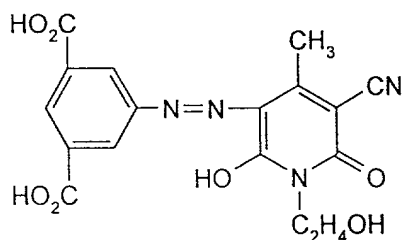
To a mixture of 4-(2',5'-dicarbomethoxyphenylazo)-N-(2-cyanoethyl)-N-ethylaniline (1.97 g, 0.005 mole) in 2-  
25 ethoxyethanol (20 mL) was added aqueous 50% NaOH (1.90 g). The reaction solution was heated at 95°C for 15 minutes and then drowned into water (200 mL). The solution was acidified and the yellow dye which precipitated was collected by filtration, washed with water and dried in

air (yield - 1.75 g). FDMS indicated the structure to be that of the starting diacid azobenzene dye of Example 74.

Example 75

5

A mixture of diacidic azo compound (38.6 g, 0.10 mole) having the structure



- 10 1,6-hexanediol, dimethanesulfonate (27.4 g, 0.10 mole), potassium carbonate (27.6 g, 0.20 mole), and DMF (350 mL) was heated at 95-100°C for 2.0 hours. The reaction mixture was drowned into a solution of acetic acid (70.0 mL) in water (1700 mL) with good stirring. After stirring
- 15 for about 15 minutes, the yellow polydye was collected by filtration, washed with hot water and dried in air (yield - 42.6 g). An absorption maximum at 422 nm was observed in the visible light absorption spectrum in DMF.

20 Example 75a

- To a mixture of the diester dye (41.4 g, 0.10 mole) [3-cyano-5-(3',5'-dicarbomethoxyphenylazo)-6-hydroxy-N-(2-hydroxyethyl)-4-methyl-2-pyridone] in 2-ethoxyethanol (400
- 25 mL) was added aqueous 50% NaOH (40.0 g) and the reaction mixture was heated at 75-80°C for about 30 minutes. Acetone (200 mL) was added to the slightly cooled reaction mixture. The yellow solid was collected by filtration, washed with acetone and then reslurried in warm water (750
- 30 mL). After acidification using conc. HCl (20 mL), the

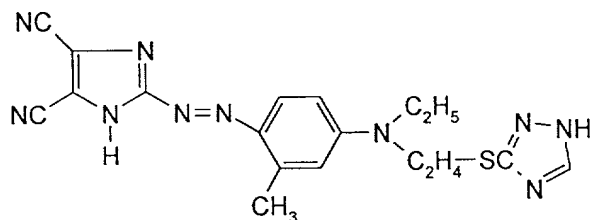
yellow diacid dye was collected by filtration, washed with hot water and dried in air (yield- 36.0 g). FDMS indicated the structure to be that of the starting diacid azo compound of Example 75.

5

Example 76

A mixture of the diacidic azo compound (2.03 g, 0.005 mole) having the structure

10



1,2-ethanediol, dimethanesulfonate (1.09 g, 0.005 mole), potassium carbonate (1.5 g) and DMF (20 mL) was heated at about 95°C with occasional stirring for 5.0 hours. The reaction mixture was drowned into methanol. Acetic acid (1.0 mL) was added and the polydye was collected by filtration and washed with water and dried in air. GPC analysis indicated a Mw of 9,876, a Mn of 3,917 and a polydispersity of 2.52. An absorption maximum at 506 nm was observed in the visible light absorption spectrum in DMF.

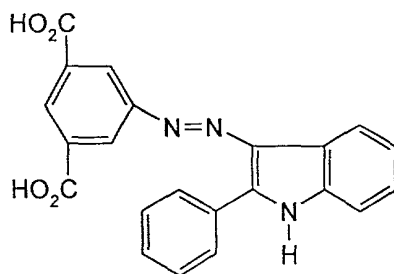
15

20

Example 77

A mixture of the diacidic azo compound (0.60 g, 0.00155 mole) having the structure

25



006237 992460  
1,2-ethanediol, dimethanesulfonate (0.34 g, 0.00155 mole),  
potassium carbonate (0.3 g) and DMF (4.0 mL) was heated at  
about 95°C for 4.0 hours. The reaction mixture was  
5 drowned into methanol (20 mL) and the yellow polydye was  
collected by filtration, washed with methanol, water  
containing acetic acid, water and then air dried (yield -  
0.5 g). GPC analysis showed a Mw of 4,566, a Mn of 2,474  
and a Mw/Mn of 1.84. In the visible light absorption  
10 spectrum in DMF an absorption maximum was observed at 420  
nm.

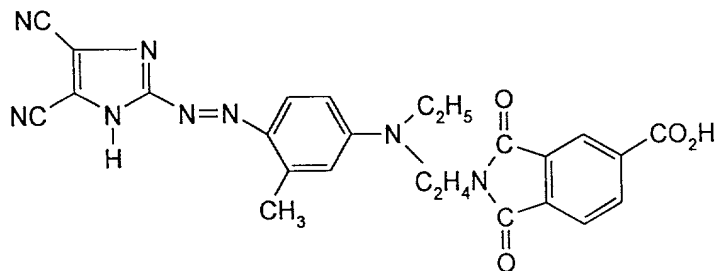
#### Example 77a

15 To a mixture of 3-(3',5'-dicarboxymethoxyphenylazo)-2-  
phenylindole (1.0 g, .00242 mole) in 2-ethoxyethanol (10  
mL) was added aqueous 50% NaOH (0.75 g) and the hydrolysis  
reaction carried out by heating at about 95°C for 30  
minutes. The reaction mixture was drowned into water (100  
20 mL) and the solution treated with acetic acid to  
precipitate the product which was collected by filtration,  
washed with water and dried in air (yield - 0.85 g). FDMS  
indicated the structure to be that of the starting  
diacidic azo compound in Example 77.

25

#### Example 78

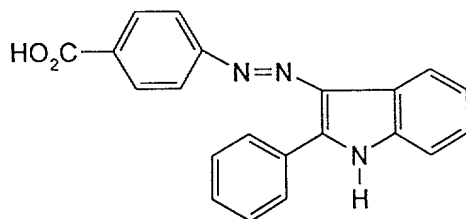
A mixture of the diacidic azo compound (0.99 g, 0.002  
mole) having the structure



1,2-ethanediol, dimethanesulfonate (0.42 g, 0.002 mole), potassium carbonate (0.5 g) and DMF (7.0 mL) was heated at about 95°C for 3.0 hours. The reaction mixture was drowned into methanol (50 mL) and the scarlet polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and then dried in air (yield - 0.18 g). GPC analysis indicated a Mw of 8,246, a Mn of 2,619 and a polydispersity of 3.15.

#### Example 79

A mixture of the diacidic azo dye (2.50 g, 0.00733 mole) having the following structure

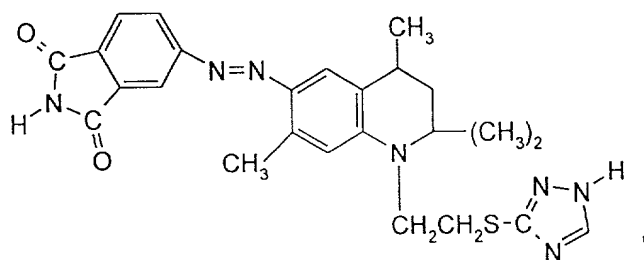


1,2-ethanediol, dimethanesulfonate (1.60 g, 0.00733 mole), potassium carbonate (2.07 g) and DMF (25 mL) was heated at 95°C for 3.0 hours. The reaction mixture was drowned into methanol and a small amount of acetic acid added. The yellow polydye was collected by filtration, washed with a little methanol, water containing acetic acid, hot water and dried in air. GPC analysis indicated a Mw of 1,949, a

Mn of 1,569 and a Mw/Mn of 1.24. An absorption maximum was observed at 411 nm the visible light absorption spectrum.

5 Example 80

A mixture of the diacidic azo compound (1.22 g, 0.0025 mole) having the structure



10

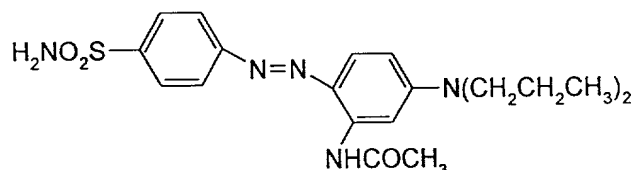
1,2-ethanediol, dimethanesulfonate (0.55 g, 0.0025 mole), potassium carbonate (0.75 g) and DMF (8.0 mL) was heated and stirred at about 95°C for 3 hours with occasional stirring. The reaction mixture was drowned into methanol (50 mL) and the polydye was collected by filtration washed with methanol, water containing acetic acid, hot water and then dried in air (yield - 0.68 g). GPC analysis indicated a Mw of 2,259, a Mn of 1,571 and a Mw/Mn of 1.44. An absorption maximum was observed at 503 nm in DMF in the visible light absorption spectrum.

20

Example 81

A mixture of the diacidic azo compound (1.25 g, 0.003 mole) having the structure

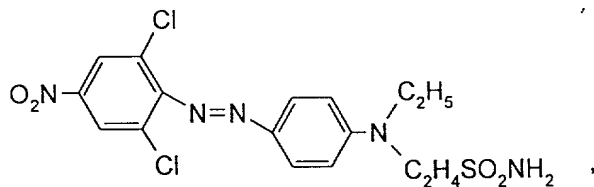
25



1,2-ethanediol, dimethanesulfonate (0.65 g, 0.003 mole), potassium carbonate (1.0 g) and DMF (10 mL) was heated at about 95°C for 3.0 hours with occasional stirring. The reaction mixture was drowned into methanol (25 mL) and the orange polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 0.75 g). GPC analysis indicated a Mw of 2,014, a Mn of 1,520 and a Mw/Mn of 1.32. An absorption maximum was observed at 493 nm in the visible light absorption spectrum in DMF.

#### Example 82

A mixture of the diacidic azo compound (1.11 g, 0.0025 mole) having the structure



1,2-ethanediol, dimethanesulfonate (0.55 g, 0.0025 mole), potassium carbonate (0.80 g) and DMF (8.0 mL) was heated at about 95°C for 2.5 hours. The reaction mixture was drowned into methanol (100 mL) and the brown polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 0.30 g). GPC analysis indicated a Mw of 2,301, a Mn of 1,345 a Mw/Mn of 1.71. In the visible light absorption

spectrum in DMF a maximum absorption was observed at 434 nm.

Example 83

5

A mixture of the diacidic azo compound (2.40 g, 0.005 mole) having the structure

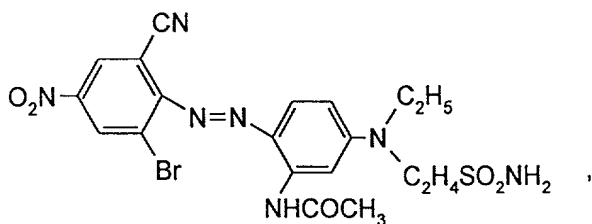


10 1,2-ethanediol, dimethanesulfonate (1.09 g, 0.005 mole), potassium carbonate (1.5 g) and DMF (25 mL) was heated at about 95°C for 3.0 hours with occasional stirring. The reaction mixture was drowned into methanol (200 mL) and the dark red polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and  
15 then dried in air (yield - 1.80 g). GPC analysis indicated Mw of 2,914, a Mn of 809 and a Mw/Mn of 3.60. An absorption maximum at 528 nm was observed in the visible light absorption spectrum in DMF.

20 Example 84

A mixture of the diacidic azo compound (1.07 g, 0.002 mole) having the structure





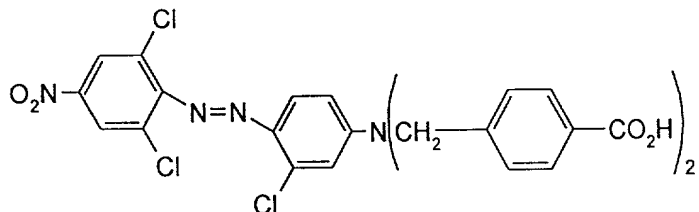
1,2-ethanediol, dimethanesulfonate (0.44 g, 0.002 mole), potassium carbonate (0.5 g) and DMF (10 mL) was heated at 95°C with occasional stirring for 5 hours. The reaction mixture was drowned into methanol (50 mL) and the reddish-blue polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 0.83 g). GPC analysis indicated a Mw of 7,038, a Mn of 832 and a Mw/Mn at 8.44. An absorption maximum was observed at 574 nm in the visible light absorption spectrum in DMF.

Example 85 - Displacement of Bromine in Polydye of Example 84 with Cyano Group

A mixture of a portion (0.5 g) of the polydye of Example 84, sodium dicyanocuprate (0.2 g) and DMF (8.0 mL) was heated at about 95°C with occasional stirring for 3.0 hours. The reaction mixture, the color of which changed from reddish-blue to neutral-blue as the displacement reaction occurred, was then drowned into methanol and the polydye was collected by filtration, washed with methanol and dried in air. GPC analysis indicated a Mw of 9,427, a Mn of 1,117 and a Mw/Mn of 8.44. An absorption maximum at 590 nm was observed in DMF in the visible light absorption spectrum.

Example 86

A mixture of diacidic azo compound (1.53 g, 0.0025 mole) having the structure



1,6-hexanediol, dimethanesulfonate (0.69 g, 0.0025 mole),  
K<sub>2</sub>CO<sub>3</sub> (0.8 g) and DMF (8.0 mL) was heated at about 95°C  
5 with occasional stirring for 2.0 hours. The reaction  
mixture was drowned into methanol (100 mL) and the brown  
polydye was collected by filtration, washed with methanol,  
water containing acetic acid, hot water and then dried in  
air (yield - 0.62 g). GPC analysis indicated a Mw of  
10 4,795, a Mn of 2,051 and a Mw/Mn of 2.33. An absorption  
maximum at 434 nm in DMF was observed in the visible light  
absorption spectrum.

#### Example 86a

15 To conc. H<sub>2</sub>SO<sub>4</sub> (33.0 mL) was added 2,6-dichloro-4-  
nitroaniline (6.21 g, 0.03 mole) with stirring. The  
solution was cooled to 0-5°C and stirred while a nitrosyl  
sulfuric acid mixture, prepared by adding sodium nitrite  
20 (2.19 g) to conc. H<sub>2</sub>SO<sub>4</sub> (15 mL) portionwise with stirring  
and allowing the temperature to rise, was added below 5°C  
with stirring. The diazotization reaction mixture was  
stirred at 0-5°C for 2.0 hours. An aliquot of the  
diazonium salt solution (0.01 mole) was added to a chilled  
25 solution of the diacid coupler (3.95 g, 0.01 mole) (N,N-  
bis(4-carboxyphenylmethyl)-3-chloroaniline) dissolved in  
1:5 (1 part propionic acid:5 parts acetic acid) (120 mL)  
containing some conc. HCl (5.0 mL) with stirring at 0-5°C.  
The coupling reaction mixture was neutralized by the  
30 addition of ammonium acetate with stirring and allowed to  
stand with occasional stirring at below 5°C for about 1.0

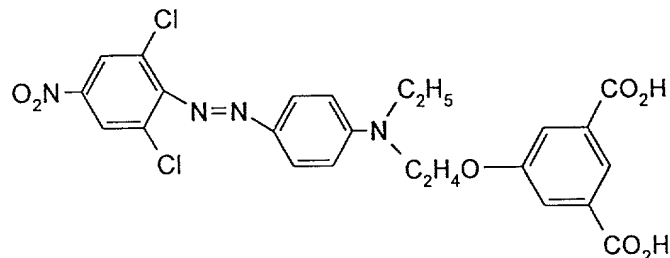
hour. Water was added to precipitate the solid dye, which was collected by filtration, washed with water and dried in air (yield - 4.0 g). The crude dye was reslurried in hot methanol and the mixture allowed to cool. The final dye was collected by filtration, washed with methanol and dried in air. An absorption maximum was observed at 431 nm in DMF. The diacid dye was used as the starting material in Example 86.

10 Example 86b

A mixture of m-chloroaniline (2.56 g, 0.02 mole), methyl 4-(bromomethyl)benzoate (10.08 g, 0.044 mole), sodium carbonate (4.66 g) and sodium iodide (0.2 g) and 2-ethoxyethanol (50 mL) was heated under nitrogen at about 90°C for 3.0 hours with stirring. The reaction mixture was drowned into water and the product was extracted into methylene chloride. Methylene chloride was removed to leave an oily product (11.0 g), which was added to 2-ethoxyethanol (100 mL). To the solution was added aqueous 50% NaOH solution (7.50 g) and the reaction mixture was warmed. At about 30°C, white solids began to precipitate and at about 50°C the reaction mixture become very thick. When the temperature had reached 70°C, water (20 mL) was added to dissolve the salts of the diacidic product. After stirring at 70°C for 1.5 hours the reaction mixture was clarified by filtering through Celite filter aid and the filtrate acidified by the addition of 10% aqueous HCl to pH of about 4.0. The white solid was collected by filtration, washed with water and dried in air (yield - 7.20 g). FDMS indicated the product to have the structure of the coupler used in Example 86a.

Example 87

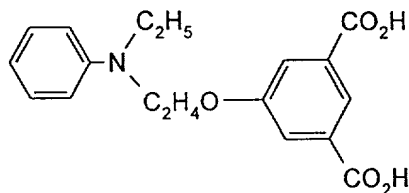
A mixture of the diacidic azo compound (1.64 g, 0.003 mole) having the structure



- 5 1,6-hexanediol, dimethanesulfonate (0.82 g, 0.003 mole), potassium carbonate (0.5 g) and DMF (8.0 mL) was heated at about 95°C for 25 hours with occasional stirring. The reaction mixture was drowned into methanol (150 mL) and
- 10 the polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 1.5 g). GPC analysis indicated a Mw of 2,741, a Mn of 1,367 and a Mw/Mn of 2.00. An
- 15 absorption maximum at 441 nm was observed in the visible light absorption spectrum in DMF.

Example 87a

- 20 An aliquot (0.01 mole) of the diazonium salt from 2,6-dichloro-4-nitroaniline prepared in Example 86a was added to a chilled solution of the coupler (3.29 g, 0.01 mole) having the formula



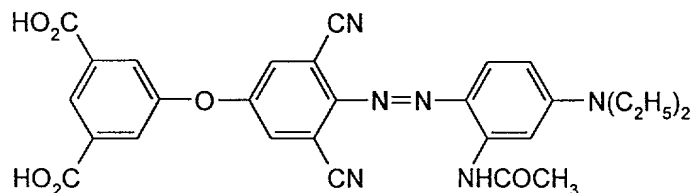
dissolved in 1:5 acid (100 mL) with stirring at 0-5°C. Ammonium acetate was added with stirring until the coupling mixture was neutral to Congo Red Test paper. After allowing to stand for 1.0 hour, water was added to the coupling mixture to precipitate the dye, which was collected by filtration, washed with water and dried in air (yield - 4.27 g). An absorption maximum was observed at 460 nm in the visible light absorption spectrum in DMF.

10 Example 87b

A mixture of N-(2-chloroethyl)-N-ethylaniline (46.0 g, 0.25 mole), dimethyl 5-hydroxyisophthalate (52.5 g, 0.25 mole), potassium carbonate (69.08), a trace of pulverized potassium iodide and DMF (350 mL) was heated at 125-30°C for 3.5 hours with stirring. The reaction mixture was allowed to cool and drowned in water/ice mixture (1.0 L). The product separated as a brown oil and the aqueous layer was removed by decantation. To the oily product was added 2-ethoxyethanol (175 mL) and aqueous 50% NaOH (50.0 g) and the hydrolysis reaction mixture was heated at 60-65°C for about 20 minutes. Acetone was added to the reaction mixture and the white solid was collected by filtration, washed with acetone and dried in air (yield - 99.0 g). The disodium salt was dissolved in water (250 mL) by stirring. Acidification with conc. HCl to a pH of about 3.0 gave a slightly sticky product which solidified in a few minutes. The pale yellow granular solid was collected by filtration, washed with water and dried in air (yield - 58.0 g). FDMS indicated the structure to be that of the coupler used in Example 87a.

Example 88

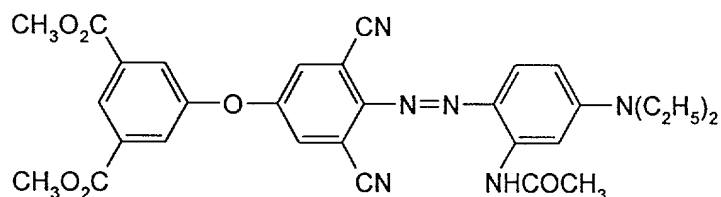
35 A mixture of the diacid azo compound (0.70 g, 0.0013 mole) having the structure



1,6-hexanediol, dimethanesulfonate (0.36 g, 0.0013 mole), potassium carbonate (0.35 g) and DMF (5.0 mL) was heated at about 95°C with occasional stirring for 2.0 hours. The reaction mixture was drowned into methanol (50 mL) and the polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 0.55 g). GPC indicated a Mw of 7,353, a Mn of 2,431 and a Mw/Mn of 3.02. An absorption maximum at 537 nm was observed in the visible light absorption spectrum in DMF.

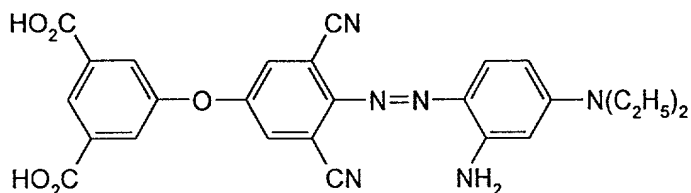
#### Example 88a

To a mixture of the diester dye (1.75 g, 0.0013 mole) having the structure



and 2-ethoxyethanol (20 mL) was added aqueous 50% NaOH solution (1.2 g) and the hydrolysis mixture was heated at about 10 minutes at about 95°C. The reaction mixture was drowned into acetone and the solid material collected by filtration. The acetone-wet material was dissolved by stirring in water (200 mL) and the diacid dye precipitated by adding acetic acid. The product was collected by

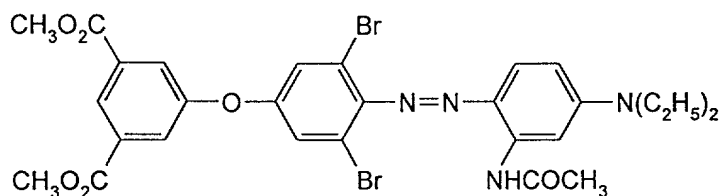
filtration washed with water and dried in air (yield - 1.35 g). FDMS showed the product to be mostly



5 indicating hydrolysis of the acetamido group in addition to the ester group. All of the product was added to acetic acid (8.0 mL) and acetic anhydride (1.0 mL). The reaction mixture was heated at 95°C for 30 minutes with occasional stirring. A bathochromic shift in color from  
10 red to magenta was observed as the amine group was acetylated. The reaction mixture was allowed to cool, whereupon a solid dark red product crystallized, and then was drowned into methanol (40 mL). The product was collected by filtration, washed with water and dried in  
15 air (yield - 0.90 g). FDMS indicated the structure to be that of the diacidic azo dye in Example 88.

#### Example 88b

20 A mixture of the dibromo azo dye (3.00 g, 0.0044 mole) having the structure



sodium dicyanocuprate (0.69 g, 0.005 mole) and DMF (30 mL)  
25 was heated at 95°C for 1.0 hour. The reaction mixture was drowned into methanol (150 mL) and the dye was collected

006221 99475460

by filtration, washed with methanol and dried in air (yield - 1.91 g). FDMS indicated the structure to be that of the dicyano dye used in Example 88a.

5 Example 88c

To conc.  $\text{H}_2\text{SO}_4$  (7.5 mL) was added dry  $\text{NaNO}_2$  (1.08 g) portionwise with stirring and the temperature allowed to rise. The nitrosyl sulfuric acid mixture was cooled and  
10 1:5 acid (15 mL) was added at less than  $10^\circ\text{C}$  with stirring. To this mixture was added at  $0-5^\circ\text{C}$  with stirring dimethyl 5-(4'-amino,2',6'-dibromophenoxy)isophthalate (6.86 g, 0.015 mole), followed by an additional 15 mL of 1:5 acid. The diazotization  
15 reaction mixture was stirred at  $0-5^\circ\text{C}$  for 2.0 hours and then an aliquot (0.0075 mole) was added to a solution of 3-acetamido-N,N-diethylaniline (1.54 g, 0.0075 mole) dissolved in 1:5 acid (75 mL) at  $0-5^\circ\text{C}$ . Ammonium acetate was added with stirring to the coupling mixture until  
20 neutral to Congo Red test paper. Coupling was allowed to continue at  $0-5^\circ\text{C}$  for 1.0 hour and the dye then precipitated by addition of water, collected by filtration, washed with water and dried in air. FDMS indicated the structure to be that of the starting dibromo  
25 azo dye in Example 88b. An absorption maximum at 546 nm was observed in the visible light absorption spectrum in DMF.

Example 88d

30

A mixture of the dimethyl 5-(4'-aminophenoxy)isophthalate (15.0 g, 0.05 mole) (Example 12c), anhydrous sodium acetate (9.6 g) and acetic acid (85 mL) was treated with stirring with bromine (17.4 g, 0.11 mole) allowing the  
35 temperature to rise. The reaction mixture was heated at  $70-80^\circ\text{C}$  for 1.5 hours, allowed to cool, and then drowned



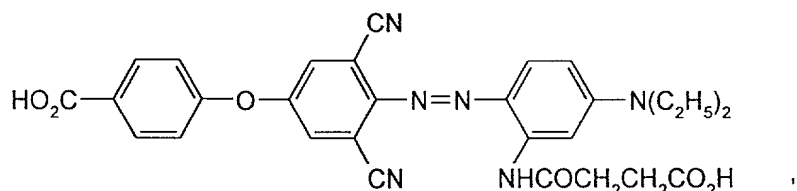
into ice water (350 mL). The product was collected by filtration, washed with water and dried in air (yield - 21.9 g). FDMS indicated the structure to be that of the amine compound diazotized in Example 88c.

5

Example 89

A mixture of the diacidic azo compound (1.39 g, 0.0025 mole) having the structure

10



1,6-hexanediol, dimethanesulfonate (0.68 g, 0.0025 mole), potassium carbonate (1.0 g) and DMF (8.0 mL) was heated at 95°C for 2.5 hrs with occasional stirring. The reaction mixture was drowned into methanol (100 mL) and the red polydye was collected by filtration, washed with water containing acetic acid, hot water and dried in air (0.85 g). GPC analysis indicated a Mw of 2,772, a Mn of 1,306 and a Mw/Mn of 2.12. An absorption maximum was observed at 538 nm in the visible light absorption spectrum in DMF.

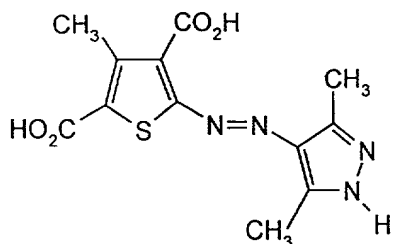
15

20

Example 90

A mixture of the diacidic azo compound (1.23 g, 0.004 mole) having the formula

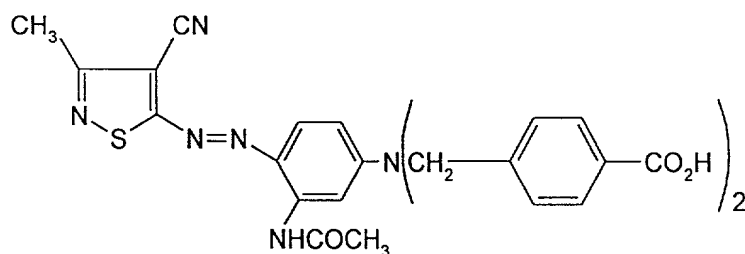
25



1,2-hexanediol, dimethanesulfonate (1.1 g, 0.004 mole), potassium carbonate (0.55 g) and DMF (8.0 mL) was heated at 95°C for 1 hour. The reaction mixture was drowned into water (250 mL) containing acetic acid (5.0 mL). The yellow polydye was collected by filtration, washed with water and dried in air (yield - 1.21 g). GPC analysis indicated a Mw of 1,726, a Mn of 1,079 and a Mw/Mn of 1.6. An absorption maximum at 400 nm was observed in the visible light absorption spectrum in DMF.

#### Example 91

A mixture of the diacidic azo compound (1.71 g, 0.003 mole) having the formula



1,6-hexanediol, dimethanesulfonate (0.82 g, 0.003 mole), potassium carbonate (0.85 g) and DMF (8.0 mL) was heated with occasional stirring at 95°C for 2.0 hours. The reaction mixture was drowned into methanol (100 ml) and the red polydye was collected by filtration, washed with methanol, water containing acetic acid, hot water and

dried in air (yield- 1.5 g). GPC indicated a Mw of 2,090, a Mn of 1,235 and a Mw/Mn of 1.69. An absorption maximum was observed at 545 nm in the visible light absorption spectrum in DMF.

5

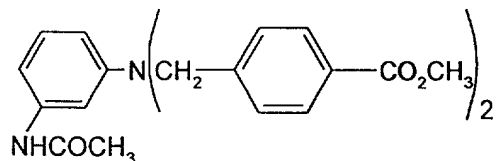
Example 91a

To conc.  $\text{H}_2\text{SO}_4$  (5.0 mL) was added dry  $\text{NaNO}_2$  (0.72 g) portionwise with stirring, allowing the temperature to  
10 rise. The nitrosyl sulfuric acid solution was stirred and cooled and 1:5 acid (10 ml was added below about 15°C, followed by 5-amino-4-cyano-3-methylisothiazole (1.39 g, 0.01 mole) and 1:5 acid (10 ml) both added at 0-5°C. After being stirred at 0-5°C for 2.0 hours an aliquot  
15 (0.005 mole) of the diazonium solution was added to a stirred solution of 3-acetamido-N,N-bis-(4-carboxyphenylmethyl)aniline (2.09 g, 0.005 mole) dissolved in 1:5 acid (30 ml) at 0-5°C. Ammonium acetate was added to neutralize the coupling mixture until neutral to Congo  
20 Red test paper. Water was added to the coupling mixture to precipitate the red dye, which was collected by filtration and dried in air (yield- 2.67 g). The product was reslurried in hot methanol, allowed to cool and the solid collected by filtration, washed with methanol and  
25 dried in air (yield- 2.10 g). FDMS indicated the structure to be that of the diacid azo compound used as a starting material for Example 91.

Example 91b

30

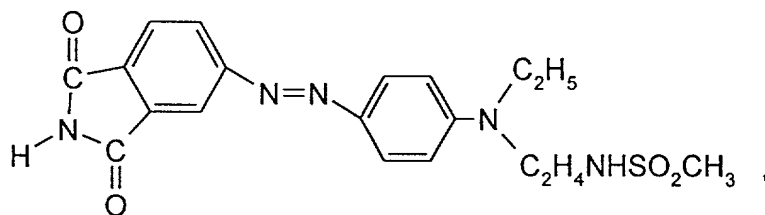
To a slurry of the diester compound (12.00 g, 0.0269 mole) having the structure



in water (150 ml) was added aqueous 50% NaOH solution (10.80 g) and 2-ethoxyethanol (20 ml). The reaction mixture was heated at about 70-80°C for 2.0 hours and allowed to cool. The cloudy reaction mixture was clarified by filtering through Celite filter aid and the filtrate was drowned into ice/water mixture (150 g). Conc. HCl was added dropwise with stirring to bring the pH to about 2.5. The tan solid was collected by filtration, washed with water and dried at 40°C under nitrogen (yield- 10.04 g). FDMS indicated the product to have the structure of the coupler used in Example 91a.

#### Example 92

A mixture of the diacidic azo compound (0.83 g, 0.002 mole) having the structure



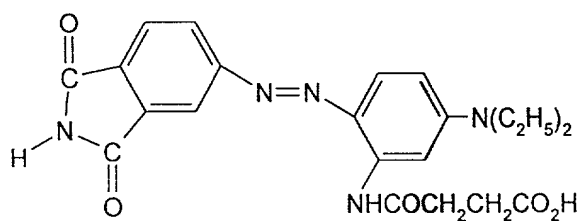
1,2-ethanediol, dimethanesulfonate (0.44 g, 0.002 mole), potassium carbonate (0.5 g) and DMF (7.5 ml) was heated at about 95°C for 3.0 hours. The polydye was isolated by drowning the reaction mixture into water and acidifying with acetic acid, followed by filtering, washing with water and drying in air. GPC analysis indicated a Mw of 2,379, a Mn of 1,363 a Mw/Mn of 1.74. An absorption

maximum was observed in DMF in the visible absorption spectrum at 480 nm.

Example 93

5

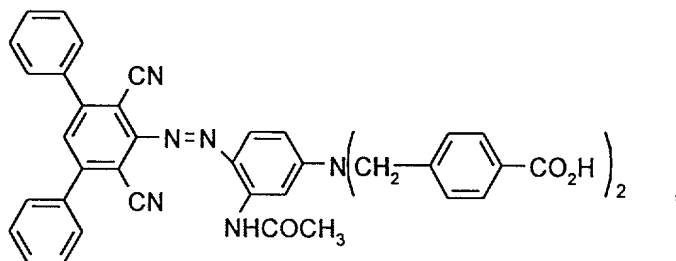
A mixture of the diacidic azo compound (1.26 g, 0.003 mole) having the structure



- 10 1,6-hexanediol, dimethanesulfonate (0.82 g, 0.003 mole),  
potassium carbonate (0.50 g) and DMF (8.0 mL) was heated  
at about 95°C for 1.5 hours. The reaction mixture was  
drowned into methanol (100 mL) and acetic acid (1.0 mL)  
was added. The initially sticky polydye solidified after  
15 standing for about 1.0 hour and was collected by  
filtration, washed with methanol, water containing acetic  
acid, hot water and dried in air (yield - 0.60 g). GPC  
analysis indicated a Mw of 2,667, a Mn of 1,695 and a  
Mw/Mn of 1.57. An absorption maximum at 508 nm was  
20 observed in the visible light absorption spectrum in DMF.

Examples 93a

- 25 A mixture of the diacidic azo compound (3.62 g, 0.005 m)  
having the structure



1,2-ethanediol, dimethanesulfonate (1.10 g, 0.005 m), potassium carbonate (1.50 g) and DMF (30 mL) was heated at about 95°C with stirring for 2.0 hours. The reaction mixture was drowned into methanol (100 mL) and the red polydye was collected by vacuum filtration and washed with methanol, water containing acetic acid, hot water and dried in air (yield- 3.08 grams). GPC analysis indicated a Mw of 7,176, a Mn of 3,533 and a Mw/Mn of 2.02. An absorption maximum was observed in the visible light absorption spectrum at 525 nm.

#### Example 93b

To conc.  $\text{H}_2\text{SO}_4$  (5.0 mL) was added dry  $\text{NaNO}_2$  (0.72 g) portionwise with stirring, allowing the temperature to rise. The nitrosyl sulfuric acid solution was stirred and cooled and 1:5 acid (1 part propionic:5 parts acetic acid) (10 mL) was added below about 15°C, followed by 2,6-dicyano-3,5-diphenylaniline (2.95 g, 0.01 m) and 1:5 acid (10 mL) both added at 0-5°C. After being stirred for 2.0 hours at 0-5°C, the diazonium solution was added to a stirred solution of 3-acetamido-N,N-bis (4-carboxyphenylmethyl)aniline (4.18 g, 0.01 m) dissolved in a mixture of 1:5 acid (75 mL) plus 15% aqueous sulfuric acid (15 mL) at 0-5°C. Ammonium acetate was added portionwise until the coupling mixture was neutral to Congo Red test paper. After about 1.0 hour, water was added to the coupling mixture and the resulting slurry

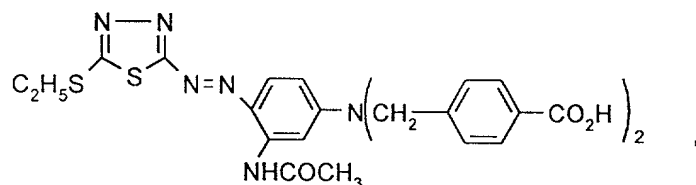
heated to about 60°C. The red product was collected by filtration, washed well with hot water and dried in air (yield - 5.43 g). FDMS analysis indicated the structure to be that of the starting material for Example 93-1.

5

Example 93c

A mixture of the diacidic azo compound (1.80 g, 0.003 m) having the structure

10



1,2-ethanediol, dimethanesulfonate (0.66 g, 0.003 m),  
15 potassium carbonate (1.0 g) and DMF (8 mL) was heated at  
about 95°C with occasional stirring. The polydye was  
isolated by drowning the reaction mixture into methanol  
(100 mL) followed by filtration and washing with methanol,  
water containing acetic acid, water and was then dried in  
20 air (yield - 0.52 g). GPC analysis using NMP (N-methyl-2-  
pyrrolidinone) solvent indicated a Mw of 5,413, a Mn of  
2,196 and a Mw/Mn of 2.46. An absorbance maximum at 517  
nm was observed in the visible absorption maximum in DMF.

25 Example 93d

A sample of 2-amino-5-ethylthio-1,3,4-thiadiazole (1.61 g,  
0.01 m) was diazotized and coupled with 3-acetamido-N,N-  
bis(4-carboxyphenylmethyl)aniline (4.18 g, 0.01 m) and the  
30 red product isolated using the procedure described above  
in Example 93-1a. FDMS indicated the structure of the azo  
compound to be that of the starting material for Example  
93-2.

Examples 94-118

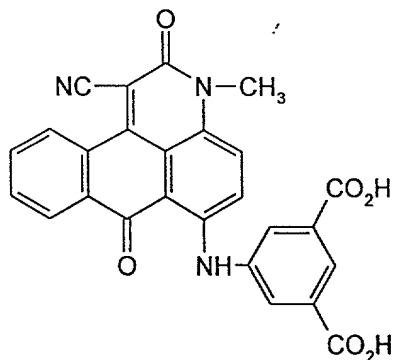
-  
Colored EASTAR® PETG 6763 film was produced by melt  
5 blending the polydyes of Examples 69-93 and extruding  
according to the following procedures to produce Examples  
94-118 (Table 2).

EASTAR® PETG polyester 6763, a  
10 poly(ethylene-cyclohexanedimethylene) terephthalate  
(Eastman Chemical Company) (300 g of previously dried  
pellets) was dry blended with the azo dye composition  
(0.12 g) and the blend extruded and finally a 18-20 mil  
thick film prepare as described above for Examples 37-66.

15

Example 119

A mixture of the diacidic anthrapyridone compound (0.93 g,  
0.002 mole) having the structure



20

1,2-ethandiol, dimethanesulfonate (0.44 g, 0.002 mole),  
potassium carbonate (.5 g) and DMF (8.0 mL) was heated at  
about 95°C for 3.0 hours with occasional stirring. The  
25 reaction mixture was drowned into methanol (100 mL) and  
the violet polydye was collected by filtration, washed



with methanol, water containing acetic acid, water and dried in air (yield - 1.09 g). A number average molecular weight of 1,228 was obtained by GPC analysis. Absorption maxima at 544 and 583 nm were observed in the visible light absorption spectrum in DMF.

Example 119a

To a mixture of 1-cyano-6-(3',5'-dicarbomethoxyphenylamino)-3-methyl-3H-dibenz[f,i]isoquinoline-2,7-dione (2.00 g, 0.00405 mole) stirred in 2-ethoxyethanol (50 mL) was added aqueous 50% NaOH solution (2.47 g). The reaction mixture was heated at 90-95°C for 50 minutes and then was drowned into water. The mixture was acidified by addition of acetic acid and the solid product was collected by filtration, washed with water and dried in air (yield - 1.78 g). FDMS indicated the product to be the diacidic anthrapyridone compound reacted in Example 119.

Example 119b

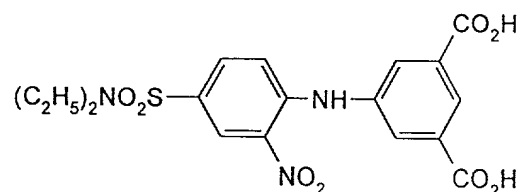
A mixture of 6-bromo-1-cyano-3-methyl-3H-dibenz[f,i]isoquinoline-2,7-dione (11.0 g, 0.03 mole), dimethyl 5-aminoisophthalate (25.1 g, 0.12 mole), cupric acetate (3.6 g), potassium carbonate (3.0 g) and DMF (90 mL) was heated and stirred under nitrogen to about 135-40°C. The reaction mixture became very thick and turned violet. Additional DMF (40 mL) was added and heating was continued at 135-40°C for 2.0 hours. The reaction mixture was allowed to cool to about 60°C and poured on a coarse fritted glass funnel for vacuum filtration. The product was washed with DMF and water and the water-wet cake was reslurried in boiling acetone (250 mL). After cooling, the product was collected by filtration, washed with acetone and dried in air (yield - 10.8 g). FDMS indicated

the product to be the diester anthrapyridone compound used in Example 119a.

Example 120

5

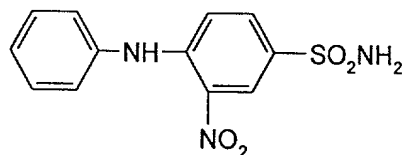
A mixture of the diacidic nitroarylamine compound (2.50 g, 0.0057 mole) having the structure



10 1,2-ethanediol, dimethanesulfonate (1.25 g, 0.0057 mole), potassium carbonate (1.6 g) and DMF (15 mL) was heated at 95°C for 2.5 hours. The reaction mixture was drowned into methanol (200 mL) and the yellow polydye was collected by filtration, washed containing acetic acid, water and dried at 40°C (yield - 0.77 g). An absorption maximum was  
15 observed at 412 nm in the visible absorption spectrum in DMF.

Example 121

20 A mixture of the diacidic nitroarylamine compound (4.40 g, 0.015 mole) having the structure



25 1,2-ethanediol, dimethanesulfonate (3.27 g, 0.015 mole), potassium carbonate (2.0 g) and DMF 40 mL) was heated at 90-95°C with stirring for 4.0 hours. The reaction mixture was drowned into methanol (200 mL) and the yellow polydye was collected by filtration, washed with methanol,

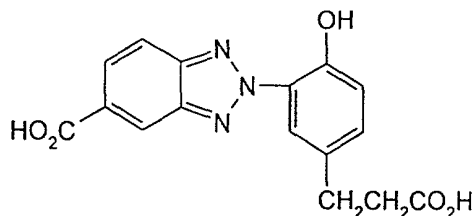
water containing acetic acid, water and dried in air  
(yield - 1.80 g). GPC analysis indicated a Mw of 1,585, a  
Mn of 1,024, a Mw/Mn of 1.54. An absorption maximum at  
416 nm was observed in the visible light absorption  
5 spectrum in DMF.

Examples 122-124

Colored polyester film was produced by melt blending and  
10 extruding EASTAR® PETG polyester 6763 (Eastman Chemical  
Company) (300 g previously dried pellets) which had dry  
blended with the polydyes of Examples 119, 120, 121 to  
produce Examples 122-124, respectively, according to the  
procedure used to produce Examples 37-66. The film of  
15 Example 122 was violet and those of Examples 123 and 124  
were bright yellow.

Example 125

20 A mixture of the benzotriazole UV light absorbing compound  
(3.27 g, 0.01 mole) having the structure

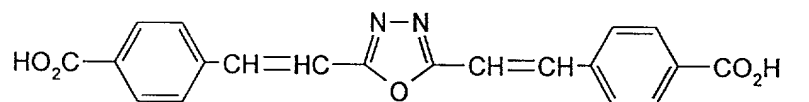


1,2-ethanediol, dimethanesulfonate (2.18 g, 0.01 mole),  
potassium carbonate (2.76 g) and DMF (25 mL) was heated at  
25 about 95°C for 6.0 hours. The reaction mixture was  
drowned into methanol (200 mL) and a little acetic acid  
added. The polymeric UV light absorbing compound was  
collected by filtration, washed with water containing a  
little acetic acid, hot water and then dried in air (yield  
30 - 2.88 g). GPC analysis indicated a Mw of 7,561, a Mn of

2,632 and a Mw/Mn of 2.87. An absorption maximum was observed at 350 nm in the UV light absorption spectrum in methylene chloride.

5 Example 126

A benzylidene type UV light fluorescent compound (1.0 g, 0.0028 mole) having the structure



- 10 1,6-hexenediol, dimethanesulfonate (0.0028 mole),  
potassium carbonate (0.97 g) and DMF (10 mL) were mixed  
and the reaction mixture was heated at for 3.0 hours at  
about 120-130°C. The reaction mixture was drowned into  
methanol (100 mL) and the polymer was collected by  
15 filtration, washed with methanol, water containing acetic  
acid, hot water and dried in air (yield - 0.69 g). GPC  
indicated a Mw of 50,717, a Mn of 16,044 and a MW/Mn of  
3.16.

20 Example 127

- EASTAPAK® PET 7352, a poly(ethyleneterephthalate) (Eastman  
Chemical Company) (400 g of previously dried pellets) was  
dry blended with the polymeric UV light fluorescent  
25 material of Example 126 (0.16 g). The blend was extruded  
with a C. W. Brabender ¼ inch extruder, equipped with a  
mixing screw, at 285°C into a water bath and the extrudate  
pelletized. The pellets which contained about 400 ppm of  
the UV light absorber showed a strong blue white  
30 fluorescence under UV light.

Example 128

Example 127 was repeated except that 8 mg of the UV light fluorescent material of Example 126 was added to the  
5 EASTAPAK® PET 7352. The resulting pellets showed a strong blue-white fluorescence under UV light and appeared very white in sunlight.

Example 129

10 A mixture of Pc-Al-O-C<sub>6</sub>H<sub>3</sub>-3,5-diCO<sub>2</sub>H (Pc = phthalocyanine) (1.74 g, 0.0024 mole), 1,6-hexanediol, dimethanesulfonate (0.66 g, 0.0024 mole), potassium carbonate (0.83 g) and DMF (10 mL) was heated and stirred at about 125°C for 1  
15 hour and then at about 140°C for 1 hour. The reaction mixture was drowned into methanol (50 mL) and the polymeric product was collected by filtration, washed with methanol, water containing acetic acid, hot water and dried in air (yield - 1.48 g).

Example 130

EASTAPAK® PET 7352, a poly(ethyleneterephthalate) (Eastman Chemical Company) (400 g of previously dried pellets) was  
25 dry blended with the polymeric phthalocyanine compound of Example 129 (0.12g). The blend was extruded with a C. W. Brabender ¾ inch extruder, equipped with a mixing screw, at 285° into a water bath and the extrudate pelletized. The cyan pellets were redried at 70°C for about 17 hrs at  
30 a pressure of about 1-5 torr. A portion of the dried pellets (1.40 g) was pressed into a film at 285°C using a 2-inch diameter circular mold in a Pasadena Hydraulic, Inc. press using 12,000 pounds ram force (4-inch ram). A transparent cyan film was produced by quenching in water  
35 and had an absorption maximum at 684 nm in the light absorption spectrum.

Example 131

Example 130 was repeated except that 4 mg of the polymeric  
5 phthalocyanine compound of Example 129 was added to the  
PET . The final film contained about 10 ppm and had a  
light absorption maximum at 685 nm.

Example 132

10 EASTAPAK® PET 7352, a poly(ethyleneterephthalate) (Eastman  
Chemical Company) (400 g of dried pellets) was dry blended  
with the polydye of Example 18 (0.6 g). The blend was  
extruded with a C. W. Brabender  $\frac{3}{4}$  inch extruder, equipped  
15 with a mixing screw, at 285°C into a water bath and the  
extrudate pelletized. Good color production resulted with  
no evidence of color loss by sublimation to give dark red  
pellets containing about 0.15% by weight of the polydye.

20 Example 133

Example 132 was repeated using 0.6 g of the polydye of  
Example 75 as the colorant to give yellow pellets having  
about 0.15% by weight of the polydye. No loss of color by  
25 sublimation was observed.

Examples 134-182

The diacidic azo compounds of Formula VI in Table 3 are  
30 reacted with essentially equimolar amounts of 1,2-  
ethanediol, dimethanesulfonate in DMF in the presence of  
potassium carbonate to yield the polydyes of Examples 134-  
182 in Table 3.

Examples 183-193

The diacidic diazo compounds of Formula VII in Table 4 are reacted with essentially equimolar amounts of 1,4-  
5 butanediol, dimethanesulfonate in DMF in the presence of potassium carbonate to yield the polydyes of Examples 183-193 in Table 4.

Examples 194-202

10 The diacidic bisazo compounds of Formula VIIa in Table 5 are reacted with essentially equimolar amounts of 1,3-propanediol, dimethanesulfonate in DMF in the presence of sodium carbonate to yield the polydyes of Examples 194-202  
15 in Table 5.

Examples 203-211

20 The diacidic benzylidene (methine) compounds in Table 6 are reacted with essentially equimolar amounts of 1,4-cyclohexanedimethanol, dimethanesulfonate in DMF in the presence of sodium carbonate to yield the polydyes of Examples 203-211 in Table 6.

25 Examples 212-220

The diacidic 3-aryl-2,5-dioxypyrroline compounds of Formula X in Table 7 are reacted with essentially  
equimolar amounts of diethylene glycol, dimethanesulfonate  
30 in DMF in the presence of potassium carbonate to yield the polydyes of Examples 212-220 in Table 7.

Examples 221-230

35 The diacidic 3-aryl-5-dicyanomethylene-2-oxypyrroline compounds of Formula XI in Table 8 are reacted with

006627 997560

essentially equimolar amounts of triethylene glycol, dimethanesulfonate to yield the polydyes of Examples 221-230 in Table 8.

5    Examples 231-239

10    The diacidic azo-methine compounds of Formula XIII in Table 9 are reacted with essentially equimolar amounts of 1,4-butanediol, dimethanesulfonate in DMF in the presence of potassium carbonate to yield the polydyes of Examples 231-239 in Table 9.

Examples 240-269

15    The diacidic anthraquinone compounds of Formula XIV in Table 10 are reacted with essentially equimolar amounts of 2,2,4,4-tetramehtyl-1,3-cyclobutanediol, dimethanesulfonate in N,N-dimethylacetamide in the presence of potassium carbonate to yield the polydyes of  
20    Examples 240-269 in Table 10.

Examples 270-326

25    The diacidic anthraquinone compounds of Formula XV in Table 11 are reacted with essentially equimolar amounts of 1,2-ethanediol, dimethanesulfonate in DMF in the presence of potassium carbonate to yield the polydyes of Examples 270-326 in Table 11.

30    Examples 327-344

35    The diacidic anthraquinone compounds of Formula XVI in Table 12 are reacted with essentially equimolar amounts of 1,6-hexanediol, dimethanesulfonate in N-methyl-2-pyrrolidinone in the presence of sodium carbonate to yield the polydyes of Examples 327-344 in Table 12.



Examples 345-361

5 The diacidic anthrapyridine compounds of Formula XVIII in Table 13 are reacted with essentially equimolar amounts of 1,4-butanediol, di-p-toluenesulfonate in the presence of DMF to yield the polydyes of Examples 345-361 in Table 13.

Examples 362-381

10 The diacidic anthraquinone compounds of Formula XIX in Table 14 are reacted with 2,2-dimethyl-1,3-propanediol, dimethanesulfonate in essentially equimolar amounts in DMF in the presence of potassium carbonate to yield the  
15 polydyes of Examples 362-381 in Table 14.

Examples 382-396

20 The diacidic anthraquinone compounds of Formula XIXc of Table 15 are reacted with essentially equimolar amounts of 1,2-ethanediol, dimethanesulfonate in DMF in the presence of potassium carbonate to yield the polydyes of Examples 382-396 in Table 15.

25 Examples 397-414

The diacidic anthraquinone compounds of Formula XIXd in Table 16 are reacted with essentially equimolar amounts of 1,6-hexanediol, dimethanesulfonate in DMF in the presence  
30 of potassium carbonate to yield the polydyes of Examples 397-414 in Table 16.

Examples 415-435

35 The diacidic anthraquinone compounds of Formula XIXe in Table 17 are reacted in essentially equimolar amounts with

ethylene glycol, dimethanesulfonate in DMF in the presence of potassium carbonate to yield the polydyes of Examples 414-435 in Table 17.

5    Examples 436-449

10    The diacidic anthraquinone compounds of Formula XIXf in Table 18 are reacted in essentially equimolar amounts with 1,4-cyclohexanedimethanol, dimethanesulfonate in DMF in the presence of potassium carbonate to yield the polydyes of Examples 436-449 in Table 18.

15    Examples 450-455

20    The diacidic anthrapyridine compounds of Table 19 are reacted with essentially equimolar amounts of 1,6-hexanediol, di-p-toluenesulfonate in DMF in the presence of potassium carbonate to yield the polydyes of Examples 450-455 in Table 19.

25    Examples 456-465 ,

30    The diacidic nitroarylamine compounds of Table 20 are reacted with 1,4-butanediol, dimethanesulfonate in essentially equimolar amounts in DMF in the presence of potassium carbonate to yield the polydyes of Examples 456-465 in Table 20.

35    Examples 466-505

40    The miscellaneous diacidic compounds of Table 21 are reacted with essentially equimolar amounts of the disulfonate compounds of Table 21 in DMF in the presence of potassium carbonate to yield the polydyes of Examples 466-505 in Table 21.

Examples 506-522

5 The diacidic UV light absorbing compounds of Table 22 are reacted with essentially equimolar amounts of the disulfonate compounds of Table 22 in DMF in the presence of potassium carbonate to yield the polymeric UV absorbers of Examples 506-522 in Table 22.

Examples 523-536

10 The diacidic infrared light absorbing compounds of Table 23 are reacted with essentially equimolar amounts of the disulfonate compounds of Table 23 in DMF in the presence of potassium carbonate to yield the polymeric infrared  
15 light absorbing compounds of Examples 523-536 in Table 23.

006347 99116 133900

Table 1

Example No.	Polydye Melt Blended and Extruded With EASTAR® PETG	Color of Film
37	Polydye of Example 7	Blue
38	Polydye of Example 8	Blue
39	Polydye of Example 9	Blue
40	Polydye of Example 10	Blue
41	Polydye of Example 11	Blue
42	Polydye of Example 12	Blue
43	Polydye of Example 13	Greenish-blue
44	Polydye of Example 14	Reddish-blue
45	Polydye of Example 15	Blue
46	Polydye of Example 16	Green
47	Polydye of Example 17	Bright blue
48	Polydye of Example 18	Bluish-red
49	Polydye of Example 19	Yellow
50	Polydye of Example 20	Orange
51	Polydye of Example 21	Red
52	Polydye of Example 22	Green
53	Polydye of Example 23	Yellow
54	Polydye of Example 24	Yellow
55	Polydye of Example 25	Yellow
56	Polydye of Example 26	Yellow
57	Polydye of Example 27	Blue
58	Polydye of Example 28	Red
59	Polydye of Example 29	Greenish-yellow
60	Polydye of Example 30	Yellow
61	Polydye of Example 31	Greenish-yellow
62	Polydye of Example 32	Blue
63	Polydye of Example 33	Yellow
64	Polydye of Example 34	Yellow
65	Polydye of Example 35	Greenish-blue
66	Polydye of Example 36	Greenish-blue

Table 2  
Azo Polydyes in EASTAR® PETG 6763  
(300 ppm)

Example No.	Polydye Melt Blended and Extruded With EASTAR® PETG	Color of Film
94	Polydye of Example 69	Violet
95	Polydye of Example 70	Blue
96	Polydye of Example 71	Yellow-brown
97	Polydye of Example 72	Red
98	Polydye of Example 73	Orange
99	Polydye of Example 74	Yellow
100	Polydye of Example 75	Greenish-yellow
101	Polydye of Example 76	Scarlet
102	Polydye of Example 77	Yellow
103	Polydye of Example 78	Scarlet
104	Polydye of Example 79	Yellow
105	Polydye of Example 80	Red
106	Polydye of Example 81	Orange
107	Polydye of Example 82	Reddish-brown
108	Polydye of Example 83	Red
109	Polydye of Example 84	Reddish-blue
110	Polydye of Example 85	Blue
111	Polydye of Example 86	Brown
112	Polydye of Example 87	Reddish-brown
113	Polydye of Example 88	Magenta
114	Polydye of Example 89	Magenta
115	Polydye of Example 90	Yellow
116	Polydye of Example 91	Red
117	Polydye of Example 92	Orange
118	Polydye of Example 93	Scarlet

Table 3  
Polydyes From Diacid Compounds of Formula VI

$R_4 - N = N - Z$

Example No.	$R_4$	Z	Color
134			violet
135			red
136			magenta
137			violet
138			scarlet
139			red
140			violet

Table 3  
Polydyes From Diacidic Compounds of Formula VI

$R_4 - N = N - Z$

Example No.	$R_4$	Z	Color
141			blue
142			orange
143			scarlet
144			magenta
145			magenta
146			bluish-red
147			violet

Table 3  
Polydyes From Diacidic Compounds of Formula VI

$R_4 - N = N - Z$

Example No.	$R_4$	Z	Color
148			blue
149			red
150			violet
151			violet
152			orange
153			red
154			navy blue



Table 3  
Polydyes From Diacidic Compounds of Formula VI

$R_6 - N = N - Z$

Example No.	$R_6$	Z	Color
155			blue
156			red
157			orange
158			red
159			blue
160			blue
161			red

Table 3  
Polydyes From Diacidic Compounds of Formula VI

$R_6 - N = N - Z$

Example No.	$R_6$	Z	Color
162			red
163			red
164			orange
165			yellow
166			yellow
167			orange
168			yellow

Table 3  
Polydyes From Diacidic Compounds of Formula VI

$R_6 - N = N - Z$

Example No.	$R_6$	Z	Color
169			red
170			red
171			orange
172			blue
173			red
174			yellow
175			yellow

Table 3  
Polydyes From Diacidic Compounds of Formula VI

$R_4 - N = N - Z$

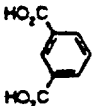
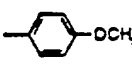
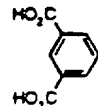
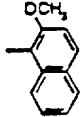
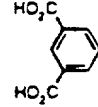
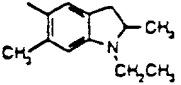
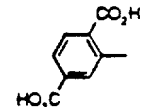
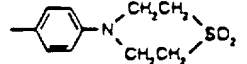
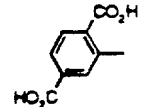
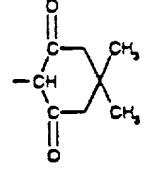
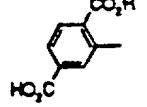
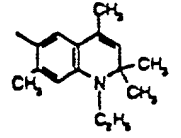
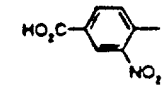
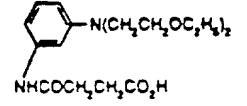
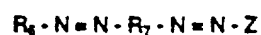
Example No.	$R_4$	Z	Color
176			yellow
177			yellow
178			orange
179			yellow
180			yellow
181			orange
182			red

Table 4  
Polydyes From Diacidic Compounds of Formula VII



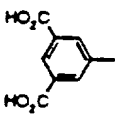

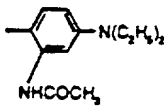
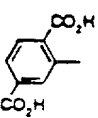
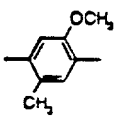
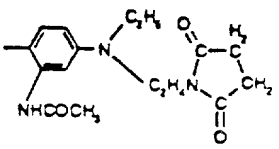
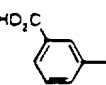

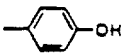


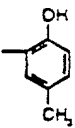
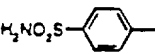

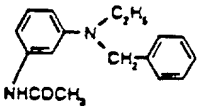
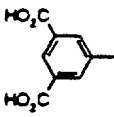
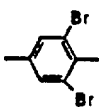
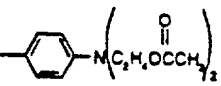
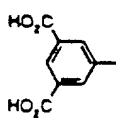
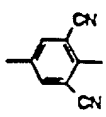
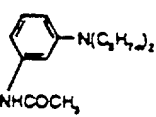
Example No.	R <sub>6</sub>	R <sub>7</sub>	Z	Color
183				red
184				red
185				reddish yellow
186				reddish yellow
187				red
188				yellow brown
189				blue

Table 4  
Polydyes From Diacidic Compounds of Formula VII



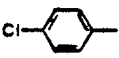
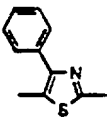
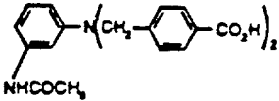
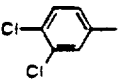
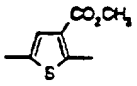
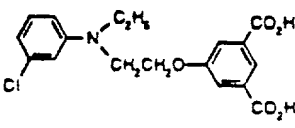


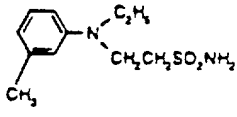
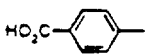
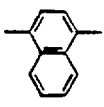
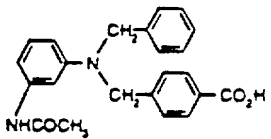
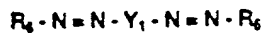
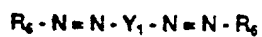
Example No.	R <sub>6</sub>	R <sub>7</sub>	Z	Color
190				red
191				red
192				reddish orange
193				red

Table 5  
Polydyes From Diacidic Compounds of Formula VIIa



Example No.	R <sub>5</sub>	Y <sub>1</sub>	Color
194			red
195			orange
196			orange
197			orange
198			violet
199			red
200			violet

Table 5  
Polydyes From Diacidic Compounds of Formula VIIa




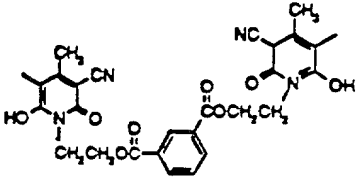
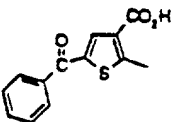
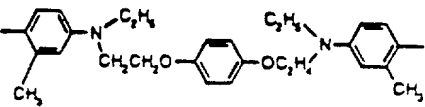
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201			yellow
202			blue



Table 6  
Polydyes From Diacidic Compounds of Formula VIII

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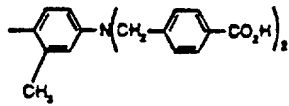

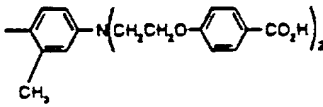
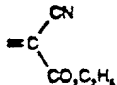
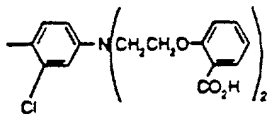
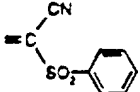
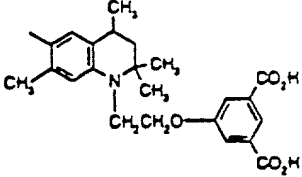
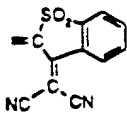
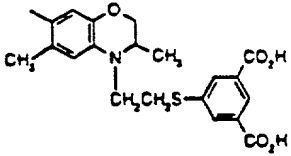
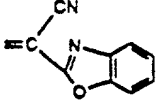
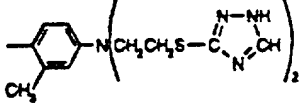
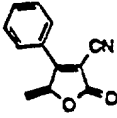
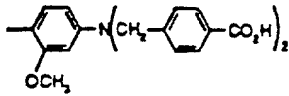
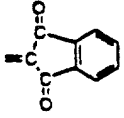
Example No.	$R_{11}$	D	Color
203			yellow
204			yellow
205			yellow
206			blue
207			yellow
208			red
209			red

Table 6  
Polydyes From Diacidic Compounds of Formula VIII

$R_{11} - CH = D$

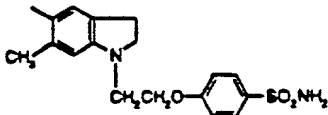
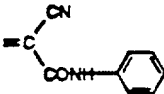
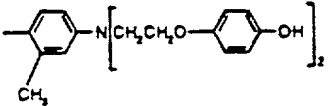
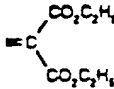
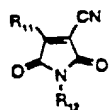
Example No.	$R_{11}$	D	Color
210			yellow
211			yellow

Table 7  
Polydyes From Diacidic Compounds of Formula X

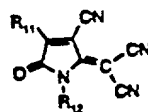


Example No.	$R_{11}$	$R_{12}$	Color
212		$C_2H_5$	red
213		H	red
214		H	red
215		$CH_2C_6H_5$	red
216		$CH_3$	violet
217		$CH_2CH_2OH$	violet

\*C1=C(C#N)C(=O)N(R12)C1=O

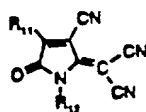
115890.xls

Table 8  
Polydyes From Diacidic Compounds of Formula XI



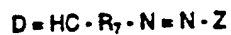
Example No.	R <sub>11</sub>	R <sub>12</sub>	Color
221		H	blue
222		H	greenish blue
223		CH <sub>2</sub> CH=CH <sub>2</sub>	reddish blue
224		CH <sub>2</sub> C <sub>6</sub> H <sub>11</sub>	blue
225		H	blue
226		H	blue

Table 8  
Polydyes From Diacidic Compounds of Formula XI



Example No.	R <sub>11</sub>	R <sub>12</sub>	Color
227		H	blue
228		H	blue
229		H	blue
230			blue

Table 9  
Polydyes From Diacidic Compounds of Formula XIII



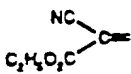
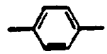
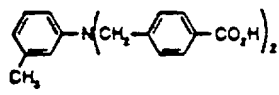
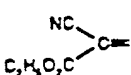
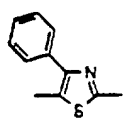
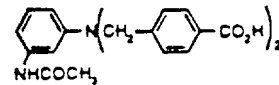
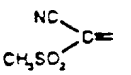
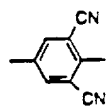
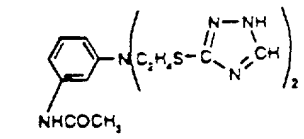
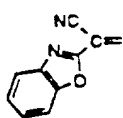

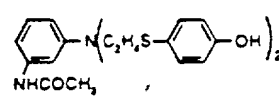


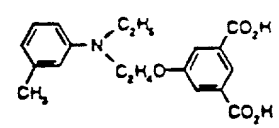
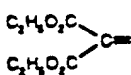

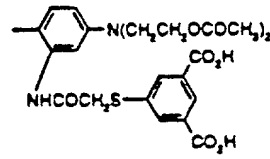
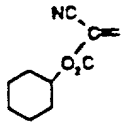
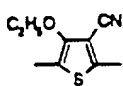
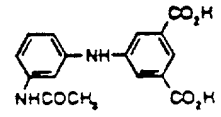
Example No.	D	R <sub>7</sub>	Z	Color
231				red
232				blue
233				blue
234				blue
235				blue
236				blue
237				blue

Table 9  
Polydyes From Diacidic Compounds of Formula XIII

D = HC - R<sub>7</sub> - N = N - Z

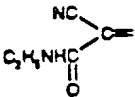

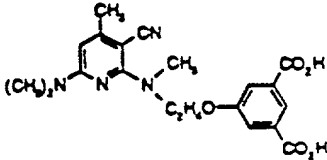
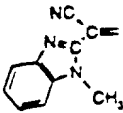
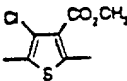
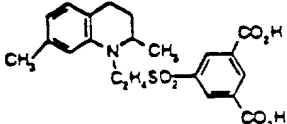
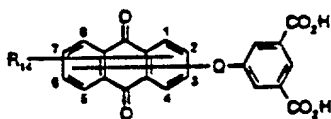
Example No.	D	R <sub>7</sub>	Z	Color
238				blue
239				blue

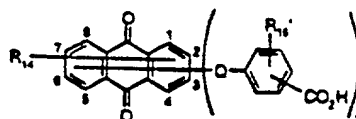


Table 10  
Polydyes From Diacidic Anthraquinone Compounds of Formula XIV



Example No.	Q	R <sub>14</sub>	Color
240	5 - S	1,4-diNHCH <sub>2</sub> C(CH <sub>3</sub> ) <sub>2</sub> CH <sub>2</sub> OH	blue
241	2 - O -	1 - NH <sub>2</sub> , 4 - OH	red
242	2 - S -	1 - NH <sub>2</sub> , 4 - NHSO <sub>2</sub> CH <sub>3</sub>	violet
243	2 - S -	1 - NH <sub>2</sub> , 4 - NHSO <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	violet
244	2 - SO <sub>2</sub> -	1 - NH <sub>2</sub> , 4 - NHC <sub>6</sub> H <sub>5</sub>	blue
245	2 - SO <sub>2</sub> -	1 - NH <sub>2</sub> , 4 - NHC <sub>6</sub> H <sub>4</sub> - 4 - CH <sub>3</sub>	blue
246	2 - SO <sub>2</sub> -	1 - NH <sub>2</sub> , 4 - SC <sub>6</sub> H <sub>5</sub>	violet
247	2 - S -	1 - NH <sub>2</sub> , 4 - NHCOC <sub>6</sub> H <sub>5</sub>	violet
248	4 - S -	1 - NH <sub>2</sub>	red
249	4 - S -	1 - NHC <sub>6</sub> H <sub>11</sub>	violet
250	4 - S -	1 - NHC <sub>6</sub> H <sub>5</sub>	violet
251	4 - NH -	1 - NH <sub>2</sub> , 2 - OCH <sub>3</sub>	violet
252	4 - NH -	1 - NHC <sub>6</sub> H <sub>5</sub>	green
253	4 - NH -	1 - NHC <sub>6</sub> H <sub>5</sub> , 2,6 - diC <sub>2</sub> H <sub>5</sub>	blue
254	4 - NH -	1 - OH	violet
255	2 - S -	1,4 - di - OH	orange
256	2 - SO <sub>2</sub> -	1,4 - di - OH	orange
257	4 - S -	1 - NHCH <sub>3</sub>	violet
258	4 - S -	1 - NHCH <sub>2</sub> CH(C <sub>2</sub> H <sub>5</sub> )C <sub>6</sub> H <sub>5</sub>	violet
259	6(7)S -	1,4 - diNHC <sub>6</sub> H <sub>5</sub> , 2,6 - diC <sub>2</sub> H <sub>5</sub>	cyan
260	6(7)S -	1,4 - diNHC <sub>6</sub> H <sub>5</sub> , 2,4,6 - triCH <sub>3</sub>	cyan
261	6(7)SO <sub>2</sub> -	1,4 - diNHC <sub>6</sub> H <sub>5</sub> , 2 - CH <sub>3</sub> , 6 - C <sub>2</sub> H <sub>5</sub>	cyan
262	4 - NH -	1,8 - diOH, 5 - NO <sub>2</sub>	blue
263	4 - NH -	1,8 - diOH, 5 - NH <sub>2</sub>	blue
264	4 - NH -	1,8 - diOH, 5 - NHC <sub>6</sub> H <sub>5</sub>	blue
265	4 - NH -	1,5 - diOH, 8 - NO <sub>2</sub>	blue
266	4 - NH -	1 - NH <sub>2</sub> , 2 - CN	cyan
267	4 - NH -	1 - NH <sub>2</sub> , 2 - S - C <sub>6</sub> H <sub>5</sub>	blue
268	4 - NH -	1 - NH <sub>2</sub> , 2 - S - C <sub>8</sub> H <sub>4</sub> N <sub>2</sub>	blue
269	4 - NH -	1 - NH <sub>2</sub> , 2 - SO <sub>2</sub> - C <sub>6</sub> H <sub>3</sub> Cl <sub>2</sub>	blue

Table 11  
Polydyes From Diacidic Anthraquinone Compounds of Formula XV






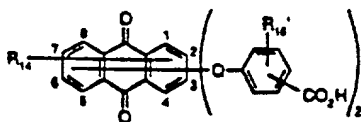
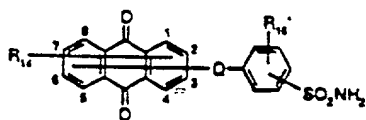
Example No.	$\left( \text{O} - \text{C}_6\text{H}_4 - \text{CO}_2\text{H} \right)_2$	$R_{14}$	Color
270	2,4-di-S-C <sub>6</sub> H <sub>4</sub> -3-CO <sub>2</sub> H	1-NH <sub>2</sub>	red
271	2,3-di-S-C <sub>6</sub> H <sub>4</sub> -4-CO <sub>2</sub> H	1,4-diNH <sub>2</sub>	blue
272	2,4-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1-NHCH <sub>3</sub>	violet
273	2-SO <sub>2</sub> C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H, 4-NHC <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1-NH <sub>2</sub>	blue
274	2-OC <sub>6</sub> H <sub>4</sub> -4-CO <sub>2</sub> H, 4-NHC <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1-NH <sub>2</sub>	violet
275	2-OC <sub>6</sub> H <sub>4</sub> -3-CO <sub>2</sub> H, 4-S-NHC <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1-NH <sub>2</sub>	red
276	2,4-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1-OH	orange
277	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>3</sub>	blue
278	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -3-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> CH(CH <sub>3</sub> ) <sub>2</sub>	blue
279	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -4-CO <sub>2</sub> H	1,8-diNH(CH <sub>2</sub> ) <sub>7</sub> CH <sub>3</sub>	blue
280	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> (C <sub>2</sub> H <sub>5</sub> )C <sub>2</sub> H <sub>5</sub>	blue
281	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHC <sub>6</sub> H <sub>4</sub> -4-CH <sub>3</sub>	blue
282	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHC <sub>6</sub> H <sub>11</sub>	blue
283	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNH(CH <sub>2</sub> ) <sub>5</sub> OH	blue
284	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> C(CH <sub>3</sub> ) <sub>2</sub> CH <sub>2</sub> OH	blue
285	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	blue
286	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> CH <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	blue
287	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> CH=CH <sub>2</sub>	blue
288	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> C≡CH	blue
289	4,5-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,8-diNHCH <sub>2</sub> -  -CH <sub>2</sub> OH	blue
290	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNHC <sub>2</sub> H <sub>5</sub>	blue
291	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNHCH <sub>2</sub> CH(CH <sub>3</sub> )CN	blue
292	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -4-CO <sub>2</sub> H	1,5-diNHCH <sub>2</sub> CH <sub>2</sub> NHCOCH <sub>3</sub>	blue
293	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNH(CH <sub>2</sub> ) <sub>3</sub> OC <sub>2</sub> H <sub>5</sub>	blue
294	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNHCH <sub>2</sub> C <sub>6</sub> H <sub>10</sub> -4-CH <sub>3</sub>	blue
295	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNHCH <sub>2</sub> - 	blue
296	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNHCH <sub>2</sub> - 	blue
297	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNH(CH <sub>2</sub> ) <sub>3</sub> OC <sub>6</sub> H <sub>5</sub>	blue
298	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNHCH(CH <sub>3</sub> )(CH <sub>2</sub> ) <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	blue
299	4,8-di-S-C <sub>6</sub> H <sub>4</sub> -2-CO <sub>2</sub> H	1,5-diNHCH(CH <sub>2</sub> CH <sub>3</sub> ) <sub>2</sub>	blue

Table 11  
Polydyes From Diacidic Anthraquinone Compounds of Formula XV



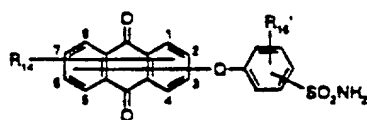
Example No.	$\left( \text{O} - \text{C}_6\text{H}_4 - \text{CO}_2\text{H} \right)_2$	$R_{14}$	Color
300	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diSCH <sub>2</sub> CH <sub>2</sub> OH	red
301	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diSCH <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	red
302	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diSC <sub>6</sub> H <sub>5</sub>	red
303	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diSC <sub>6</sub> H <sub>11</sub>	red
304	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diSC <sub>6</sub> H <sub>4</sub> - 4 - OCH <sub>3</sub>	red
305	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diSC <sub>6</sub> H <sub>4</sub> - 4 - Cl	red
306	4,5 - di - S - C <sub>6</sub> H <sub>4</sub> - 3 - CO <sub>2</sub> H	1,8 - diSC <sub>6</sub> H <sub>4</sub> - 4 - CH <sub>3</sub>	red
307	4,5 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diSC <sub>6</sub> H <sub>3</sub> - 3,4 - diCl	red
308	4,5 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diSC <sub>6</sub> H <sub>4</sub> - 2 - NHCOCH <sub>3</sub>	red
309	4,5 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diSC <sub>6</sub> H <sub>4</sub> - 4 - NHCOC <sub>6</sub> H <sub>5</sub>	red
310	4,5 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diSCH <sub>2</sub> CH <sub>2</sub> OCOCH <sub>3</sub>	red
311	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diSC <sub>6</sub> H <sub>4</sub> - 4 - C(CH <sub>3</sub> ) <sub>3</sub>	red
312	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - dibenzothiazol - 2 - ythio	red
313	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - dibenzoxazol - 2 - ythio	red
314	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diS - $\overline{\text{C} = \text{N} - \text{N}(\text{CH}_3)\text{CH} = \text{N}}$	red
315	2,6 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diNH <sub>2</sub> , 4,8 - diOH	blue
316	2,6 - di - O - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,4,5,8 - tetra NH <sub>2</sub>	blue
317	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - diNH <sub>2</sub> , 2,6 - diBr	blue
318	2,7 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diNH <sub>2</sub> , 4,5 - diNHCO <sub>2</sub> CH <sub>3</sub>	blue
319	2,7 - di - SO <sub>2</sub> - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diNH <sub>2</sub> , 4,5 - diOH	cyan
320	4,5 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diNHCOCH <sub>3</sub>	orange
321	2,7 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diNH <sub>2</sub> , 4,5 - diNHC <sub>6</sub> H <sub>5</sub>	cyan
322	2,6 - di - O - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,8 - diNH <sub>2</sub> , 4,5 - diNHC <sub>6</sub> H <sub>11</sub>	blue
323	2,8 - di - SO <sub>2</sub> - C <sub>6</sub> H <sub>4</sub> - 4 - CO <sub>2</sub> H	1,4,5,8 - tetra NH <sub>2</sub>	cyan
324	4,8 - di - S - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,5 - $\overline{\text{NHCHCH}_2\text{SO}_2\text{CH}_2\text{CH}}$	blue
325	2,3 - di - O - C <sub>6</sub> H <sub>4</sub> - 4 - CO <sub>2</sub> H	1,4 - diNH <sub>2</sub>	violet
326	2,3 - di - SO <sub>2</sub> - C <sub>6</sub> H <sub>4</sub> - 2 - CO <sub>2</sub> H	1,4 - diNH <sub>2</sub>	blue

Table 12  
Polydyes From Diacidic Anthraquinone Compounds of Formula XVI



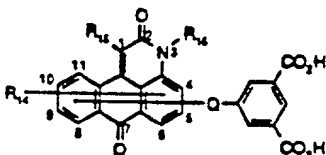
Example No.		R <sub>14</sub>	Color
327		1,4-diOH	orange
328		1-di-NH <sub>2</sub>	violet
329		1-NH <sub>2</sub> , 4-OH	red
330		1-NH <sub>2</sub> , 4-NHC <sub>6</sub> H <sub>5</sub>	violet
331		1-NH <sub>2</sub> , 4-NHC <sub>6</sub> H <sub>4</sub> -4-Cl	blue
332		1-NH <sub>2</sub> , 4-NHC <sub>6</sub> H <sub>4</sub> -4-OCH <sub>3</sub>	blue
333		1-NH <sub>2</sub> , 4-NHSO <sub>2</sub> C <sub>6</sub> H <sub>4</sub> -n	red
334		1-NH <sub>2</sub> , 4-C <sub>6</sub> H <sub>4</sub> -N <sub>2</sub> O	red
335		1-NH <sub>2</sub> , 2-SO <sub>2</sub> -C <sub>6</sub> H <sub>5</sub>	blue-green
336		1-NH <sub>2</sub> , 2-Br	blue
337		1-NH <sub>2</sub> SO <sub>2</sub> C <sub>6</sub> H <sub>3</sub> -3,4-diCl	blue

Table 12  
Polydyes From Diacid Anthraquinone Compounds of Formula XVI



Example No.		R <sub>14</sub>	Color
338		1 - NH <sub>2</sub> , 2 - CN	cyan
339		1 - NH <sub>2</sub> , 2 - NO <sub>2</sub>	cyan
340		1 - NH <sub>2</sub> , 2 - Br	blue
341		1 - NH <sub>2</sub> , 2 - SO <sub>2</sub> N(C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub>	blue
342		1 - NH <sub>2</sub> , 4 - NHC <sub>6</sub> H <sub>4</sub> - 3 - Cl	blue
343		1,8 - diOH, 5 - NO <sub>2</sub>	blue
344		1,5 - diOH, 8 - NH <sub>2</sub>	blue

Table 13  
Polydyes From Diacid Anthrapyridone Compounds of Formula XVIII



Example No.		R <sub>14</sub>	R <sub>15</sub>	R <sub>16</sub>	Color
345		H	CO <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	CH <sub>3</sub>	red
346		H	CN	CH <sub>2</sub> CH(CH <sub>3</sub> ) <sub>2</sub>	violet
347		H	H	C <sub>6</sub> H <sub>9-n</sub>	red
348		H	Cl	C <sub>6</sub> H <sub>11</sub>	red
349		H		CH <sub>3</sub>	red
350		H	CN	C <sub>6</sub> H <sub>5</sub>	violet
351		H		CH <sub>3</sub>	violet
352		H	SO <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	CH <sub>3</sub>	reddish blue

Table 13  
Polydyes From Diacidic Anthrapyridone Compounds of Formula XVIII

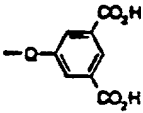
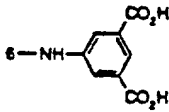
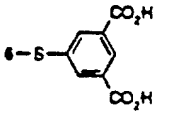
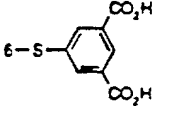
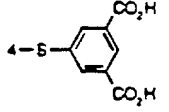
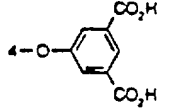
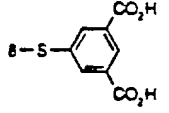
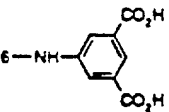
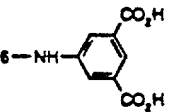
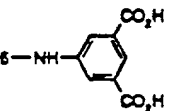
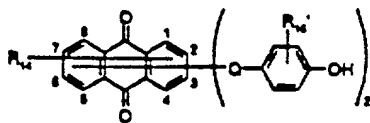
Example No.		R <sub>14</sub>	R <sub>15</sub>	R <sub>16</sub>	Color
353		4 - CH <sub>3</sub>	CO <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	H	red
354		H	CO <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	H	orange
355		H	CN	CH <sub>3</sub>	scarlet
356		6 - NHC <sub>6</sub> H <sub>5</sub>	CN	CH <sub>3</sub>	violet
357		6 - NHC <sub>6</sub> H <sub>4</sub> - 4 - CH <sub>3</sub>	CO <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	CH <sub>3</sub>	red
358		6 - NHC <sub>6</sub> H <sub>5</sub>	H	CH <sub>3</sub>	red
359		H	COC <sub>6</sub> H <sub>5</sub>	CH <sub>2</sub> CH <sub>2</sub> OC <sub>2</sub> H <sub>5</sub>	red
360		H	CN	(CH <sub>2</sub> ) <sub>7</sub> CH <sub>3</sub>	violet
361		4 - Br	CN	CH <sub>3</sub>	violet

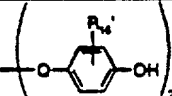



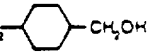






Table 14  
Polydyes From Diacidic Anthraquinone Compounds of Formula XIX



Example No.		R <sub>14</sub>	Color
362	1,5 - diNH -	H	red
363	1,5 - diNH -	H	red
364	1,6 - diNH -	H	red
365	1,6 - diNH -	H	red
366	2,3 - diS -	1,4 - diNH <sub>2</sub>	blue
367	4,5 - diS -	1,8 - diNHCH <sub>2</sub> CH(CH <sub>3</sub> ) <sub>2</sub>	blue
368	4,8 - diS -	1,5 - diSC <sub>6</sub> H <sub>5</sub>	red
369	4,5 - diS -	1,8 - di - S -	red
370	6,7 - diS -	1,4 - diNH -	cyan
371	6,7 - diSO <sub>2</sub> -	1,4 - diNH -	cyan
372	2,3 - diO -	1,4 - diNH <sub>2</sub>	violet



Table 14  
Polydyes From Diacid Anthraquinone Compounds of Formula XIX

Example No.		R <sub>14</sub>	Color
373	4,5 - diNH - 	1,8 - diOH	blue
374	4,5 - diS - 	1,8 - diNHC <sub>6</sub> H <sub>11</sub>	blue
375	4,5 - diS - 	1,8 - diNHCH <sub>2</sub> - 	blue
376	4,5 - diS - 	1,8 - diNHCH <sub>2</sub> C(CH <sub>3</sub> ) <sub>2</sub> CH <sub>2</sub> OH	blue
377	4,5 - diS - 	1,8 - diNHCH <sub>2</sub> CH(C <sub>2</sub> H <sub>5</sub> )C <sub>4</sub> H <sub>9</sub> n	blue
378	2,7 - diS - 	1,4,5,8 - tetra NH <sub>2</sub>	blue
379	2,7 - diS - 	1,8 - diNH <sub>2</sub> , 4,5 - diOH	blue
380	2,7 - diS - 	1,8 - diNH <sub>2</sub> , 4,5 - diNHC <sub>6</sub> H <sub>5</sub>	cyan
381	1,5 - diN(CH <sub>3</sub> )SO <sub>2</sub> - 	H	yellow

$$R_{14}-\text{[Quinone Ring]}-\left( \text{O}-\text{C}_6\text{H}_4(R_{15})-\text{C}_6\text{H}_4-\text{CO}_2\text{H} \right)_n$$
#15890a.xls

Table 15  
Polydyes From Diacidic Anthraquinone Compounds of Formula XIxc

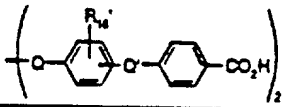
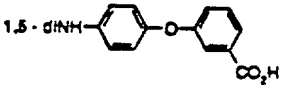
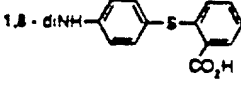
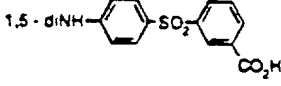
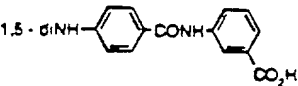
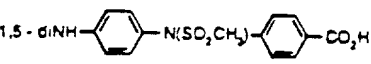
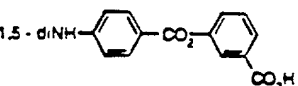
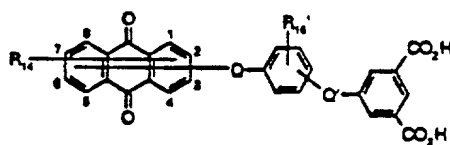
Example No.		R <sub>14</sub>	Color
391		H	red
392		H	red
393		H	red
394		H	red
395		H	red
396		H	red

Table 16  
Polydyes From Diacidic Anthraquinone Compounds of Formula XIXd



Example No.		R <sub>14</sub>	Color
397		H	red
398		1 - NHC <sub>4</sub> H <sub>9-n</sub>	blue
399		1 - NH <sub>2</sub> , 2 - CN	cyan
400		1-NH <sub>2</sub> , 2 - SO <sub>2</sub> N(CH <sub>3</sub> )C <sub>6</sub> H <sub>5</sub>	blue
401		1 - NH <sub>2</sub> , 2 - CF <sub>3</sub>	cyan
402		1-NH <sub>2</sub> , 2 - 6-	blue
403		1 - NH <sub>2</sub> , 2 - OCH <sub>2</sub> CH <sub>2</sub> OH	violet
404		1 - NH <sub>2</sub> , 2 - Br	blue

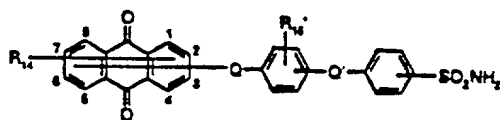
Table 16  
Polydyes From Diacidic Anthraquinone Compounds of Formula XIXd

Example No.		R <sub>14</sub>	Color
405		1-NH <sub>2</sub> , 2-SO <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	blue
406		1-NH <sub>2</sub> , 2-Br	blue
407		1-NH <sub>2</sub> , 4-OH	red
408		1,4-diNH <sub>2</sub>	violet
409		1-NH <sub>2</sub> , 4-NH-C6H4-OCH <sub>3</sub>	violet
410		1-NH <sub>2</sub> , 4-NHSO <sub>2</sub> -C6H4-CH <sub>3</sub>	red
411		1-NH <sub>2</sub> , 4-S-C <sub>8</sub> H <sub>6</sub> N	red
412		1-NH <sub>2</sub> , 4-NHC <sub>6</sub> H <sub>5</sub>	blue

Table 16  
Polydyes From Diacid Anthraquinone Compounds of Formula XIXd

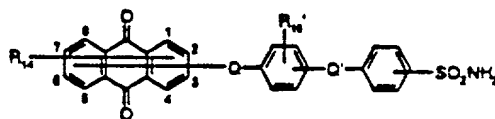
Example No.		$R_{14}$	Color
413		1-NH <sub>2</sub> , 4-NHC <sub>6</sub> H <sub>11</sub>	blue
414		1-NH <sub>2</sub> , 4-NHC <sub>6</sub> H <sub>5</sub>	blue

Table 17  
Polydyes From Diacidic Anthraquinone Compounds of Formula XIXe



Example No.		R <sub>14</sub>	Color
415		H	red
416		1 - NHCH <sub>3</sub>	blue
417		1 - OH	violet
418		1 - NH <sub>2</sub> - 2 - Br	blue
419		1 - NH <sub>2</sub> - 2 - OC <sub>6</sub> H <sub>5</sub>	violet
420		1 - NH <sub>2</sub> - 2 - SO <sub>2</sub> CH <sub>3</sub>	blue
421		1 - NH <sub>2</sub> - 2 - COC <sub>6</sub> H <sub>5</sub>	blue
422		1 - NH <sub>2</sub> - 2 - CF <sub>3</sub>	cyan
423		1 - NH <sub>2</sub> - 2 - CONH <sub>2</sub>	blue
424		1 - NH <sub>2</sub> - 2 - SO <sub>2</sub> N(CH <sub>3</sub> ) <sub>2</sub>	blue
425		1 - NHC <sub>8</sub> H <sub>11</sub>	blue

Table 17  
Polydyes From Diacidic Anthraquinone Compounds of Formula XIx



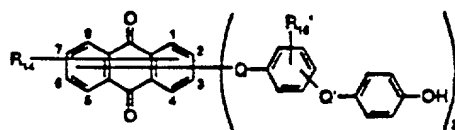
Example No.		R <sub>14</sub>	Color
426		1 - NHC <sub>6</sub> H <sub>5</sub>	green
427		1-NH <sub>2</sub> , 4 - OH	red
428		1 - NH <sub>2</sub> , 4 - NHSO <sub>2</sub> CH <sub>3</sub>	red
429		1 - NH <sub>2</sub> - 4 - NHCO <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	red
430		1 - NH <sub>2</sub> - 4 - NHSO <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	red
431		1 - NH <sub>2</sub> - 4 - NHCOC <sub>2</sub> H <sub>5</sub>	red
432		1,4 - diNH <sub>2</sub>	violet
433		1 - NH <sub>2</sub> - 4 - NHC <sub>6</sub> H <sub>5</sub>	blue
434		1,8 - diOH, 5 - NO <sub>2</sub>	blue
435		1 - NH <sub>2</sub> - 2 - SO <sub>2</sub> C <sub>6</sub> H <sub>5</sub>	blue



Table 18  
Polydyes From Diacid Anthraquinones of Formula XIXf

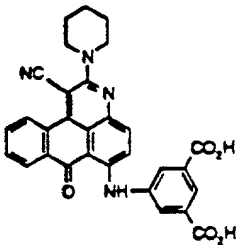
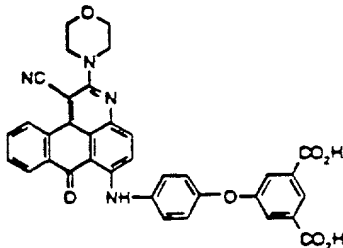
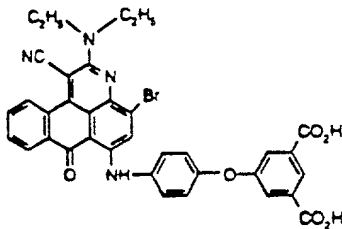
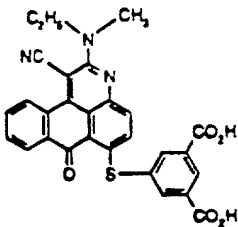
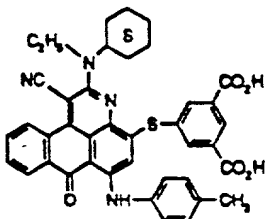
Example No.		R <sub>14</sub>	Color
436		H	blue
437		H	blue
438		H	green
439		H	green
440		H	red
441		H	red
442		1,4 - diNH <sub>2</sub>	violet
443		1,4 - diNH <sub>2</sub>	blue
444		H	red

Table 18  
Polydyes From Diacidic Anthraquinones of Formula XIXf



Example No.		R <sub>14</sub>	Color
445		4,8 - diNH <sub>2</sub> , 3,7-diBr	blue
446		1 - NH <sub>2</sub>	red
447		6,7 - diCl	cyan
448		H	blue
449		H	blue

Table 19  
Polydyes From Diacidic Anthrapyridines

Example No.	Anthrapyridines	Color
450		red
451		bluish-red
452		red
453		orange
454		violet

Example No.	Anthrapyridines	Color
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Example No.	Anthrapyridines	Color
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CN1CCCC1C2=CC=C(C=C2)C(=O)C3=C(C=C4C(=C(C=C4)C(=O)N5C(=C(C=C5)C(=O)N6C(=C(C=C6)C(=O)N7C(=C(C=C7)C(=O)N8C(=C(C=C8)C(=O)N9C(=C(C=C9)C(=O)N10C(=C(C=C10)C(=O)N11C(=C(C=C11)C(=O)N12C(=C(C=C12)C(=O)N13C(=C(C=C13)C(=O)N14C(=C(C=C14)C(=O)N15C(=C(C=C15)C(=O)N16C(=C(C=C16)C(=O)N17C(=C(C=C17)C(=O)N18C(=C(C=C18)C(=O)N19C(=C(C=C19)C(=O)N20C(=C(C=C20)C(=O)N21C(=C(C=C21)C(=O)N22C(=C(C=C22)C(=O)N23C(=C(C=C23)C(=O)N24C(=C(C=C24)C(=O)N25C(=C(C=C25)C(=O)N26C(=C(C=C26)C(=O)N27C(=C(C=C27)C(=O)N28C(=C(C=C28)C(=O)N29C(=C(C=C29)C(=O)N30C(=C(C=C30)C(=O)N31C(=C(C=C31)C(=O)N32C(=C(C=C32)C(=O)N33C(=C(C=C33)C(=O)N34C(=C(C=C34)C(=O)N35C(=C(C=C35)C(=O)N36C(=C(C=C36)C(=O)N37C(=C(C=C37)C(=O)N38C(=C(C=C38)C(=O)N39C(=C(C=C39)C(=O)N40C(=C(C=C40)C(=O)N41C(=C(C=C41)C(=O)N42C(=C(C=C42)C(=O)N43C(=C(C=C43)C(=O)N44C(=C(C=C44)C(=O)N45C(=C(C=C45)C(=O)N46C(=C(C=C46)C(=O)N47C(=C(C=C47)C(=O)N48C(=C(C=C48)C(=O)N49C(=C(C=C49)C(=O)N50C(=C(C=C50)C(=O)N51C(=C(C=C51)C(=O)N52C(=C(C=C52)C(=O)N53C(=C(C=C53)C(=O)N54C(=C(C=C54)C(=O)N55C(=C(C=C55)C(=O)N56C(=C(C=C56)C(=O)N57C(=C(C=C57)C(=O)N58C(=C(C=C58)C(=O)N59C(=C(C=C59)C(=O)N60C(=C(C=C60)C(=O)N61C(=C(C=C61)C(=O)N62C(=C(C=C62)C(=O)N63C(=C(C=C63)C(=O)N64C(=C(C=C64)C(=O)N65C(=C(C=C65)C(=O)N66C(=C(C=C66)C(=O)N67C(=C(C=C67)C(=O)N68C(=C(C=C68)C(=O)N69C(=C(C=C69)C(=O)N70C(=C(C=C70)C(=O)N71C(=C(C=C71)C(=O)N72C(=C(C=C72)C(=O)N73C(=C(C=C73)C(=O)N74C(=C(C=C74)C(=O)N75C(=C(C=C75)C(=O)N76C(=C(C=C76)C(=O)N77C(=C(C=C77)C(=O)N78C(=C(C=C78)C(=O)N79C(=C(C=C79)C(=O)N80C(=C(C=C80)C(=O)N81C(=C(C=C81)C(=O)N82C(=C(C=C82)C(=O)N83C(=C(C=C83)C(=O)N84C(=C(C=C84)C(=O)N85C(=C(C=C85)C(=O)N86C(=C(C=C86)C(=O)N87C(=C(C=C87)C(=O)N88C(=C(C=C88)C(=O)N89C(=C(C=C89)C(=O)N90C(=C(C=C90)C(=O)N91C(=C(C=C91)C(=O)N92C(=C(C=C92)C(=O)N93C(=C(C=C93)C(=O)N94C(=C(C=C94)C(=O)N95C(=C(C=C95)C(=O)N96C(=C(C=C96)C(=O)N97C(=C(C=C97)C(=O)N98C(=C(C=C98)C(=O)N99C(=C(C=C99)C(=O)N100C(=C(C=C100)C(=O)N101C(=C(C=C101)C(=O)N102C(=C(C=C102)C(=O)N103C(=C(C=C103)C(=O)N104C(=C(C=C104)C(=O)N105C(=C(C=C105)C(=O)N106C(=C(C=C106)C(=O)N107C(=C(C=C107)C(=O)N108C(=C(C=C108)C(=O)N109C(=C(C=C109)C(=O)N110C(=C(C=C110)C(=O)N111C(=C(C=C111)C(=O)N112C(=C(C=C112)C(=O)N113C(=C(C=C113)C(=O)N114C(=C(C=C114)C(=O)N115C(=C(C=C115)C(=O)N116C(=C(C=C116)C(=O)N117C(=C(C=C117)C(=O)N118C(=C(C=C118)C(=O)N119C(=C(C=C119)C(=O)N120C(=C(C=C120)C(=O)N121C(=C(C=C121)C(=O)N122C(=C(C=C122)C(=O)N123C(=C(C=C123)C(=O)N124C(=C(C=C124)C(=O)N125C(=C(C=C125)C(=O)N126C(=C(C=C126)C(=O)N127C(=C(C=C127)C(=O)N128C(=C(C=C128)C(=O)N129C(=C(C=C129)C(=O)N130C(=C(C=C130)C(=O)N131C(=C(C=C131)C(=O)N132C(=C(C=C132)C(=O)N133C(=C(C=C133)C(=O)N134C(=C(C=C134)C(=O)N135C(=C(C=C135)C(=O)N136C(=C(C=C136)C(=O)N137C(=C(C=C137)C(=O)N138C(=C(C=C138)C(=O)N139C(=C(C=C139)C(=O)N140C(=C(C=C140)C(=O)N141C(=C(C=C141)C(=O)N142C(=C(C=C142)C(=O)N143C(=C(C=C143)C(=O)N144C(=C(C=C144)C(=O)N145C(=C(C=C145)C(=O)N146C(=C(C=C146)C(=O)N147C(=C(C=C147)C(=O)N148C(=C(C=C148)C(=O)N149C(=C(C=C149)C(=O)N150C(=C(C=C150)C(=O)N151C(=C(C=C151)C(=O)N152C(=C(C=C152)C(=O)N153C(=C(C=C153)C(=O)N154C(=C(C=C154)C(=O)N155C(=C(C=C155)C(=O)N156C(=C(C=C156)C(=O)N157C(=C(C=C157)C(=O)N158C(=C(C=C158)C(=O)N159C(=C(C=C159)C(=O)N160C(=C(C=C160)C(=O)N161C(=C(C=C161)C(=O)N162C(=C(C=C162)C(=O)N163C(=C(C=C163)C(=O)N164C(=C(C=C164)C(=O)N165C(=C(C=C165)C(=O)N166C(=C(C=C166)C(=O)N167C(=C(C=C167)C(=O)N168C(=C(C=C168)C(=O)N169C(=C(C=C169)C(=O)N170C(=C(C=C170)C(=O)N171C(=C(C=C171)C(=O)N172C(=C(C=C172)C(=O)N173C(=C(C=C173)C(=O)N174C(=C(C=C174)C(=O)N175C(=C(C=C175)C(=O)N176C(=C(C=C176)C(=O)N177C(=C(C=C177)C(=O)N178C(=C(C=C178)C(=O)N179C(=C(C=C179)C(=O)N180C(=C(C=C180)C(=O)N181C(=C(C=C181)C(=O)N182C(=C(C=C182)C(=O)N183C(=C(C=C183)C(=O)N184C(=C(C=C184)C(=O)N185C(=C(C=C185)C(=O)N186C(=C(C=C186)C(=O)N187C(=C(C=C187)C(=O)N188C(=C(C=C188)C(=O)N189C(=C(C=C189)C(=O)N190C(=C(C=C190)C(=O)N191C(=C(C=C191)C(=O)N192C(=C(C=C192)C(=O)N193C(=C(C=C193)C(=O)N194C(=C(C=C194)C(=O)N195C(=C(C=C195)C(=O)N196C(=C(C=C196)C(=O)N197C(=C(C=C197)C(=O)N198C(=C(C=C198)C(=O)N199C(=C(C=C199)C(=O)N200C(=C(C=C200)C(=O)N201C(=C(C=C201)C(=O)N202C(=C(C=C202)C(=O)N203C(=C(C=C203)C(=O)N204C(=C(C=C204)C(=O)N205C(=C(C=C205)C(=O)N206C(=C(C=C206)C(=O)N207C(=C(C=C207)C(=O)N208C(=C(C=C208)C(=O)N209C(=C(C=C209)C(=O)N210C(=C(C=C210)C(=O)N211C(=C(C=C211)C(=O)N212C(=C(C=C212)C(=O)N213C(=C(C=C213)C(=O)N214C(=C(C=C214)C(=O)N215C(=C(C=C215)C(=O)N216C(=C(C=C216)C(=O)N217C(=C(C=C217)C(=O)N218C(=C(C=C218)C(=O)N219C(=C(C=C219)C(=O)N220C(=C(C=C220)C(=O)N221C(=C(C=C221)C(=O)N222C(=C(C=C222)C(=O)N223C(=C(C=C223)C(=O)N224C(=C(C=C224)C(=O)N225C(=C(C=C225)C(=O)N226C(=C(C=C226)C(=O)N227C(=C(C=C227)C(=O)N228C(=C(C=C228)C(=O)

red

Table 20  
Polydyes From Diacidic Nitroarylamines

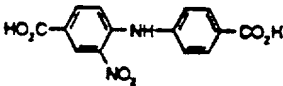
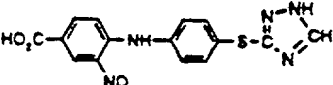
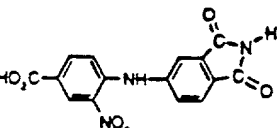
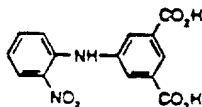
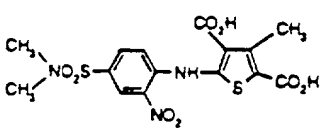
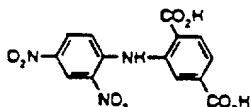
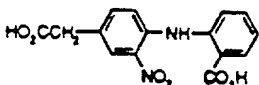
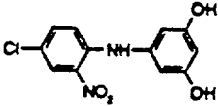
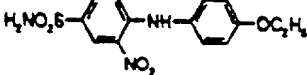
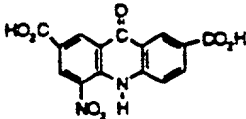
Example No.	Nitroarylamine Compound	Color
456		yellow
457		yellow
458		yellow
459		yellow
460		yellow
461		yellow
462		yellow
463		yellow
464		yellow

Table 20  
Polydyes From Diacidic Nitroarylamines

Example No.	Nitroaryamine Compound	Color
455		yellow

006341-037560

Table 21  
Miscellaneous Polydyes

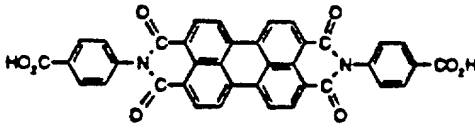
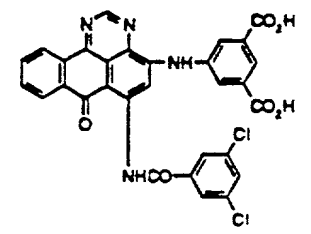
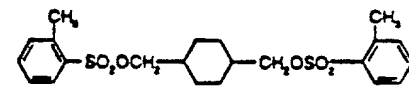
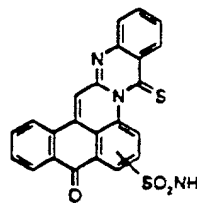
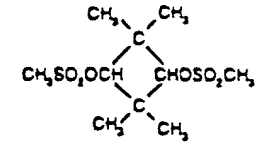
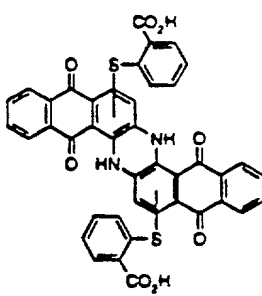
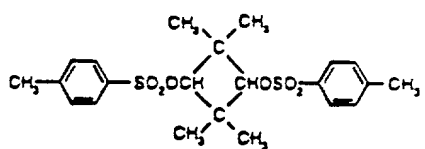
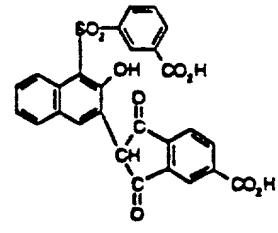
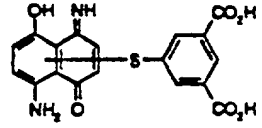
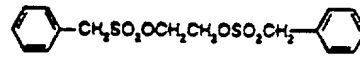
Example No.	Diacidic Compound Reacted	Disulfonate Compound Reacted	Color
466		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$	red
467			orange
468			yellow
469			blue
470		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{O}-\text{C}_6\text{H}_4-\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	yellow
471			blue

Table 21  
Miscellaneous Polydyes

Example No.	Diacid Compound Reacted	Disulfonate Compound Reacted	Color
472			yellow
473		$\text{CH}_3\text{SO}_2(\text{OCH}_2\text{CH}_2)_2\text{OSO}_2\text{CH}_3$	yellow
474		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{SCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	red
475		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{N}(\text{CH}_2\text{CH}_2\text{OSO}_2\text{CH}_3)_2$	red
476		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{N}(\text{CH}_2\text{CH}_2\text{OSO}_2\text{CH}_3)_2$	violet
477		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{N}(\text{CH}_2\text{CH}_2\text{OSO}_2\text{CH}_3)_2$	yellow
478		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	orange



Table 21  
Miscellaneous Polydyes

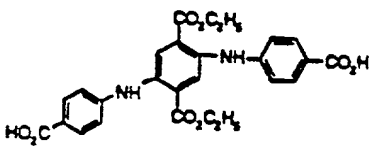
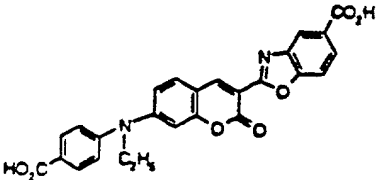
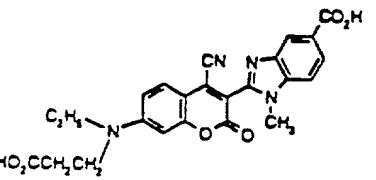
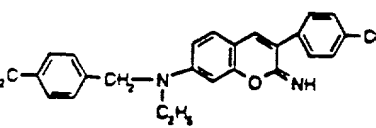
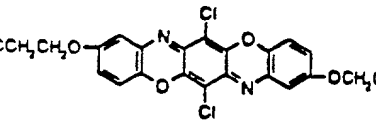
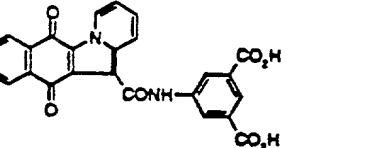
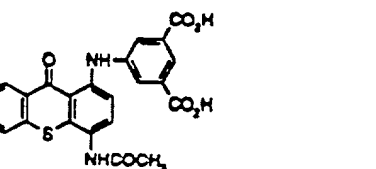
Example No.	Diacidic Compound Reacted	Disulfonate Compound Reacted	Color
479		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	orange
480		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{OSO}_2\text{CH}_3$	yellow
481		$\text{CH}_3\text{CH}_2\text{SO}_2\text{OCH}_2\text{CH}_2\text{CH}_2\text{OSO}_2\text{CH}_2\text{CH}_3$	red
482		$n\text{-C}_4\text{H}_9\text{SO}_2\text{OCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OSO}_2\text{C}_4\text{H}_9$	yellow
483		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_4\text{OSO}_2\text{CH}_3$	red
484		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{---}(\text{cyclohexyl})\text{---CH}_2\text{OSO}_2\text{CH}_3$	orange
485		$\text{C}_6\text{H}_5\text{---SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	reddish-yellow

Table 21  
Miscellaneous Polydyes

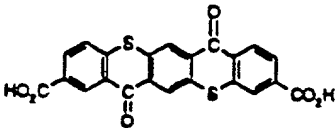
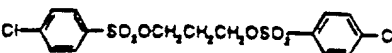
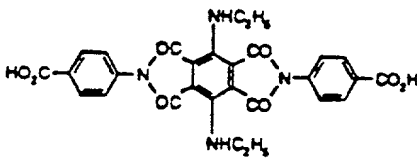
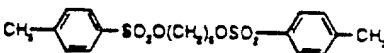
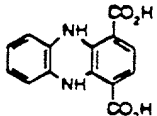
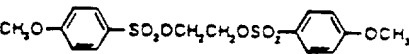
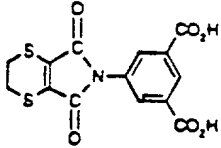
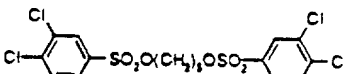
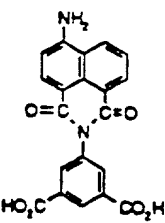
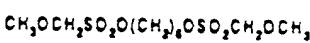
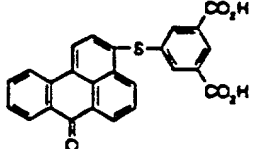
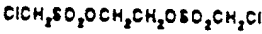
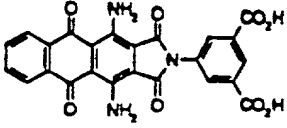
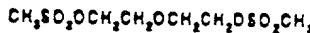
Example No.	Diacid Compound Reacted	Disulfonate Compound Reacted	Color
486			yellow
487			blue
488			orange
489			yellow
490			yellow
491			yellow
492			blue

Table 21  
Miscellaneous Polydyes

Example No.	Diacid Compound Reacted	Disulfonate Compound Reacted	Color
493		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{C}_6\text{H}_4\text{CH}_2\text{OSO}_2\text{CH}_3$	greenish-blue
494		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{C}_6\text{H}_4\text{SCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	greenish-blue
495		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{C}_6\text{H}_4\text{SCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	red
496		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{SO}_2\text{CH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	orange
497		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_{12}\text{OSO}_2\text{CH}_3$	red

Table 21  
Miscellaneous Polydyes

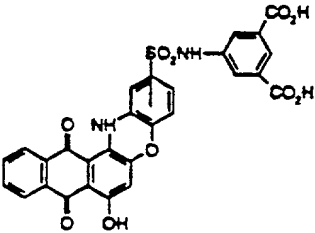
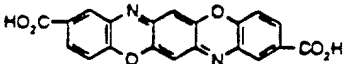
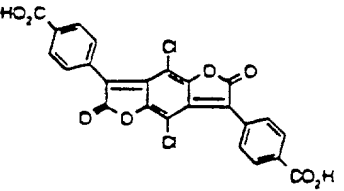
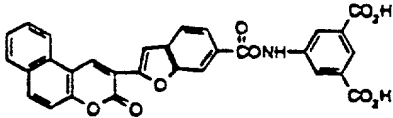
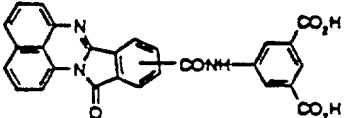
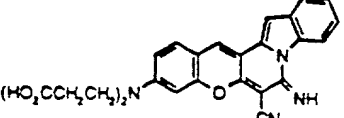
Example No.	Diacid Compound Reacted	Disulfonate Compound Reacted	Color
498		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{C}(\text{CH}_3)_2\text{CH}_2\text{OSO}_2\text{CH}_3$	blue
499		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{O}-\text{C}_6\text{H}_4-\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	orange
500		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$	red
501	$[\text{CuPc}] \text{SO}_2\text{NH}_2$ Pc = phthalocyanine	$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	blue
502	$[\text{CuPc}] \text{SO}_2\text{NH}-\text{C}_6\text{H}_3(\text{CO}_2\text{H})_2$	$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	blue
503		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{SCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	yellow
504		$\text{CH}_3\text{SO}_2\text{OCH}_2-\text{C}_6\text{H}_{10}-\text{OCH}_2\text{OSO}_2\text{CH}_3$	reddish-yellow
505		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$	red

Table 22  
Polymeric UV Absorbers

Example No.	Diacid Compound Reacted	Disulfonate Compound Reacted
506		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$
507		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
508		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{C}_6\text{H}_4\text{CH}_2\text{OSO}_2\text{CH}_3$
509		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$
510		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$
511		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$
512		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
513		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$

Table 22  
Polymeric UV Absorbers

Example No.	Diacid Compound Reacted	Disulfonate Compound Reacted
514		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$
515		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{OSO}_2\text{CH}_3$
516		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2\text{CH}_2\text{O})_3\text{SO}_2\text{CH}_3$
517		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
518		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{SCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
519		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_6\text{OSO}_2\text{CH}_3$
520		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}(\text{Cl})\text{CH}_2\text{OSO}_2\text{CH}_3$
521		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}(\text{Cl})\text{CH}_2\text{OSO}_2\text{CH}_3$
522		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}(\text{Cl})\text{CH}_2\text{OSO}_2\text{CH}_3$

Table 23  
Polymeric Infrared Light Absorbers

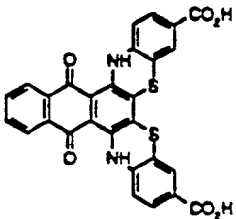
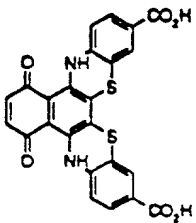
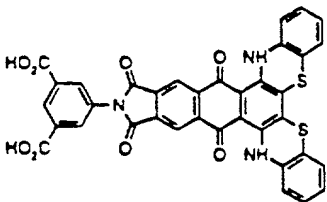
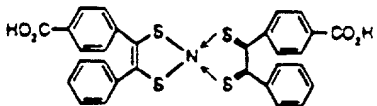
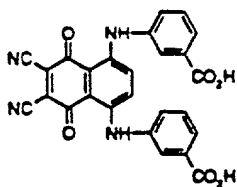
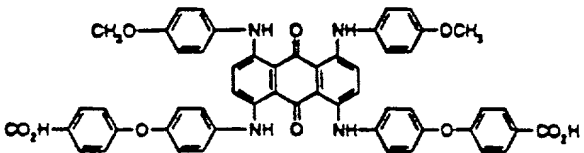
Example No.	Diacidic Compound Reacted	Disulfonate Compound Reacted
523		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
524		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_4\text{OSO}_2\text{CH}_3$
525		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_4\text{OSO}_2\text{CH}_3$
526		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_4\text{OSO}_2\text{CH}_3$
527		$\text{CH}_3\text{SO}_2\text{O}(\text{CH}_2)_{12}\text{OSO}_2\text{CH}_3$
528		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$

Table 23  
Polymeric Infrared Light Absorbers

Example No.	Diacid Compound Reacted	Disulfonate Compound Reacted
529		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
530		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
531		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
532		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
533		$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
534	$[\text{PcAlOH}] - \text{SO}_2\text{NH}_2$ Pc = phthalocyanine	$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$
535	$\text{NcSi}(\text{OC}_2\text{H}_5)_2 - \text{Si}(\text{OC}_2\text{H}_5)_2$ Nc = naphthalocyanine	$\text{CH}_3\text{SO}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$



Table 23  
Polymeric Infrared Light Absorbers

Example No.	Diacidic Compound Reacted	Disulfonate Compound Reacted
536	$[PcAlOH](CO_2H)_2$ <p>Pc = phthalocyanine</p>	$CH_3SO_3O(CH_2)_6OSO_3CH_3$

006267-9375260

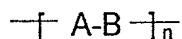
Example 537

- To dimethylformamide (DMF, 45.0 mL) was added 1,5-bis(2-carboxyphenyl-thio)anthraquinone (17.9 g, 0.035 mole).
- 5 The mixture was stirred for about 10 minutes and then 1,8-diazabicyclo[5,4,0]undec-7-ene (DBU, 10.4 g, 0.068 mole) was added followed by 1,2-ethanediol, dimethanedisulfonate (7.64 g, 0.035 mole) and additional DMF (10.0 mL). The reaction mixture was stirred at about 110°C for 3.0 hours
- 10 and allowed to cool to about 55°C. Methanol (35.0 g) was added dropwise with stirring followed by water (20 mL) and acetic acid (3.0 mL). The yellow solid was collected by filtration, washed with methanol (25 mL), warm water (50 mL) and then finally with methanol to facilitate drying.
- 15 The yield of polymeric and cyclic yellow product was 18.2 g.

Examples 538-543

- 20 The procedure described in Example 537 was repeated exactly except that the reactant X-B-X<sub>1</sub> used in each example was as follows:
- Example 538: 1,3-propanediol, dimethanedisulfonate
- Example 539: 1,4-butanediol, dimethanedisulfonate
- 25 Example 540: diethylene glycol, dimethanedisulfonate
- Example 541: 1,6-hexanediol, dimethanedisulfonate
- Example 542: 1,4-cyclohexanedimethanol, dimethanedisulfonate
- Example 543: 1,12-dodecanediol, dimethanedisulfonate
- 30 The weight yields (Weight, g), percent yields (Yield), weight average molecular weights (M<sub>w</sub>), number average molecular weights (M<sub>n</sub>), and polydispersities (M<sub>w</sub>/M<sub>n</sub>, by GPC) for each of the light absorbing (yellow) compositions prepared in Examples 537-543 are presented in Table 24.

The percent yields were calculated by dividing the actual weight of the polymer



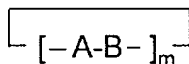
obtained in grams by the theoretical number of moles of repeating unit multiplied by the gram molecular weight of the repeating unit and then multiplying the number thus obtained by 100.

TABLE 24

10

<u>Example</u>	<u>Weight</u>	<u>Yield</u>	<u>Mw</u>	<u>Mn</u>	<u>Mw/Mn</u>
537	18.2	96.6	7,512	1,106	6.8
538	18.8	97.2	5,224	1,051	5.0
539	19.4	98.0	13,856	1,202	11.5
15 540	20.1	98.7	9,849	1,840	5.4
541	20.2	97.4	7,396	870	8.5
542	20.7	95.3	3,685	808	4.6
543	22.7	95.5	5,503	1,116	4.9

20 The approximate amount of cyclic compounds of formula I-A present in each of the light absorbing compositions produced in Examples 537-543 was determined by a combination of GPC, NMR and FDMS analyses. The weight percentages of the cyclic compounds having the structure



25

I-A

wherein m is 1, 2, 3, or 4 present in the composition of Examples 537-543 is set forth in Table 25. The remainder or balance of the light absorbing compositions was presumed to be a linear polymer.

30

TABLE 25

	<u>Example</u>	<u>m=1</u>	<u>m=2</u>	<u>m=3</u>	<u>m=4</u>
	537	0.1	3.4	1.5	0.7
5	538	0.8	3.3	1.3	<0.1
	539	6.9	4.1	1.3	0.2
	540	9.5	3.1	0.8	<0.1
	541	26.7	4.0	1.1	0.2
	542	20.4	4.0	0.7	<0.1
10	543	20.1	2.2*	0.5*	<0.1

\*These peaks are unidentified but appear not to be oligomers.